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Abstract

Decagonal quasicrystals are made of pairs of atomic planes with pentagonal symmetry periodically stacked along a 10-fold axis. We have investigated the atomic structure of the 2-fold surface of a decagonal Al-Ni-Co quasicrystal using scanning tunneling microscopy. The surface consists of terraces separated by steps of heights 1.9, 4.7, 7.8, and 12.6 Å containing rows of atoms parallel to the 10-fold direction with an internal periodicity of 4 Å. The rows are arranged aperiodically, with separations that follow a Fibonacci sequence and inflation symmetry. The results indicate that the surfaces are preferentially Al-terminated and in general agreement with bulk models.

Keywords

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Disciplines

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Comments

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Decagonal quasicrystals are made of pairs of atomic planes with pentagonal symmetry periodically stacked along a 10-fold axis. We have investigated the atomic structure of the 2-fold surface of a decagonal Al-Ni-Co quasicrystal using scanning tunneling microscopy. The surface consists of terraces separated by steps of heights 1.9, 4.7, 7.8, and 12.6 Å containing rows of atoms parallel to the 10-fold direction with an internal periodicity of 4 Å. The rows are arranged aperiodically, with separations that follow a Fibonacci sequence and inflation symmetry. The results indicate that the surfaces are preferentially Al-terminated and in general agreement with bulk models.

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Unraveling the relationship between atomic structure and physical properties drives much of physics and materials science. Quasicrystals provide a special challenge because their atomic structure—well-ordered but not periodic (Ref. 1)—is clearly associated with an array of unusual mechanical^{2,3} and electronic⁴ properties. While most known quasicrystals are icosahedral, with quasiperiodic bulk structure in all three dimensions, the decagonal quasicrystals have a structure that is periodic in one dimension and quasiperiodic in two.⁵ They possess a 10-fold symmetry axis along the periodic direction, and two sets of five equivalent 2-fold symmetry axes rotated by 36° in the quasiperiodic plane. The 2-fold surfaces thus provide, within the same surface, a periodic direction along the 10-fold axis, and an aperiodic direction perpendicular to it. Surface properties that might depend on the existence of periodicity, such as friction, can thus be uniquely studied.³

In this paper, we present a microscopy study of such a surface to unravel its atomic structure. We show that, as predicted from the bulk, it contains both periodic and aperiodic atomic arrangements that, together with a characteristic step-terrace structure, indicate that it exposes the Al-rich planes of the bulk.

The bulk structure of decagonal quasicrystals has been studied by high-resolution transmission electron microscopy and x-ray and electron diffraction.⁵ These studies show that it consists of ordered arrangements of columnar clusters with pentagonal symmetry and diameter of 20 Å. The 10-fold surface of the Al-Ni-Co decagonal quasicrystal—the surface that is quasicrystalline in both dimensions—has been studied by low-energy electron diffraction (LEED), ion scattering spectroscopy, and scanning tunneling microscopy^{6–10} (STM), but there are still important issues, such as atomic structure and composition, that remain controversial.^{8,9} Some authors have attributed sites of high tunneling probability in the STM

images to transition metal atoms,⁸ while others have identified these sites as Al atoms.^{7,9} A further point of discussion is whether the surface termination is enriched in Al relative to the bulk composition.^{6,7} In contrast to the 10-fold surface however, only a few studies of the 2-fold surface^{6,11,12} have been reported.

The experiments were performed in an ultrahigh vacuum chamber (base pressure in the 10^{−10} Torr range) equipped with LEED, Auger electron spectroscopy, and an atomic force microscopy (AFM)-STM head from RHK-tech.¹³ Samples and cantilevers were transferred from air through a load lock. Two types of conductive cantilevers were used, coated with approximately 30 nm of either W₂C or TiN, with a spring constant of 48 or 90 N/m.¹⁴ The high stiffness of the cantilevers suppressed the jump to contact instability and ensured stable tunneling.

Single grain decagonal Al₇₂Ni₁₁Co₁₇ quasicrystals were grown at the Ames Laboratory at Iowa State University, and cut and polished to expose the (10 000) surface.¹⁵ In the UHV chamber they were cleaned by cycles of Ar⁺ sputtering at 1 keV, and annealed for 1 to 2 h by electron bombardment at a temperature between 1100 and 1200 K, as monitored by an optical pyrometer.^{13,16} As reported by Cox *et al.* the surface morphology is critically dependent on annealing temperature.¹⁰ We found indeed that the surface undergoes a phase transformation above 1200 K.

Figure 1(a) shows a STM image of the surface at $V_s = +1.0$ V, $I_t = 0.1$ nA. Terraces separated by several types of atomic steps are visible. From the line profile [Fig. 1(b)], the heights of the steps are 1.9 ± 0.1 Å, 4.7 ± 0.1 Å (short, S_H), and 7.8 ± 0.2 Å (long, L_H). We also observed steps of 12.6 ± 0.3 Å height ($S_H + L_H$). Given the lack of periodicity, the existence of steps with such peculiar heights is a clear indication that particular atomic terminations, aluminum as

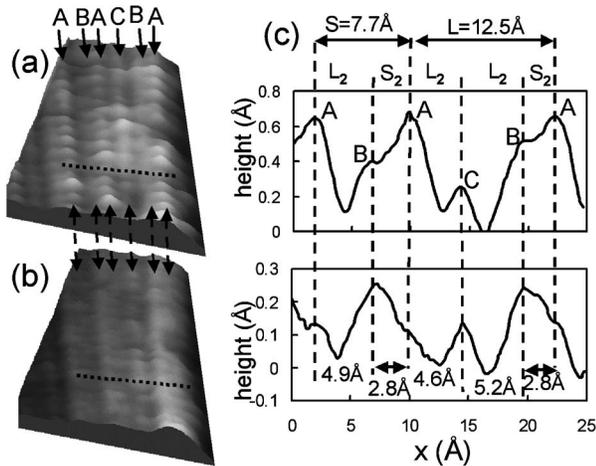


FIG. 3. STM images of the same area acquired at positive and negative bias. (a) $V_s = +1.2$ V, $I = 0.1$ nA (b) $V_s = -1.2$ V, $I = 0.1$ nA. (c) Line profiles across the dashed lines in (a) and (b) reveal three electronically different atomic rows, labeled A, B, and C.

high temperature (HT) treatment performed by these authors. According to our observations this changes the structure of the surface into a new phase, probably due to preferential evaporation of Al. Additional information is provided by images acquired at different polarities. Since the tunneling current depends on the local density of states, it provides information on the electronic character of the surface. Figures 3(a) and 3(b) show STM empty and filled states images of the same area. The profiles across the rows, shown in Fig. 3(c), indicate that the contrast depends on the polarity and that there are three types of electronically different rows, denoted A, B, and C. Type A rows have the highest contrast in the empty state images [marked by lines in Fig. 2(c)], and lowest in the filled state images. Type B rows have opposite contrast behavior, i.e., low in empty state images, and high in filled state images. Type C rows are separated from A and B by an L_2 segment and have relatively constant contrast in both polarities.

With this information we can proceed to examine proposed structural models based on bulk x-ray diffraction.⁵ We examined three models.^{8,20–22}

First we consider the model of Deloudi *et al.*²¹ for the Co-rich phase of decagonal Al-Co-Ni. This model is based on a columnar cluster with 20 Å diameter extracted from the W approximant, $\text{Al}_{72.5}\text{Co}_{20}\text{Ni}_{7.5}$. This model proposes 8-Å periodicity, inconsistent with the 4-Å periodicity found in our STM results.

The second model, due to Cervellino *et al.*^{8,22} is for the Ni-rich phase of Al-Ni-Co. Although the 4-Å periodicity is present in this model, the observed L and S sequence, as well as that of the daughter segments L_2 and S_2 , are not predicted by the model.

Hence, we conclude that the best match (but not perfect as we shall see) is with the model of Takakura *et al.*²⁰ for $\text{Al}_{72}\text{Ni}_{20}\text{Co}_8$. First, its interlayer periodicity matches the 4 Å observed in the STM images. Second, among the many possible planes that can be generated by cuts parallel to the 2-fold plane, the planes made of pure Al are the only ones

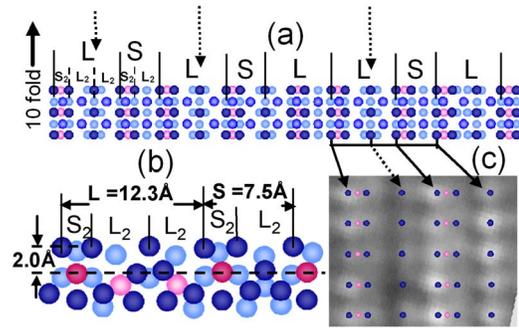


FIG. 4. (Color online) Atomic model of the 2-fold surface derived from the bulk x-ray diffraction model of Takakura *et al.* (Ref. 20). The Al is blue and transition metal (Ni, Co) is red. The plane shown here is the one containing only Al atoms, which provides the best fit with the experimental STM images. (a) Top view ($8 \text{ \AA} \times 90 \text{ \AA}$). Three levels of color of the Al atoms corresponding to relative distances below the surface (dark=top; medium= -0.5 \AA , light= -2 \AA). The spacing between Al atoms along the 10-fold direction is 4.1 \AA , and the sequence along the 2-fold direction shows the same Fibonacci sequence ($LSLSLLSL$ or $S_2L_2L_2S_2L_2S_2L_2L_2S_2L_2L_2S_2L_2L_2S_2L_2L_2S_2L_2L_2S_2L_2L_2$) as the STM image in Fig. 2(c). Dashed arrows mark the positions of local mirror planes. (b) A cross section showing the atomic arrangement perpendicular to the surface (perpendicular to the 10-fold direction). A transition metal rich layer is visible below ($\sim 2.0 \text{ \AA}$ lower) the topmost Al one. (c) A STM image of section of L and S with the Al sites indicated by blue dots.

that: (i) contain rows with spacings matching the observed L and S sequences and also the daughter segments L_2 and S_2 , including the inversion discussed above, as shown in Fig. 4(a); (ii) contain three compositionally different rows that can match the polarity dependent types A, B, and C observed in the STM images; (iii) are separated by distances that match the observed step heights S_H and L_H ; and (iv) are separated by 2 \AA from a plane containing transition metals (Co, Ni) that might correspond to the 2-Å terraces.

The Al termination [Fig. 4(b)] is presumably driven by the lower surface energy of Al relative to the two transition metals, consistent with the observation that icosahedral quasicrystals exhibit Al-rich terminations, selected from among the possible bulk layers.²³

Figure 4(c) shows that the positions of the topmost Al atoms in that surface map directly onto the positions of high tunneling probability in the STM image. These positions do not depend upon bias voltage, although their relative intensities do, as in Fig. 3. Hence the “bumps” in the STM image represent Al atoms on the 2-fold surface.

The agreement with the Takakura model, however, is not perfect. The model exhibits local mirror symmetry planes, shown by the broken line arrows in Fig. 4(a). This symmetry is lost by the different contrast of the Al atoms [Fig. 4(c)], which varies with bias voltage. Even the partially good agreement with the model is remarkable however, because of the large number of structural details that can be explained with it. Nonetheless, the exceptions to the agreement should not be minimized. Another is the fact that the model is for the Ni rich alloy (Ni:Co=20:8), which has a 4-Å periodicity, while the nominal composition of our alloy is Ni:Co

=11:17, which borders on a phase with 8-Å periodicity at low temperature.⁵ However, our sample could be quenched in from the 5*f* HT state, which is a high-temperature basic Co-rich phase with a 4-Å periodicity.²⁴ Since both HT Co-rich and Ni-rich decagonal phases correspond to pentagonal tilings and have a 4-Å period, the Takakura model could provide a reasonable structural interpretation.

In conclusion, we have shown that the atomic structure of the 2-fold (10 000) surface of Al-Ni-Co mainly consists of atomic rows of atoms with 4-Å periodicity along the 10-fold direction and aperiodically spaced in the 2-fold direction. They occupy terraces spaced by the same distances that separate the Al planes in the bulk. The rows are separated by short and long distances in a Fibonacci sequence. Nested

daughter rows are present that extend the sequence (inflation symmetry), confirming quasicrystallinity in the aperiodic direction. The observations follow closely, but not exactly, the bulk model of Al₇₂Ni₂₀Co₈ derived by Takakura *et al.*

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¹D. Shechtman, I. Blech, D. Gratias, and J. W. Cahn, *Phys. Rev. Lett.* **53**, 1951 (1984).

²J.-M. Dubois, *Useful Quasicrystals* (World Scientific, London, 2005).

³J. Y. Park, D. F. Ogletree, M. Salmeron, R. A. Ribeiro, P. C. Canfield, C. J. Jenks, and P. A. Thiel, *Science* **309**, 1354 (2005).

⁴E. Rotenberg, W. Theis, K. Horn, and P. Gille, *Nature (London)* **406**, 602 (2000).

⁵W. Steurer, *Z. Kristallogr.* **219**, 391 (2004).

⁶M. Kishida, Y. Kamimura, R. Tamura, K. Edagawa, S. Takeuchi, T. Sato, Y. Yokoyama, J. Q. Guo, and A. P. Tsai, *Phys. Rev. B* **65**, 094208 (2002).

⁷J. Yuhara, J. Klikovits, M. Schmid, P. Varga, Y. Yokoyama, T. Shishido, and K. Soda, *Phys. Rev. B* **70**, 024203 (2004).

⁸N. Ferralis, K. Pussi, E. J. Cox, M. Gierer, J. Ledieu, I. R. Fisher, C. J. Jenks, M. Lindroos, R. McGrath, and R. D. Diehl, *Phys. Rev. B* **69**, 153404 (2004).

⁹H. R. Sharma, K. J. Franke, W. Theis, A. Riemann, S. Fölsch, P. Gille, and K. H. Rieder, *Phys. Rev. B* **70**, 235409 (2004).

¹⁰Erik J. Cox, Julian Ledieu, R. McGrath, R. D. Diehl, C. J. Jenks, and I. Fisher, in *Quasicrystals—Preparation, Properties, and Applications*, edited by E. Belin-Ferre, P. A. Thiel, A. P. Tsai, and K. Urban, MRS Symposia Proceeding No. 643 (Materials Research Society, Pittsburgh, 2001), p. K11.3.1.

¹¹W. Theis, E. Rotenberg, K. J. Franke, P. Gille, and K. Horn, *Phys. Rev. B* **68**, 104205 (2003).

¹²H. R. Sharma, K. J. Franke, W. Theis, A. Riemann, S. Fölsch, K.

H. Rieder, and P. Gille, *Surf. Sci.* **561**, 121 (2004).

¹³J. Y. Park, D. F. Ogletree, M. Salmeron, C. J. Jenks, and P. A. Thiel, *Tribol. Lett.* **17**, 629 (2004).

¹⁴NT-MDT Co., Zelenograd Research Institute of Physical Problems, Moscow, Russia.

¹⁵I. R. Fisher, M. J. Kramer, Z. Islam, A. R. Ross, A. Kracher, T. Wiener, M. J. Sailer, A. I. Goldman, and P. C. Canfield, *Philos. Mag. B* **79**, 425 (1999).

¹⁶J. Y. Park, D. F. Ogletree, M. Salmeron, R. A. Ribeiro, P. C. Canfield, C. J. Jenks, and P. A. Thiel, *Phys. Rev. B* **71**, 144203 (2005).

¹⁷J. Y. Park, D. F. Ogletree, M. Salmeron, R. A. Ribeiro, P. C. Canfield, C. J. Jenks, and P. A. Thiel (unpublished).

¹⁸T. Cai, V. Fournee, T. Lograsso, A. Ross, and P. A. Thiel, *Phys. Rev. B* **65**, 140202(R) (2002).

¹⁹J. Ledieu, J. T. Hoefl, D. E. Reid, J. A. Smerdon, R. D. Diehl, T. A. Lograsso, A. R. Ross, and R. McGrath, *Phys. Rev. Lett.* **92**, 135507 (2004).

²⁰H. Takakura, A. Yamamoto, and A.-P. Tsai, *Acta Crystallogr., Sect. A: Found. Crystallogr.* **57**, 576 (2001).

²¹S. Deloudi, M. Kobas, and W. Steurer, *Philos. Mag. B* (to be published).

²²A. Cervellino, T. Haibach, and W. Steurer, *Acta Crystallogr., Sect. B: Struct. Sci.* **58**, 8 (2002).

²³L. Vitos, A. V. Ruban, H. L. Skriver, and J. Kollar, *Surf. Sci.* **411**, 186 (1998).

²⁴A. Baumgarte, J. Schreuer, M. A. Estermann, and W. Steurer, *Philos. Mag. A* **75**, 1665 (1997).