

6-15-2014

Mixed valency and site-preference chemistry for cerium and its compounds: A predictive density-functional theory study

Aftab Alam

Indian Institute of Technology Bombay

Duane D. Johnson

Iowa State University, ddj@iastate.edu

Follow this and additional works at: http://lib.dr.iastate.edu/ameslab_pubs

 Part of the [Condensed Matter Physics Commons](#), [Engineering Physics Commons](#), and the [Materials Science and Engineering Commons](#)

The complete bibliographic information for this item can be found at http://lib.dr.iastate.edu/ameslab_pubs/271. For information on how to cite this item, please visit <http://lib.dr.iastate.edu/howtocite.html>.

This Article is brought to you for free and open access by the Ames Laboratory at Iowa State University Digital Repository. It has been accepted for inclusion in Ames Laboratory Publications by an authorized administrator of Iowa State University Digital Repository. For more information, please contact digirep@iastate.edu.

Mixed valency and site-preference chemistry for cerium and its compounds: A predictive density-functional theory study

Abstract

Cerium and its technologically relevant compounds are examples of anomalous mixed valency, originating from two competing oxidation states—itinerant Ce⁴⁺ and localized Ce³⁺. Under applied stress, anomalous transitions are observed but not well understood. Here we treat mixed valency as an “alloy” problem involving two valences with competing and numerous site-occupancy configurations. We use density-functional theory with Hubbard U (i.e., DFT+ U) to evaluate the effective valence and predict properties, including controlling the valence by pseudoternary alloying. For Ce and its compounds, such as (Ce,La)₂(Fe,Co)₁₄B permanent magnets, we find a stable mixed-valent α state near the spectroscopic value of $v_s=3.53$. Ce valency in compounds depends on its steric volume and local chemistry. For La doping, Ce valency shifts towards γ -like Ce³⁺, as expected from steric volume; for Co doping, valency depends on local Ce-site chemistry and steric volume. Our approach captures the key origins of anomalous valency and site-preference chemistry in complex compounds.

Keywords

Materials Science and Engineering

Disciplines

Condensed Matter Physics | Engineering Physics | Materials Science and Engineering

Comments

This article is from Phys. Rev. B **89**, 235126 (2014), doi:[10.1103/PhysRevB.89.235126](https://doi.org/10.1103/PhysRevB.89.235126). Posted with permission.

Mixed valency and site-preference chemistry for cerium and its compounds: A predictive density-functional theory study

Aftab Alam^{1,*} and D. D. Johnson^{2,3,†}¹*Department of Physics, Indian Institute of Technology, Bombay, Powai, Mumbai 400 076, India*²*The Ames Laboratory, U.S. Department of Energy, Ames, Iowa 50011-3020, USA*³*Materials Science & Engineering, Iowa State University, Ames, Iowa 50011-2300, USA*

(Received 15 December 2013; revised manuscript received 3 June 2014; published 23 June 2014)

Cerium and its technologically relevant compounds are examples of anomalous mixed valency, originating from two competing oxidation states— itinerant Ce^{4+} and localized Ce^{3+} . Under applied stress, anomalous transitions are observed but not well understood. Here we treat mixed valency as an “alloy” problem involving two valences with competing and numerous site-occupancy configurations. We use density-functional theory with Hubbard U (i.e., DFT+ U) to evaluate the effective valence and predict properties, including controlling the valence by pseudoternary alloying. For Ce and its compounds, such as $(Ce,La)_2(Fe,Co)_{14}B$ permanent magnets, we find a stable mixed-valent α state near the spectroscopic value of $\nu_s = 3.53$. Ce valency in compounds depends on its steric volume and local chemistry. For La doping, Ce valency shifts towards γ -like Ce^{3+} , as expected from steric volume; for Co doping, valency depends on local Ce-site chemistry and steric volume. Our approach captures the key origins of anomalous valency and site-preference chemistry in complex compounds.

DOI: [10.1103/PhysRevB.89.235126](https://doi.org/10.1103/PhysRevB.89.235126)

PACS number(s): 71.28.+d, 71.15.Ap, 71.20.Eh, 75.30.Mb

Mixed-valence compounds exhibit interesting anomalies when external parameters, such as pressure, are varied. Reliably predicting their properties and determining the origin of mixed-valence effects remains open and dependent upon correlating the properties with the molecular and crystal structure, and with local chemistry. Apart from the field of metallurgy and pigments in artwork and ceramics (e.g., Fe_2O_3 and Fe_3O_4) [1], it also serves as an active area of research in complex biophysical problems, such as photosynthesis, and in organic-conjugated materials [2] for artificial electronic devices [3]. What makes these compounds different from other materials is the coexistence of wide s - d bands and very heavy atomlike f electrons at/near the Fermi energy. Anomalous properties then arise from a competition between the itinerant and localized nature of the f electrons. Numerous experiments have revealed a growing list of such compounds from the rare-earth (RE) series to late actinides, and some transition-metal compounds.

Cerium is the first RE element that exhibits phases with enormous (15–17%) volume differences [4]. As the most abundant RE, Ce is often considered as a RE replacement in permanent magnets [5]. The low-pressure γ phase of Ce exhibits a local magnetic moment, associated with a trivalent Ce^{3+} configuration. With applied pressure, Ce first transforms into a mixed-valent α state with quenched moments, and eventually transforms into a tetravalent Ce^{4+} α' state at higher pressure. With alloying, a local “chemical” pressure can be exerted, so the Ce valency in its compounds can be more sensitive, e.g., $Ce_2Fe_{14}B$, where it remains in a strongly mixed-valent α -like state, which is incompatible with a local $4f$ moment. The central problem is to understand the mechanism that controls Ce valency both in pure Ce and its compounds.

Here, we treat mixed valency by mapping to an “alloy” problem, where the Ce^{3+} and Ce^{4+} are considered as two different atoms, and we include the binomial distribution of RE-site occupancy in the lattice. Our approach captures the key electronic and chemical effects, reproducing the observed mixed valency of Ce and its complex compounds, such as $(Ce-La)_2(Fe-Co)_{14}B$ magnets. We show that the Ce valency in compounds depends on the steric volume of Ce sites and local chemistry surrounding the RE site. Mixed valency of Ce is then predicted similarly to studies of rare-earth systems using model Hamiltonians [6], and is consistent with a correlated and multielectron picture of Ce with semi-isolated $4f$ states in contact with a bath of spd valence electrons, as found experimentally [7].

Addressing mixed valency using a density-functional theory (DFT) treats magnetism, atomic multiplet effects, and crystal-field splitting on an equal footing, and identifies the electronic mechanisms responsible for the anomalous valence behavior. While a first-principles dynamical mean-field theory (DMFT) may better describe the fluctuating mixed valency, our approach captures the key effects in complex compounds with dramatically less computational intensity. Notably, within DMFT, δ Pu is found to be a superposition of two atomic valences (60% f^5 and 40% f^6) [8]. Yet, experimentally, α and δ Pu have a superposition of three $5f$ states [9] ($\sim 20\%$ f^4 , 40% f^5 , and 40% f^6), i.e., a ternary alloy (two independent fractions).

We use the Vienna *ab initio* simulation package VASP [10] with a pseudopotential and projected augmented-wave basis [11] using Perdew-Burke-Ernzhorf (PBE) exchange-correlation and spin-orbit coupling. With different sized Ce^{3+} and Ce^{4+} ions, relaxations (ignored in previous studies) are crucial to predict reliable energetics and ground states. Localized Ce^{3+} f electrons are addressed via a PBE+ U approach [12] with a Hubbard U (set to 5 eV from previous work [5]) introduced in a screened Hartree-Fock manner. All of the other components (La, Ce^{4+} , Fe, Co, and C) are addressed within PBE only, as the effect of Hubbard U on

*aftab@phy.iitb.ac.in

†ddj@ameslab.gov

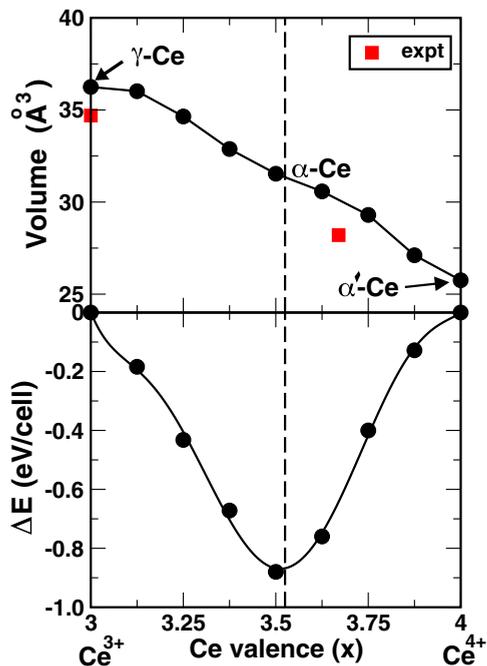


FIG. 1. (Color online) Top: Relaxed volume. Bottom: Formation enthalpy vs Ce valence for pure fcc Ce where trivalent (3+) and tetravalent (4+) Ce are mixed. Square symbols indicate experimental volume (from Ref. [14]). The vertical dashed line is a guide to the eye for the estimated minimum.

them is very small. All potentials are obtained via standard fully self-consistent DFT-based solutions. See [13] for more computational details.

Formation enthalpy (ΔE) is defined as the relative energy of the alloy [(Ce³⁺, Ce⁴⁺) in the present case] with respect to the concentration-weighted sum of the energies of the α -constituent elements in the compound, i.e.,

$$\Delta E = E^{\text{Alloy}}(V^0) - \sum_{\alpha=1}^s c_{\alpha} E^{\alpha}(V_{\alpha}^0), \quad (1)$$

where c_{α} is the concentration of the α constituents. V^0 and V_{α}^0 are the equilibrium volume of the alloy and the constituents in their ground states, respectively.

The formation enthalpy ΔE and volume V versus Ce valence in fcc Ce mixed-valence alloys are shown in Fig. 1. Here, Ce³⁺ and Ce⁴⁺ potentials are occupied over eight sites in a fcc supercell to find the energetically most favorable configuration; hence, we have discrete jumps of 0.125 valence “composition.” From Fig. 1, the energetically most favorable mixed-valence state occurs near 3.5 (a fifth-order polynomial fit yields $v_s = 3.55$), close to the assessed value of 3.67 [4]. We find that the mixed-valence α state of Ce arises from an energetically favorable distribution of Ce³⁺ and Ce⁴⁺ states. (Using more sites fills in the curve, which is needed in skewed distributions.) Atomic positions and the cell volumes are fully relaxed in each data. The relaxed, DFT+ U calculated V 's are compared with known experimental volumes [14] (red squares), and are within 8% and have the correct trend for the γ , α , and α' phases. For pure Ce, the mixed valency corresponds to a volume between those of the purely γ and α phases.

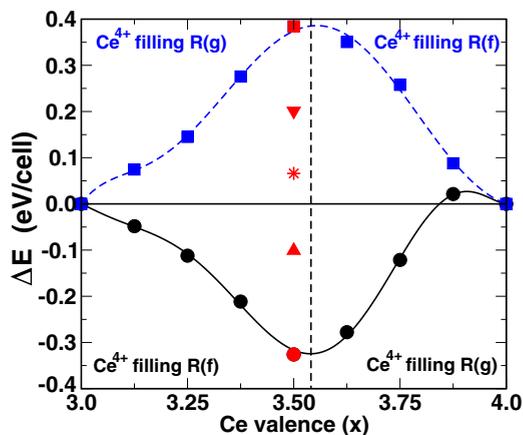


FIG. 2. (Color online) Mixed valency of Ce in Ce₂Fe₁₄B. Solid (dashed) curve indicates the formation enthalpies of mixing Ce³⁺ and Ce⁴⁺, with the larger 3+ ions distributed over 4g (4f) sites and smaller 4+ ions distributed over 4f (4g) sites. Star and triangles indicate the energies for intermediate sets of distribution where 3+ and 4+ ions are mixed over both of the sites.

The inclusion of on-site U for Ce³⁺ (with 4f electron) and the spin-orbit coupling is important to get the correct ground state for the intermediate valency.

Next, we investigate the Ce mixed valency in Ce₂Fe₁₄B, a challenging 68-atom-per-cell, tetragonal structure with space group $P4_2/mnm$ [15]. The mixed valency of Ce significantly affects its magnetic behavior. Due to the complex nature of a 2-14-1 structure [5], the Ce mixed valency is associated with Ce-site preferences. The 2-14-1 structure contains two inequivalent rare-earth (R) sites— $R(4f)$ and $R(4g)$ —each with multiplicity 4 [5]. From the coordination shell around each site, 4g sites acquire a larger volume than 4f sites. As such, Ce³⁺ (larger ion) prefers to occupy the 4g sites, while Ce⁴⁺ (smaller ion) prefers to occupy the 4f sites. Energetically favored configurations are found by mixing Ce³⁺ and Ce⁴⁺ on the eight R sites in all possible ways.

The formation enthalpy ΔE versus Ce valence in (Ce³⁺-Ce⁴⁺)₂Fe₁₄B for all eight R -site configurations is shown in Fig. 2. The filled circles indicate energies when Ce³⁺ is favorably distributed over 4g sites and Ce⁴⁺ on 4f sites; filled squares are the results with the opposite (unfavorable) distribution of Ce ions. Other symbols indicate intermediate sets of distribution where 3+ and 4+ ions are mixed over both of the sites with a binomial distribution. Notice the asymmetric nature of the energy curves comparing the lower vs upper half of ΔE , which is due to a different filling on the two inequivalent RE sites; that is, the collective effect of filling Ce³⁺ ions preferentially over 4f sites is very different from that of 4g sites. The favorable mixed valency occurs near $v_s = 3.5$ (3.53 from polynomial fit), close to the assessed 3.44 [16].

To improve the magnetic properties requires engineering Ce or Fe sites in such a way as to push the Ce valency either towards 3+ or 4+. Due to the dependence of the Ce valence on the steric volume [5], one way to manipulate the Ce valency is to vary the unit cell volume by forming pseudoternary compounds. We studied two compounds, i.e., (Ce,La)₂Fe₁₄B and Ce₂(Fe,Co)₁₄B; the former (latter) should

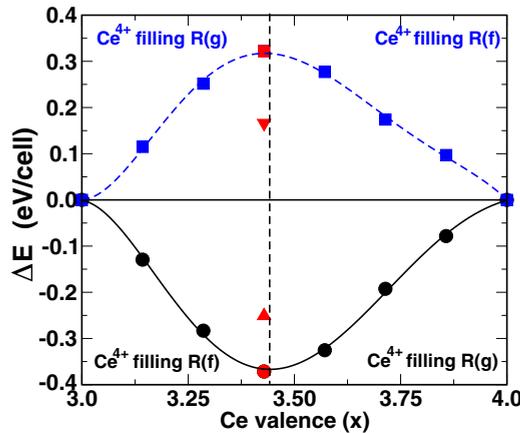


FIG. 3. (Color online) Similar to Fig. 2 but for $(\text{Ce}_{87.5}\text{La}_{12.5})_2\text{Fe}_{14}\text{B}$. Here La is substituted at its favored $4g$ site and Ce^{3+} and Ce^{4+} are distributed over the remaining seven RE sites in all possible ways.

increase (reduce) the unit cell volume. First, we dope RE Ce sites by La in $\text{Ce}_2\text{Fe}_{14}\text{B}$. Our calculation shows that out of the two inequivalent RE sites $4f$ and $4g$, La (being larger than Ce) energetically prefers to occupy the $4g$ sites. This was also observed from x-ray absorption near-edge spectroscopy (XANES) measurement [17].

Figure 3 shows the ΔE vs Ce valence with 12.5% of La doping (one out of eight sites) in $\text{Ce}_2\text{Fe}_{14}\text{B}$. From the data in the vicinity of the minimum ($x \sim 3.5$), La doping clearly moves the Ce valency towards $3+$ relative to the undoped case. This effect is in accord with the steric volume argument: a La ion, being larger than Ce, expands the lattice when doped in $\text{Ce}_2\text{Fe}_{14}\text{B}$, enhancing the steric volume of Ce site(s), and supports a more trivalentlike state. The energetically most favorable mixed-valence state occurs near ≤ 3.5 ($\nu_s \sim 3.43$ from polynomial fit), which compares well with the observed value of $\nu_s \sim 3.46$ from the XANES experiment [17] for $(\text{Ce}_{90}\text{La}_{10})_2\text{Fe}_{14}\text{B}$. Steric volume is an important factor controlling the Ce chemical valence, as also evidenced in hydrogenated $\text{Ce}_2\text{Fe}_{14}\text{B}$ and $\text{Ce}_2\text{Fe}_{17}$ compounds [16,18].

Next, we study the effects of Co doping on Fe sites in $\text{Ce}_2\text{Fe}_{14}\text{B}$, which is known to enhance the Curie temperature and magnetic anisotropy and, hence, is a reason for our choice. $\text{Ce}_2(\text{Fe},\text{Co})_{14}\text{B}$ crystallizes in the same $P4_2/mnm$ structure as $\text{Ce}_2\text{Fe}_{14}\text{B}$, which has six inequivalent Fe sites [5], i.e., $\text{Fe}(k1)$, $\text{Fe}(k2)$, $\text{Fe}(j1)$, $\text{Fe}(j2)$, $\text{Fe}(e)$, and $\text{Fe}(c)$. First we verified the Co-site preference on these symmetry distinct Fe sites. Figure 4 (top) shows the site-preference energy for both Ce^{3+} and Ce^{4+} when one out of 14 Fe sites is doped with a Co atom. The results indicate that the $\text{Fe}(j2)$ site has the strongest preference for not occupying Co, as it costs the highest energy.

This finding is supported by two arguments: (1) Co and Fe differ a little in size ($R_{\text{Co}} < R_{\text{Fe}}$). Out of six Fe sites in 2-14-1, $\text{Fe}(j2)$ has the largest coordination volume—a reason for Co to avoid $j2$ sites. (2) $j2$ sites in 2-14-1 [and c (dumbbell) sites in a rhombohedral $\text{Ce}_2\text{Fe}_{17}$ structure [15]] are reported from neutron-diffraction measurement [19] to be crystallographically, as well as magnetically, cognates. In other words, each of these sites has the largest number of near-neighbor Fe ions and the largest moment. Also, these

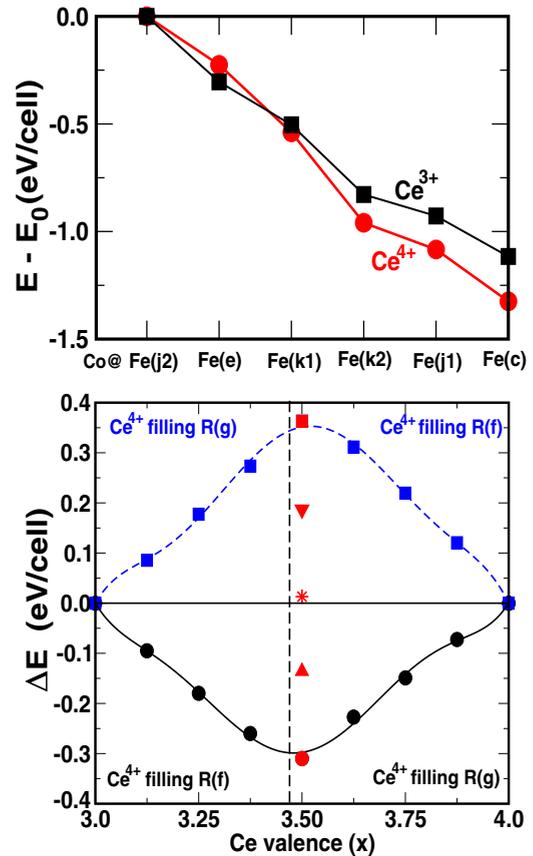


FIG. 4. (Color online) Top: Site-preference energy for Co doping at various Fe sites in $\text{Ce}_2\text{Fe}_{14}\text{B}$. E_0 is a reference when Co is on the $\text{Fe}(j2)$ site. Bottom: Same as Fig. 2, but for $\text{Ce}_2(\text{Fe}_{13}\text{Co})\text{B}$ with Co occupying $\text{Fe}(c)$ sites.

sites are the only transition-metal sites that have a major ligand line—perhaps another reason behind the unfavorability of Co to occupy $j2$ sites. $\text{Fe}(c)$ and $\text{Fe}(j1)$ sites have the strongest preference towards Co. This site preference can be justified from the large affinity of Co towards rare earth, i.e., those transition-metal sites that acquire the highest RE coordination will prefer to have Co on them. $\text{Fe}(c)$ and $\text{Fe}(j1)$, indeed, have the highest coordination of RE around them.

Figure 4 shows ΔE vs Ce valence for $\text{Ce}_2(\text{Fe}_{13}\text{Co})\text{B}$ with Co doped on the energetically most favorable $\text{Fe}(c)$ site. Unlike $\text{Ce}_2\text{Fe}_{14}\text{B}$ (Fig. 2), 7.14% Co doping (1 out of 14) already favors the mixing of the Ce valence at the Ce^{4+} end, i.e., no positive (unfavorable) ΔE . Again, compared to the Ce valence in $\text{Ce}_2\text{Fe}_{14}\text{B}$, Co doping pushes the valency of Ce towards $3+$ (a similar argument for the polynomial fit to locate minima holds here). This, however, does not jibe with the volume argument used for La doping. Co being smaller than Fe leads to a volume reduction that should move the Co valence towards $4+$ via steric volume, instead of $3+$. In this case, the local chemistry and the associated local steric volume of RE sites play an important role in determining the Ce valency compared to the simple concept based on global volume reduction.

It is well known that Co has a strong affinity to RE elements (Ce in this case), and as such Ce favors a high coordination number of Co. In 2-14-1, $\text{Fe}(c)$ sites with 4 Ce and $\text{Fe}(j1)$ sites with 3 Ce have the highest number of RE neighbors.

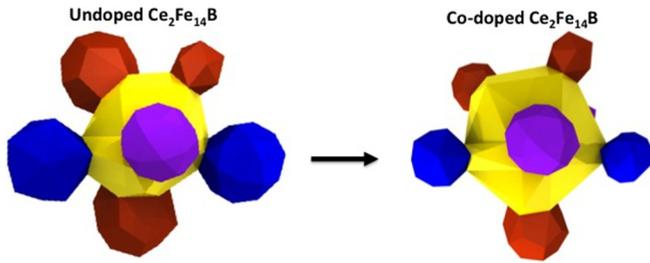


FIG. 5. (Color online) Expansion of the steric (Voronoi) volume of the central Ce site (yellow) after Co doping. Inequivalent Fe sites are denoted by blue, red, and purple. With Co doping, the steric volume of the Ce site is enlarged, at the expense of the Fe sites and associated charge.

These Fe sites are, indeed, the energetically most favorable sites for Co; see Fig. 4 (top). Now, because $R_{\text{Co}} < R_{\text{Fe}}$, the accumulation of a large number of Co around the Ce site causes the formation of major ligands, given by lines connecting faces of Voronoi polyhedra, and allowing more room and an expansion of the local Voronoi volume around the Ce site. (These Voronoi polyhedra and volumes were determined by inscribed radii given by saddle points in the electronic density [20].) Note how the central Ce polyhedra expands (Fig. 5) due to Co doping on Fe sites. Thus, although Co doping reduces the unit cell volume, the local steric volume around the Ce site is enhanced, which shifts the Ce valency towards $3+$. This phenomenon is based on the local chemistry and the nature of hybridization of the Ce ion with its neighboring atoms, instead of the simple volume argument alone. Another argument can be based on the increase in band filling due to one additional electron (arising from Co doping). Now, because of the location of Ce- $4f$ bands in the vicinity of

the Fermi energy, this can also cause a decrease in the Ce valence.

In summary, we have presented an alloy approach to reliably predict the mixed-valency properties in complex compounds, for which DFT+ U methods are essential, and reveal the electronic origin for such behavior. For Ce-based materials, cerium does not have a well-defined valence; rather, its f electrons fluctuate between two extreme valence states ($3+$ and $4+$) depending upon local atomic configurations (site occupancy). The energy difference between these states/configurations is a few meV atom $^{-1}$, so associated anomalies are observed, e.g., under pressure. The mechanism for such a transition and the reason for differing valence states is not yet well understood. Doping puts the material under chemical pressure. In $\text{Ce}_2\text{Fe}_{14}\text{B}$, La doping at Ce sites expands the lattice (as expected from steric volume arguments), while a transition-metal dopant such as Co at Fe sites shrinks it. Here we predicted the mixed valency of Ce in pure Ce and $\text{Ce}_2\text{Fe}_{14}\text{B}$, in agreement with experiment. Then, we addressed two different types of doping (La at Ce sites and Co at Fe sites) to reveal how both steric volume and local chemistry influence the Ce valence in compounds. In Co-doped $\text{Ce}_2\text{Fe}_{14}\text{B}$, a simple argument based on steric volume does not hold, rather the nature of hybridization and the electron affinity of Co towards RE elements are crucial to predict the mixed valency of Ce.

Work at Ames Laboratory was supported by the US Department of Energy (DOE) ARPA-E REACT Program (Contract No. 0472-1526), using capabilities maintained and supported by the Office of Basic Energy Sciences in our Division of Materials Science and Engineering. The Ames Laboratory is operated for the US DOE by Iowa State University under Contract No. DE-AC02-07CH11358.

-
- [1] Peter Day, Noel S. Hush, and Robin J. H. Clark, *Philos. Trans. A* **366**, 5 (2008).
- [2] A. Heckmann and C. Lambert, *Angew. Chem. Int. Ed.* **51**, 326 (2012).
- [3] Oliver S. Wenger, *Chem. Soc. Rev.* **41**, 3772 (2012).
- [4] D. C. Koskenmaki and K. A. Gschneidner, Jr., in *Handbook on the Physics and Chemistry of Rare Earths*, edited by K. A. Gschneidner, Jr. and L. Eyring (North-Holland, Amsterdam, 1978), p. 337.
- [5] Aftab Alam, Mahmud Khan, R. W. McCallum, and D. D. Johnson, *Appl. Phys. Lett.* **102**, 042402 (2013).
- [6] S. K. Ghatak, M. Avignon, and K. H. Bennemann, *J. Phys. F: Metal Phys. B* **6**, 1441 (1976).
- [7] J. G. Tobin, S. W. Yu, T. Komesu, B. W. Chung, S. A. Morton, and G. D. Waddill, *Europhys. Lett.* **77**, 17004 (2007).
- [8] H. J. Shim, K. Haule, and G. Kotliar, *Nature (London)* **446**, 513 (2007), and references therein.
- [9] C. H. Booth, Yu Jianga, D. L. Wang, J. N. Mitchell, P. H. Tobash, E. D. Bauer, M. A. Walle, P. G. Allen, D. Sokaras, D. Nordlund, T.-C. Weng, M. A. Torrez, and J. L. Sarrao, *PNAS* **109**, 10205 (2012).
- [10] G. Kresse and J. Furthmüller, *Phys. Rev. B* **54**, 11169 (1996); *Comput. Mater. Sci.* **6**, 15 (1996).
- [11] G. Kresse and D. Joubert, *Phys. Rev. B* **59**, 1758 (1999).
- [12] A. I. Liechtenstein, V. I. Anisimov, and J. Zaanen, *Phys. Rev. B* **52**, R5467 (1995).
- [13] A Monkhorst-pack Brillouin-zone integration with $24 \times 24 \times 12$ ($4 \times 4 \times 4$) \mathbf{k} mesh is used for 8-atom fcc (68-atom 2-14-1 compound). We use *high* precision with large plane-wave cutoffs (400–450 eV), giving a convergence within 4 meV/cell (8 kBar) for energy (stress tensor). More details are available in Ref. [5], including selection of U .
- [14] B. Johansson, I. A. Abrikosov, M. Aldén, A. V. Ruban, and H. L. Skriver, *Phys. Rev. Lett.* **74**, 2335 (1995).
- [15] J. F. Herbst, *Rev. Mod. Phys.* **63**, 819 (1991), and references therein.
- [16] T. W. Capehart, R. K. Mishra, G. P. Meisner, C. D. Fuerst, and J. F. Herbst, *Appl. Phys. Lett.* **63**, 3642 (1993).
- [17] T. W. Capehart, R. K. Mishra, C. D. Fuerst, G. P. Meisner, F. E. Pinkerton, and J. F. Herbst, *Phys. Rev. B* **55**, 11496 (1997).
- [18] D. Fruchart, F. Vaillant, A. Yaouanc, J. M. D. Coey, R. Fruchart, Ph. L'Heritier, T. Riesterer, J. Osterwalder, and L. Schlappbach, *J. Less-Common Met.* **130**, 97 (1987).
- [19] J. F. Herbst and W. B. Yelon, *J. Appl. Phys.* **60**, 4224 (1986).
- [20] Aftab Alam, S. N. Khan, Brian G. Wilson, and D. D. Johnson, *Phys. Rev. B* **84**, 045105 (2011).