Design, synthesis, characterization and study of novel conjugated polymers

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Design, synthesis, characterization and study of novel conjugated polymers

by

Wu Chen

A dissertation submitted to the graduate faculty
in partial fulfillment of the requirements for the degree of

DOCTOR OF PHILOSOPHY

Major: Organic Chemistry

Major Professor: Thomas J. Barton

Iowa State University

Ames, Iowa

1997
This is to certify that the doctoral dissertation of
Wu Chen
has met the dissertation requirements of Iowa State University

Signature was redacted for privacy.

Major Professor
Signature was redacted for privacy.

For the Major Program
Signature was redacted for privacy.

For the Graduate College
DEDICATION

TO MY PARENTS,
TO MY WIFE, HONGLING LI
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INTRODUCTION

Conjugated organic polymers are polymers which contain overlapping p-orbitals throughout the polymer main chains. Three types of p-orbital units used to construct conjugated polymers are aromatic segments like phenyl, thiophene and other hetero-aromatic rings, non-aromatic C-C double bonds or triple bonds or C-N double bonds, and hetero atoms having nonbonding electron pairs like sulfur or nitrogen. The different combinations of these units can form a wide variety of conjugated polymers. The conjugated polymers which have been most studied are polyacetylene (PA), polydiacetylene (PDA), polyparaphenylene (PPP), polythiophene (PT), polyphenylenesulfide (PPS), polyparaphenylenevinylene (PPV), polyparaphenylenethynylene (PPE), polyaniline (PANI), polyphenylenesulfide (PPS) and polyacenes. The structures of these polymers are shown in Figure 1.

Conjugated polymers have been attracting attention for more than two decades because some can exhibit high electroconductivity (after doping), photoconductivity, strong non-linear optical response or intense fluorescence. Conjugated polymers were first studied as conducting polymers. They are often capable of being doped, through chemical or electrochemical oxidation or reduction, to states of moderate to high electrical conductivity. Twenty years ago, the idea of a conducting polymer was purely a subject of theoretical debate; however, a fortunate coincidence of events in the late 1970s led to the first reports of polymeric materials with semiconductor properties. Since then, electrically conducting polymers have been a center of scientific interest and active multidisciplinary research. These materials have attracted academic and industrial research groups not only because of their theoretically interesting properties but also because of their technologically promising future.

A number of applications have been proposed for conducting polymers. They include field-effect transistors, light weight rechargeable batteries, electrochromic displays, erasable compact discs and photovoltaic devices. During the last twenty years, many attempts have been made to produce polymers with a high level of electrical conduction and much of this effort has been directed towards polyacetylenes. Oxidatively doped and stretch oriented
Figure 1: The structures of conjugated polymers
Table 1: Conductivities of conducting polymers

<table>
<thead>
<tr>
<th>Polymer</th>
<th>Doping materials</th>
<th>Conductivities (Scm&lt;sup&gt;-1&lt;/sup&gt;)</th>
</tr>
</thead>
<tbody>
<tr>
<td>Polyacetylene</td>
<td>I&lt;sub&gt;2&lt;/sub&gt;, Br&lt;sub&gt;2&lt;/sub&gt;, Li, Na, AsF&lt;sub&gt;5&lt;/sub&gt;</td>
<td>10,000</td>
</tr>
<tr>
<td>Polypyrrole</td>
<td>BF&lt;sub&gt;4&lt;/sub&gt;&lt;sup&gt;-&lt;/sup&gt;, ClO&lt;sub&gt;4&lt;/sub&gt;&lt;sup&gt;-&lt;/sup&gt;</td>
<td>500-7500</td>
</tr>
<tr>
<td>Poly(3-alkylthiophene)</td>
<td>BF&lt;sub&gt;4&lt;/sub&gt;&lt;sup&gt;-&lt;/sup&gt;, ClO&lt;sub&gt;4&lt;/sub&gt;&lt;sup&gt;-&lt;/sup&gt;, FeCl&lt;sub&gt;4&lt;/sub&gt;&lt;sup&gt;-&lt;/sup&gt;</td>
<td>1000-10,000</td>
</tr>
<tr>
<td>Polyphenylenevinylene</td>
<td>AsF&lt;sub&gt;5&lt;/sub&gt;</td>
<td>10,000</td>
</tr>
<tr>
<td>Polyphylene sulfide</td>
<td>AsF&lt;sub&gt;5&lt;/sub&gt;</td>
<td>500</td>
</tr>
<tr>
<td>Polyphenylene</td>
<td>AsF&lt;sub&gt;5&lt;/sub&gt;, Li, K</td>
<td>1000</td>
</tr>
<tr>
<td>Polyaniline</td>
<td>HCl</td>
<td>200</td>
</tr>
</tbody>
</table>

trans-PA is the most conductive (σ=10<sup>3</sup>-10<sup>5</sup> S/cm)<sup>11</sup> among conducting polymers. Table 1 gives the examples of the conductivities of these polymers.

Conjugated polymers can also exhibit photoconductive properties.<sup>47</sup> In the absence of light most polymers are fairly insulators. The lowest excited state in conjugated polymers is a band-to-band transition leading to the formation of free electron-hole pairs.<sup>48</sup> In this picture photoluminescence is interpreted in terms of radiative recombination of charge carriers after relaxing into polaronic states of the polymeric backbone. After being exposed to light (usually laser pulse), the conjugated polymers can generate electron-hole pairs. These pairs must dissociate to produce separated electrons and positive holes in such way that the electrons end up in the conducting states. Poly(2-phenyl-1,4-phenylenevinylene) (PPPV) has been studied as photoconductive polymer.<sup>47(a)</sup>

Even though only doped conjugated polymers are conducting polymers, presently the underlying chemistry and physics of these materials in their undoped states may be more interesting. These undoped polymers have photoluminescent and electroluminescent properties. Although electroluminescence in organic materials has been known since the
1960s, the first observation of electroluminescence from polymeric devices was reported by physicists and chemists from the University of Cambridge in 1990. This organic light-emitting device (LED) was constructed from undoped PPV sandwiched between indium tin oxide (ITO) and aluminum electrodes. Figure 2 shows the construction of an LED.

![Figure 2: Schematic of an LED device](image)

There are at least four steps involved in the operating mechanism of these devices: charge injection, charge transport, charge recombination and light emission. When a voltage is applied across the polymer film, the cathode injects electrons (in the form of radical anions) into the polymer film, while the anode injects holes (as radical cations). The electrical field makes these electrons and holes migrate toward each other through the film. When they meet in the middle, the energetic electrons can drop into the holes to form triplet or singlet excitons, which give up their excess energy as photons of light (Figure 3(a)).

In most conjugated polymers including PPV, the mobility of holes is faster than that of electrons. The differences in mobilities cause unbalanced charges in the polymer film and the charge recombination occurs close to or inside the negative electrode, which results in diminishing the electroluminescence (EL) quantum efficiency by the self-absorption. In order to achieve a high recombination efficiency, various types of cell structures have been proposed. One of the most common is a double-layer-type cell that has a layered structure composed of materials with different carrier transport properties, one transporting holes and the other electrons (Figure 3(b)). In this system, electrons and holes are injected from electrodes into the corresponding organic (carrier transporting) layers. Eventually, the recombination of the carriers takes place at or near the interface between the two organic
layers. Since the hole transport layer blocks electrons, and the electron transport layer blocks holes, the carriers are confined to the organic layer, thus maximizing the recombination efficiency. An improvement of EL quantum efficiency was also observed when PPV as replaced by a partially conjugated PPV.\textsuperscript{24(c)}

![Diagram of LED configuration](image)

**Figure 3:** Configuration of typical LED

Since the first report of a PPV-based LED, there has been a rapidly increasing number of publications demonstrating the applications of these conjugated polymers in LED devices.\textsuperscript{12} Other conjugated polymers used for EL application include PPE,\textsuperscript{25b,c} polyanalkylthiophenes (PATs),\textsuperscript{13} polyalkylfluorenes (PAFs),\textsuperscript{14} PPP\textsuperscript{15} and poly(1,4-naphthalene vinylene) (PNV).\textsuperscript{16}

Polymeric LEDs are likely to replace the inorganic LEDs in the future because of their advantages. The polymer-based light-emitting materials are light, inexpensive, flexible and efficient. They can be used for everything from lightweight backlights for computer displays to TVs that can be hanged flat on the wall or rolled up.

Conjugated polymers are emerging as an important class of third-order nonlinear optical (NLO) materials.\textsuperscript{17} NLO materials in devices allow light (photons) to be used instead of electrons for the representation, manipulation and transmission of information. Devices such as optical switches and modulators will perform the same functions as their electronic counterparts but with increased speed. As mentioned before, the devices are light, flexible and inexpensive.
The equations describing the polarization of a material as a function of applied field on both molecular level and the bulk level are shown in Figure 4.

Conjugated polymers such as polyacetylene, poly(phenylene vinylene) (PPV), polythiophene, polydiacetylene and some other polymers have been shown to exhibit very large nonlinear optical response (Table 2). 

<table>
<thead>
<tr>
<th>Molecular Level</th>
<th>Bulk Level</th>
</tr>
</thead>
<tbody>
<tr>
<td>$p = \alpha E + \beta E \cdot E + \gamma E \cdot E \cdot E + \ldots$</td>
<td>$p = \chi^{(1)} E + \chi^{(2)} E \cdot E + \chi^{(3)} E \cdot E \cdot E + \ldots$</td>
</tr>
<tr>
<td>$p =$ polarization of individual molecule</td>
<td>$p =$ polarization of bulk material</td>
</tr>
<tr>
<td>$\alpha =$ molecular polarizability</td>
<td>$\chi^{(1)} =$ linear susceptibility tensor</td>
</tr>
<tr>
<td>$\beta =$ 1st hyperpolarizability</td>
<td>$\chi^{(2)} =$ 1st nonlinear susceptibility tensor</td>
</tr>
<tr>
<td>$\gamma =$ 2nd hyperpolarizability</td>
<td>$\chi^{(3)} =$ 2nd nonlinear susceptibility tensor</td>
</tr>
</tbody>
</table>

Figure 4: Polarization of a medium induced by an external electric field
Table 2: Selected polymers that exhibit third-order nonlinearities

<table>
<thead>
<tr>
<th>Polymers</th>
<th>Nonresonant $\chi^{(3)}$ (esu) [wavelength]</th>
<th>Resonant $\chi^{(3)}$ (esu) [wavelength]</th>
<th>Measurement method</th>
</tr>
</thead>
<tbody>
<tr>
<td>Polyacetylene</td>
<td>$1.0 \times 10^{-10}$ [950nm]</td>
<td>$1.1 \times 10^{9}$ [620nm]</td>
<td>THG * spectroscopy</td>
</tr>
<tr>
<td>Polydiacetylene</td>
<td>$4.0 \times 10^{-10}$ [602nm]</td>
<td></td>
<td>DFWM **</td>
</tr>
<tr>
<td>Polythiophene</td>
<td></td>
<td>$4.0 \times 10^{-10}$ [602nm]</td>
<td>DFWM **</td>
</tr>
<tr>
<td>Poly(phenylbenzobisthiazole)</td>
<td>$4.5 \times 10^{-10}$ [602nm]</td>
<td></td>
<td>DFWM **</td>
</tr>
<tr>
<td>Polyquinoline</td>
<td>$2.3 \times 10^{-12}$ [2.38μm]</td>
<td>$1.1 \times 10^{-11}$ [1.2μm]</td>
<td>THG * spectroscopy</td>
</tr>
<tr>
<td>Polyphenylenevinylene</td>
<td></td>
<td>$4.0 \times 10^{-10}$ [602nm]</td>
<td>DFWM **</td>
</tr>
<tr>
<td>Polyaniline (emeraldine base)</td>
<td></td>
<td>$3.7 \times 10^{-11}$ [1.83μm]</td>
<td>THG * spectroscopy</td>
</tr>
</tbody>
</table>

* THG, third harmonic generation; ** DFWM, degenerate four-wave mixing
I: SYNTHESES OF ELECTROLUMINESCENT POLYMERS CONTAINING
CONJUGATED ARYL, OLEFINIC, THIOPHENE AND ACETYLENIC UNITS AND
THEIR STUDIES FOR USE IN LIGHT-EMITTING DIODES (LED)

Literature Survey

Conjugated polymers today are most widely used in emitters. The first conjugated
dpolymer applied to an electroluminescent (EL) device was PPV. Since then, PPV-based
LEDs have been extensively studied. PPVs usually have two identical, long side chains to
increase their solubilities and processabilities (Figure 5). These polymers, used as hole-
transporting layers, have low quantum efficiencies. To improve their quantum efficiencies,
several adjustments have been made. The unsymmetric MEH-PPV was employed in a LED
device with a lower work function (ionization potential) metal, such as calcium, as the
cathode. The quantum efficiency can be high as 1%. Recently, electron-accepting cyano (-
CN) groups have been introduced into PPV derivatives so that the polymer (CN-PPV) can
transport electrons and have a lower HOMO-LUMO gap energy. A cell with double layers
of PPV (as the hole transport layer) and CN-PPV (as electron transport layer) exhibits a
higher quantum efficiency of 4%- one of the highest efficiencies ever reported.

In contrast to the PPV, poly(p-phenyleneethynylene) (PPE) type polymers having
structures similar to those of the PPV type polymers have received much less attention due to
synthetic inaccessibility of processable PPE polymers. Principal interest in much of this
work has focused on the liquid crystalline properties of PPEs due to their rigid-rod structures.
However, due to their highly fluorescent nature, PPEs and their similar structures have been
examined as potential materials in LED devices recently. Electron delocalizations through
a triple bond in phenyleneethynylene (PE) and a double bond in phenylenevinylene (PV) have
been compared theoretically and experimentally. Some evidence indicates that the double
bond allows better electron delocalization than the triple bond. The direct observation is that
the wavelength of electronic absorption maxima ($\lambda_{\text{max}}$) of the PE compounds are shorter than
those of the PV analogues. Figure 6 shows the examples of $\lambda_{\text{max}}$ of UV/Vis absorption.
Figure 5: Molecular structures of PPV

Figure 6: Electronic absorptions of PE, PV, PPE and PPV
Photoluminescent properties of PE derivatives and PV derivatives have been studied.\(^{28}\) The reported quantum yields of photoluminescence of many of the PE derivatives are comparable to those of the PV analogues. The quantum yield of alkoxy PPE in toluene was reported to be as high as 50%,\(^{29}\) which is about two times higher than that of PPV. Despite being highly fluorescent, studies of electroluminescence of the PPE system have not shown satisfactory results compared to the PPV system.

The syntheses of PPE and its derivatives were recently reviewed.\(^{25(a, b, d)}\) Recently, our group\(^{25(b, c)}\) and another group\(^{25(d)}\) synthesized soluble PPE with high molecular weight using the Heck-type coupling reaction (Scheme 1) and studied their use in LED. However the yellow-greenish light from the PPE-based LED's was of low intensity and the device was short-lived.

\[
\begin{align*}
\begin{array}{c}
\text{I} \quad \text{OR} \\
\text{RO} \\
\text{I} \quad \text{OR'} \\
\text{R'O} \\
\hline \\
\text{R, R' = C}_4\text{H}_9, \text{C}_6\text{H}_{13}, \text{C}_8\text{H}_{17}, \text{C}_{12}\text{H}_{25} \\
\end{array}
\end{align*}
\]

Scheme 1: Synthesis of soluble PPE

By similar Heck-type coupling, other PPE type polymers have been synthesized and studied (Figure 7).
Figure 7: Other PPE type polymers
In our continuing effort to find more efficient PPE type polymers for LED application, we have synthesized a series of conjugated polymers and copolymers with combination of repeating units of PPV, PPE, and PPP and have begun to examine the effects of electronegative substituents on their optical properties.

Results and Discussions

Preparation and Characterization of Polymers

Monomer Preparation. The key monomers with which to synthesize PPE-type polymers via Heck-type coupling are diethynyl-arylenes and arylene dihalides. The diethynylarylene unit used in all the polymers is 1,4-diethynyl-2,5-didodecyloxybenzene 1, which was synthesized as described in Scheme 2.\(^{25(b)}\)

![Scheme 2: Synthesis of monomer 1](#)

Hydroquinone was reacted with 1-bromo-dodecane and CH\(_3\)ONa in CH\(_3\)OH, giving 1,4-didodecyloxy-benzene 2 in 86% yield. Iodination of 2 by excess ZnCl\(_2\) and benzyltrimethylammonium dichloroiodate (BTMAICl\(_2\)) in acetic acid at 50-60\(^\circ\)C for 1 day gave 1,4-didodecyloxy-2,5-diiodobenzene 10 in 77% yield. The white diiiodo product 10
could be easily purified by recrystallization of the crude product in isopropanol or precipitation by adding the saturated crude product in THF into methanol. The palladium-coupling reaction between 10 and trimethylsilylacetylene in benzene/triethylamine, followed by desilylation, gave monomer 1. Monomer 1 is a greenish crystalline material which is stable in the air.

A PPV unit was put into the other monomer in order to form PPE and PPV copolymers, which may exhibit properties of both PPV and PPE. The double bond in the monomers was formed via a Wittig reaction, a McMurry Reaction or a Knoevenagel condensation.

4,4'-Dibromostilbene 2 was synthesized by either a Wittig reaction or a McMurry reaction in moderate yields (Scheme 3). The conditions of this modified Wittig reaction were to use $\text{K}_2\text{CO}_3$ as a base and 18-crown-6 as a catalyst for a solid/liquid transfer process compared with the usual conditions of a McMurry reaction, which requires an absolute moisture- and air-free environment. The cis-isomer was converted to trans-isomer by refluxing in benzene with a catalytic amount of iodine for 8 hours.

\[
\begin{align*}
\text{Br} & \quad + \quad \text{CHO} \quad \text{PPh}_3\text{Br} \\
\text{Br} & \quad \text{K}_2\text{CO}_3/\text{THF} \quad \text{Br} \\
\text{2) cat. I}_2, \text{toluene} \quad \text{reflux, 8 h} \\
\text{Yield 53%} \\
\text{Br} & \quad \text{Br} \\
\end{align*}
\]

Scheme 3: Synthesis of 4,4'-dibromostilbene 2

1,2-Di-(4-bromo-phenyl)-1-cyano-ethene 3 was synthesized by a Knoevenagel reaction (Scheme 4). Stirring of a mixture of 4-bromobenzaldehyde and 4-bromophenyl-acetonitrile in KO-t-Bu/t-BuOH at 55°C for 3 hours gave monomer 3 in 68% yield.
A trifloromethyl group was introduced into monomer 4 for comparison with monomer 3 with regard to the effect of substituents on the luminescence properties of the polymers. Halomethylation of compound 2 with paraformaldehyde and HBr (generated from reaction of NaBr and H₂SO₄) in acetic acid gave compound 11 in 68% yield, which was reacted with PPh₃ in CHCl₃ to generate phosphonium salt 12 in 91% yield. A Wittig reaction between halophenyl trifluoromethyl ketone and phosphonium salt 12 in CH₃CN produced monomer 4 (Scheme 5).

Monomer 6 was synthesized by a Wittig reaction using KO-t-Bu as a base in benzene in 43% yield (Scheme 6). The pure trans product was obtained after the cis product was isomerized to its trans form as earlier described.
PPP, which contains only phenyl units in the main chain, has also been studied for use in LED devices. The PPP-based LED gives bright blue light. The syntheses of PPP are usually based on the nickel(0) catalyzed homocoupling of di(methylsulfonyl)-, di(trifluoromethanesulfonyl)- and dihalo-benzenes in the presence of excess zinc or by palladium(0) catalyzed heterocoupling (Suzuki coupling) in toluene and aqueous K₂CO₃ solution (Scheme 7).

**Scheme 6: Synthesis of monomer 5**

PPP, which contains only phenyl units in the main chain, has also been studied for use in LED devices. The PPP-based LED gives bright blue light. The syntheses of PPP are usually based on the nickel(0) catalyzed homocoupling of di(methylsulfonyl)-, di(trifluoromethanesulfonyl)- and dihalo-benzenes in the presence of excess zinc or by palladium(0) catalyzed heterocoupling (Suzuki coupling) in toluene and aqueous K₂CO₃ solution (Scheme 7).

**Scheme 7: Synthesis of poly(paraphenylene) (PPP)**
In order to obtain special optical properties, the PPP unit was added into the polymer main chains. The PPP unit was synthesized by a Suzuki reaction as shown in Scheme 8.

\[ \text{CHO-} \overset{10, \text{Pd(PPh}_3)_4, \text{Na}_2\text{CO}_3}{\text{Toluene}} \overset{\text{K-O-t-Bu, benzene, Yield 55\%}}{\rightarrow} \text{CHO-} \]

\[ \text{Br-} \overset{\text{PPh}_3\text{Br}}{\rightarrow} \text{Br-} \]

\[ \text{CN} \]

\[ \text{OC}_{12}\text{H}_{25} \]

\[ \text{10} \]

\[ \text{13} \]

\[ \text{7} \]

\[ \text{8} \]

Scheme 8: Synthesis of monomer 7 and 8

\( o \)-Formylphenylboronic acid\(^{40} \) and Na\(_2\)CO\(_3 \) were allowed to react with compound 10 in refluxing DME solution for 2 days, according the Suzuki reaction, to give dialdehyde 13 in 51\% yield. Compound 13 is a yellow solid, which emits blue luminescence in THF solution. A Wittg reaction of dialdehyde 13 and phosphonium salt in benzene gave monomer 7 in 55\% yield. A Knoevenagel condensation\(^{24} \) between dialdehyde 13 and 4-bromophenylacetonitrile
in t-BuOH/THF solution gave monomer 8 in 85% yield. Both monomers 7 and 8 are yellow solids and have strong blue luminescence in THF solution.

All the monomers were characterized by mass spectroscopy, $^1$H- and $^{13}$C-NMR.

**Polymer Preparation.** Preparation of the polymers was based on Heck-type coupling as in Scheme 9. Several palladium catalyst systems have been used to synthesized a variety of polymers\(^\text{52}\) by this coupling reaction. It is well known that when $X = I$, high molecular weights can be obtained while when $X = Br$, the reactions usually require high temperature, long reaction time and give low molecular weights. In 1994, Moigne developed a catalyst system PdCl$_2$/CuAc$_2$/PPh$_3$ to obtain high molecular weight polymers when $X = Br$.\(^\text{25d}\) Thus, our polymerizations were carried out using this catalyst system in Et$_3$N/THF solution. The exclusion of oxygen is necessary to prevent catalyst deactivation and to minimize the copper-catalyzed coupling of acetylenes to diacetylenes which can compete with the desired reaction. The solutions of two monomers in Et$_3$N/THF in the presence of the catalysts were refluxed for three days. After work up, the polymers were simply purified by adding their saturated THF solutions into methanol to precipitate them several times. Most monomers were dibromo- compounds, while the others were diiodo- compounds. The yields of polymerizations were high as 90%, except polymers 19 and 20 (about 65%).

**Polymer Characterization.** The polymers were characterized by GPC, $^1$H- and $^{13}$C-NMR. Molecular weight measurements were performed by gel permeation chromatography (GPC) calibrated by standard polystyrene using THF as an eluent. The weight average molecular weights ($M_w$) of the polymers range from $5 \times 10^3$ to $8 \times 10^4$ depending on the molecular weights of the monomers (Table 3). Polymer 14 has a molecular weight of $8.53 \times 10^3$ which is close to that of polymer 15 ($8.7 \times 10^3$). The higher yield and higher molecular weight of polymer 16 ($M_w = 2.47 \times 10^4$, yield 90.2%) were obtained when diiodo-monomer 4fb) were used instead of dibromo-monomer 4fa) ($M_w = 9.84 \times 10^3$, yield 75.4%). The molecular weight of polymer 18 ($M_w = 4.58 \times 10^3$) is low compared with the other polymers. This is because of the impurity of 4,4'-diiodobiphenyl (only 95%, Aldrich). The higher molecular weight monomers gave the higher molecular weight polymers but the yields were
Scheme 9: Synthesis of the polymers
low (polymer 19 with $M_w = 3.21 \times 10^4$ and yield 65.5%, Polymer 20 with $M_w = 7.74 \times 10^4$ and yield 66.0%). However the degree of polymerization (DP) are about 10 for polymers 14-18 and 25 for polymer 19, and 60 for polymer 20. Polydispersities (PD) of these polymers were 1.45-2.57, which are typical for step growth polymerizations, except polymer 20 (PD = 5.71) due to the large monomer. Table 3 summarizes the results of polymerization.

Table 3: The yields, weight average molecular weights ($M_w$), polydispersities (PD) and degrees of polymerization (DP) of the polymers

<table>
<thead>
<tr>
<th>Polymer</th>
<th>Yield (%)</th>
<th>$M_w$</th>
<th>PD ($M_w/M_n$)</th>
<th>DP</th>
</tr>
</thead>
<tbody>
<tr>
<td>14</td>
<td>92.5</td>
<td>$8.53 \times 10^3$</td>
<td>2.44</td>
<td>13.7</td>
</tr>
<tr>
<td>15</td>
<td>89.0</td>
<td>$8.70 \times 10^3$</td>
<td>2.14</td>
<td>13.4</td>
</tr>
<tr>
<td>16</td>
<td>75.4 (X=Br)</td>
<td>$9.84 \times 10^3$</td>
<td>2.12</td>
<td>8.0</td>
</tr>
<tr>
<td></td>
<td>90.2 (X=I)</td>
<td>$2.47 \times 10^4$</td>
<td>1.98</td>
<td>20.0</td>
</tr>
<tr>
<td>17</td>
<td>90.0</td>
<td>$8.92 \times 10^3$</td>
<td>1.38</td>
<td>13.6</td>
</tr>
<tr>
<td>18</td>
<td>91.1</td>
<td>$4.58 \times 10^3$</td>
<td>1.45</td>
<td>7.7</td>
</tr>
<tr>
<td>19</td>
<td>65.5</td>
<td>$3.21 \times 10^4$</td>
<td>2.57</td>
<td>25.6</td>
</tr>
<tr>
<td>20</td>
<td>66.0</td>
<td>$7.74 \times 10^4$</td>
<td>5.71</td>
<td>59.4</td>
</tr>
</tbody>
</table>

The $^1$H-NMR spectra of the polymers are slightly different from those of the respective monomers. All the resonances are broader, and no signal was detected in the 3.3 ppm region for the acetylenic proton in any case. The protons of -CH$_2$- groups attached by the oxygen in the side chain are at ~4.00-3.20 ppm. The $^{13}$C-NMR spectra show two signals between 100 and 84 ppm, which represent two unsymmetric acetylenic carbons. The carbons of aromatic ring with alkoxy side groups and acetylene groups are detected at about 150, 115, and 113 ppm. The carbons belonging to the alkoxy side chains appear as very strong signals.
Figure 8: The $^1$H-NMR and $^{13}$C-NMR spectra of polymer 14
at 70 ppm and from 32 to 13 ppm. A specific example of the $^1$H-NMR and $^{13}$C-NMR spectra of polymer 14 is shown in Figure 8. In the $^{13}$C-NMR spectrum, the acetylenic carbons are at 94.07 and 86.15 ppm. The aromatic ring with side chains are at 152.58, 115.78 and 112.93 ppm. There are a total of 8 carbon signals between 153 to 112 ppm corresponded to the aromatic and vinyl carbons. There is only one vinyl carbon signal due to the symmetry of the polymer chain.

As for the polymers which contain two phenyl rings with alkoxy side groups in the main chain, the $^1$H-NMR spectra show two kinds of -CH$_2$- groups attached by the oxygen in the side chains. The $^{13}$C-NMR spectra show these two kinds of methylene signals at ~70 ppm and the two kinds of aromatic carbons attached by alkoxy groups at ~150 ppm receptively (see Figure 9 for an example). The $^{13}$C-NMR spectrum shows that polymer 16 has two acetylenic carbons at 93.38 and 86.65 ppm, two aromatic carbons of two phenyl rings attached by side chains at 152.96 and 150.06 ppm, two methylene carbons attached by oxygen in the side chains at 68.75 and 67.95 ppm. These methylene protons also can be seen clearly in the $^1$H-NMR spectrum at 4.00 and 3.26 ppm.

Properties of the Polymers

Solubilities. All the polymers are yellow in the solid states except polymer 17 which is reddish. They are all quite soluble in organic solvents such as THF, CHCl$_3$, CH$_2$Cl$_2$ and toluene due to the long alkoxy side group. They are green-yellow in THF and have green luminescence, except polymer 17, which is red, and has strong green-blue luminescence.

Thermal Behaviors. Thermal behaviors of these polymers were studied by thermogravimetric analysis (TGA) and differential scanning calorimetric analysis (DSC) under nitrogen atmosphere. The TGA was performed from room temperature to 600$^\circ$C with a ramp of 20$^\circ$C/min. From the TGA spectra, these polymers have similar thermal stabilities. They start to lose weight rapidly at about 350-390$^\circ$C (about 5% weight loss) and after 450$^\circ$C they slowly lose weight. At 600$^\circ$C, char yields are about 45%, except polymer 17 (29%). For example, polymer 14 starts weight loss at 395$^\circ$C (3% weight loss), and after 474$^\circ$C (48%
Figure 9: The $^1$H-NMR and $^{13}$C-NMR spectra of polymer 16
weight loss) the weight loss is slow. At end (600°C), the weight loss is 52%. A summary of the TGA data is shown in Table 4.

The thermal transitions of these polymers were studied by DSC as shown in Figure 10. The endothermic peaks usually show the melting process ($T_m$) while exothermic peaks show crosslinking. All these polymers have an exothermic peak before they start to decompose except polymers 15 and 20. DSC for polymer 14 shows melting at ~127°C and exothermic reactions starting at ~194°C and 283°C due to crosslinking. Polymer 15 has only one exothermic peak at ~178°C starting at 153°C, which shows crosslinking happens before polymer melting. There are two endothermic peaks at ~56°C and 77°C and an exothermic reaction at ~197°C for the polymer 16. Polymer 17 has a small endothermic peak at 119°C and starts to crosslink at 135°C and 256°C. Polymer 18 has two endothermic peaks at 69°C and 109°C, and starts to crosslink at 286°C. Polymer 19 melts at about 74°C with a small endothermic peak and start to crosslink at 161°C. There is no melting peak before polymer 20 starts to crosslink at ~158°C. After cooling to room temperature, the brown crystals were

### Table 4: TGA and DSC results of the polymers

<table>
<thead>
<tr>
<th>Polymer</th>
<th>TGA</th>
<th>DSC</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>Temperature (°C)</td>
<td>% remained weight at 600°C</td>
</tr>
<tr>
<td>14</td>
<td>395 (97)</td>
<td>48</td>
</tr>
<tr>
<td>15</td>
<td>390 (93)</td>
<td>46</td>
</tr>
<tr>
<td>16</td>
<td>381 (95)</td>
<td>45</td>
</tr>
<tr>
<td>17</td>
<td>353 (95)</td>
<td>29</td>
</tr>
<tr>
<td>18</td>
<td>333 (94)</td>
<td>50</td>
</tr>
<tr>
<td>19</td>
<td>352 (95)</td>
<td>42</td>
</tr>
<tr>
<td>20</td>
<td>373 (96)</td>
<td>49</td>
</tr>
</tbody>
</table>
Figure 10: DSC thermogram of polymers 14-20
usually recovered. All these crystals were not soluble in normal organic solvents anymore, indicating crosslinking.

**Conductivities.** Conjugated polymers are capable of being doped, through chemical or electrochemical oxidation or reduction, to states of higher electrical conductivity. The conductivities of these polymers were measured by two in-line probes. All these polymers were insulators when undoped. After doping with I₂ vapor under vacuum, they became semiconducting. Polymer 14 has a conductivity of \( \sigma = 4-10 \times 10^{-4} \text{ S/cm} \), while polymer 15 with its -CN group has a conductivity of \( \sigma = 2-4 \times 10^{-3} \text{ S/cm} \). Other polymers also have very low conductivities of \( \sigma \) about \( 1 \times 10^{-4} - 10^{4} \text{ S/cm} \), which are similar to those of PPE, except polymer 17 with conductivity of 50-100 S/cm.

**UV/Vis Absorption.** The UV/Vis absorption spectra of these polymers were recorded and are shown in Figure 11. The absorption edges (\( \lambda_{\text{e}} \)) were estimated within \( \pm 5 \text{ nm} \). All the polymers show strong absorptions in the visible range between 384 to 472 nm(Table 5). Polymer 14 was synthesized as comparison to PPE and PPV. The \( \lambda_{\text{max}} \) polymer 14 in THF solution is 414 nm, which is smaller that those of PPE (429 nm²⁵(a,d)) and PPV (459 nm²¹), probably due to the low molecular weight of polymer 14 compared with PPE and PPV. The absorption edge (\( \lambda_{\text{e}} \)) of polymer 14 is 464 nm, which could be used to estimate a bandgap of 2.67 eV. The -CN group was introduced into polymer 15 as a study of effects of the substituent on the electronic properties of the polymer. The \( \lambda_{\text{max}} \) of polymer 15 in THF is 14 nm red-shifted compared with that of polymer 14, indicating the bandgap of polymer 15 is lower than that of polymer 14. The effect of an electron-withdrawing group of CN was studied by Holmes et al in CN-PPV (Figure 5).²⁴ They reasoned that the electron withdrawing group -CN would make polymer more electronegative, lower the LUMO and HOMO levels, thus improving the electron injection to the polymer layer in LED devices. The CN-PPV-based LEDs had internal quantum efficiencies of 4-10%, the highest efficiencies reported. In 1994, Heeger et al. calculated the influence of donor and acceptor substituents on the electronic characteristics of PPV and PPP. It also showed that CN lowers the LUMO and HOMO energies. Figure 12 shows the bandgaps, HOMO and LUMO energies. The low LUMO level can match to the low workfunction metal like calcium, thus improving the efficiency of
Figure 11: The UV spectra of the polymers in THF solution
Figure 12: Sketch of the band structures of the polymers, relative to the workfunctions of Ca and Al

Another electron withdrawing group -CF₃ was introduced into polymer 16 as a comparison with polymers 14 and 15. Both -CF₃ and -CN are electron withdrawing groups, but -CN exhibits electron-withdrawing by resonance while -CF₃ exhibits electron-withdrawing by sigma induction. However, the $\lambda_{\text{max}}$ of polymer 16 is at a much shorter wavelength than that of polymer 15, and even polymer 14. The most likely rationalization is that the bulky trifluoromethyl group reduces conjugation by causing the phenyl to twist. In 1996, Holmes et al. also studied the effect of electron-withdrawing groups -CF₃ and CN on the model compounds of PPV (Figure 13).²⁴ It was found that CF₃ substituted derivatives have shorter wavelength of UV/Vis absorption than those CN substituted derivatives.

Replacement of the phenyl by thiophene in the polymer main chain greatly changes the UV absorption. Thus $\lambda_{\text{max}}$ of polymer 17 (472 nm) is considerable red-shifted from $\lambda_{\text{max}}$ of all of the other polymers. Conjugated polymers with thiophene rings usually allow a better delocalization of $\pi$-orbitals than the corresponding phenylene rings. It is because thiophene has less aromaticity than phenyl, thus allowing the $\pi$-electrons more mobility.²⁵ For example
PPP has an electronic bandgap of 3.0 eV (413 nm) as compared with 2.1 eV (590 nm) for PT.\textsuperscript{41}

Increasing PPP units in the polymer increases the bandgap energies. The UV/Vis absorption $\lambda_{\text{max}}$ of polymer 18 (396 nm in THF) is 33 nm longer than that of PPE (429 nm) even polymer 18 has only one more phenyl in the repeating units. Compared the UV/Vis absorption of polymer 19 with polymer 14 and polymer 20 with polymer 15, polymer 19 is 26 nm longer than polymer 14 while polymer 20 is 12 nm longer than polymer 15. As mentioned earlier, the more phenyl rings in the polymer main chain, the more aromaticity of the polymer, thus the more localization of the $\pi$-electrons in the phenyl region.

The UV/Vis absorption spectra of these polymer films were also measured and the absorption edges ($\lambda_\infty$) were estimated (Table 5). The films were made by dropping polymer THF solutions onto spinning quartz plates. After the solvent was evaporated, the films formed on the plates were directly measured. The films of polymers 14, 16, 18 and 20 have similar absorption peaks as their solutions, but with a red shift of the main peak of about 27 to 58 nm. Polymer 14 in the film has a 44 nm red-shift compared with in THF solution. Polymer 15 has the largest red-shift of 58 nm. Polymer 16 has a 43 nm red-shift. Polymer 16 has a red-shift of 43 nm. Polymer 17 has the smallest red-shift of 27 nm. Polymer 18 is about 36 nm red-shifted. Polymer 19 has a red-shift of 34 nm. Polymer 20 has a larger red-shift of 56 nm. The films of polymers 15, 17 and 19 have quite different absorption peaks from their
solutions. Besides the main peaks red-shift, the peaks in film states are more broad and the tails of the peaks more extend. The tails of absorption of polymer 15 ($\Delta \lambda_{\text{max}} = 58 \text{ nm}, \Delta \lambda_{\omega} = 78 \text{ nm}$) and 19 ($\Delta \lambda_{\text{max}} = 34 \text{ nm}, \Delta \lambda_{\omega} = 62 \text{ nm}$) films extend about 20 nm ($\Delta \lambda_{\omega} - \Delta \lambda_{\text{max}}$) longer than their main peaks. The tail of absorption of polymer 17 ($\Delta \lambda_{\text{max}} = 27 \text{ nm}, \Delta \lambda_{\omega} = 75 \text{ nm}$) extends even 48 nm longer. This suggests the conformations in the film states allow much better conjugation.

**Table 5:** UV/Vis absorption maxima ($\lambda_{\text{max}}$) and band edges ($\lambda_{\omega}$) of the polymers in THF solution and their films

<table>
<thead>
<tr>
<th>Polymer</th>
<th>THF solution</th>
<th>Film</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>$\lambda_{\text{max}}$(nm)</td>
<td>$\lambda_{\omega}$(nm) (eV)</td>
</tr>
<tr>
<td>14</td>
<td>414</td>
<td>464 (2.67)</td>
</tr>
<tr>
<td>15</td>
<td>428</td>
<td>482 (2.58)</td>
</tr>
<tr>
<td>16</td>
<td>384</td>
<td>433 (2.87)</td>
</tr>
<tr>
<td>17</td>
<td>472</td>
<td>535 (2.32)</td>
</tr>
<tr>
<td>18</td>
<td>396</td>
<td>440 (2.82)</td>
</tr>
<tr>
<td>19</td>
<td>378</td>
<td>458 (2.71)</td>
</tr>
<tr>
<td>20</td>
<td>416</td>
<td>467 (2.66)</td>
</tr>
</tbody>
</table>

**Fluorescence.** All the polymers are photoluminescent in both solutions and the solid states. The photoluminescence (PL) spectra of these polymers in THF solutions have been obtained (Figure 14). The solutions were excited by the UV/Vis light with excitation wavelength ($\lambda_{\alpha}$) about 10 nm shorter than the corresponding absorption maximum.
Figure 14: The photoluminescence (PL) spectra of the polymers
wavelength ($\lambda_{\text{max}}$). The emission bands are sharper than the absorption bands, but with broad peaks at the long tail regions. The emission bands are usually largely red-shifted compared with their absorption bands. These broad emission bands can be associated with recombination processes between defect levels within the gap which may result from interruption of the conjugation of the chains. The emission of polymer $15$ with a peak at 512 nm is dramatically shifted relative to the analogous polymer $14$ with a peak at 456 nm. The emission of polymer $14$ has two small, broad emission side bands at 485 nm and 509 nm while polymer $15$ has only one broad side band at 646 nm. Polymer $16$ has an emission band at 510 nm with a shoulder around 487 nm. The shoulder is due to recombination from the high vibration states of the excitation state to the ground state. Polymer $17$ has the longest emission band at 532 nm with two broad side bands at 568 and 658 nm. This indicates polymer $17$ has the lowest bandgap, which is also shown in its the longest UV/Vis absorption. Polymer $18$ has only a sharp emission band at 434 nm and a small band at 510 nm. The emission of polymer $20$ with a band at 510 nm is large red-shifted compared with that of polymer $19$ with a band at 448 nm. Like polymer $15$, polymer $20$ has a broad side band at 604 nm. The emission spectra of these polymers are summarized in Table 6.

<table>
<thead>
<tr>
<th>Polymer</th>
<th>$\lambda_{\text{ex}}$ (nm)</th>
<th>$\lambda_{\text{max}}$ (nm)</th>
</tr>
</thead>
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<tr>
<td>$14$</td>
<td>410</td>
<td>456, 485(br), 509(br)</td>
</tr>
<tr>
<td>$15$</td>
<td>420</td>
<td>512, 646(br)</td>
</tr>
<tr>
<td>$16$</td>
<td>375</td>
<td>510, 487</td>
</tr>
<tr>
<td>$17$</td>
<td>460</td>
<td>532, 568(br), 658(br)</td>
</tr>
<tr>
<td>$18$</td>
<td>380</td>
<td>434, 510</td>
</tr>
<tr>
<td>$19$</td>
<td>365</td>
<td>448, 510(br), 570(br)</td>
</tr>
<tr>
<td>$20$</td>
<td>405</td>
<td>510, 604(br)</td>
</tr>
</tbody>
</table>
The Polymer-Based LEDs

Because these polymers emit strong fluorescence upon excitation, the films of these polymers could be used as the emissive layer in LED devices. In a collaboration with Prof. J. Shinar (Department of Physics, Iowa State University), the polymer 14-based LED device has been made and studied. The LED device structure is shown in Figure 2. The LED devices were fabricated by spin coating Standish LCD indium-tin-oxide (ITO)-coated glass with a 6 mg/ml polymer 14 solution in toluene or chloroform either in air or in a protective nitrogen atmosphere. The Standish ITO was partially stripped by the manufacturer. The ITO was normally 50 μm thick with a sheet resistance of ~200 Ω. The polymer solutions were pumped through 0.2 or 0.5 μm syringe filters into clean glassware for storage until use. The polymer layer had typical thickness of 100-200 μm. Al contacts with an area of about 10 mm² was thermally or e-beam evaporated onto the polymer layer in vacuum.

I-V measurements was performed with the device submerged in liquid N₂ for heat-removal. Lifetimes under these conditions were indefinite. However, measurement at room temperature yielded lifetimes of only several hours. The electroluminescence (EL) and photoluminescence (PL) spectra were measured by directing the collected light from the device into a Jarrell-Ash monochromator with 7 nm resolution, followed by a Si photodiode. Figure 15 presents a typical I-V curve of a polymer 14-based device at 77 K. As clearly seen, it is symmetric in both forward and reverse bias voltages with a turn-on voltage of about 12 V. This symmetric I-V curve is quite different from the non-symmetric curves reported for normal PPV-based or PPE-based LEDs. The electroluminescence intensity (I_{EL}) was also measured and the I_{EL}-V curve is also presented in Figure 15. It closely follows the I-V curve, and the device clearly emits symmetrically with either polarity of the electrodes. In addition, the curve is similar to that of forward biased ITO/PPV/Al LEDs, suggesting a quantum efficiency of about 0.05%.

In order to study the nature of the reverse bias emission, the electroluminescence (EL) spectra was measured in both biasing directions under identical values of V and I, and are shown in Figure 16. The spectra are identical and very similar to the PL, which is also shown
Figure 15: The $I$-$V$ curve of polymer 14-based LED device (solid line) and the dependence of the integrated intensity of the EL on the applied bias voltage (dotted line).

Figure 16: The EL spectra for positive and negative bias voltage of 18V of polymer 14-based LED device (solid line) compared to the PL spectrum (dotted line).
in Figure 16. This similarity shows that the mechanism leading to emission in both biasing directions is identical and probably due to exciton recombination. This is particularly significant for the reverse bias spectra, where avalanche emission could otherwise be suspected. The coincidence of the EL and PL is also similar to that observed in PPV-based devices and distinct from that of PPE-based emitters.

In PPV-based LEDs, the $I-V$ characteristics of diodes prepared with different metal electrodes have indicated that metal and ITO work functions determine the barrier heights for electron and hole injection, respectively, at their interfaces with the polymer. This picture is invalid in the present case of polymer 14-based structure. This behavior could be explained by Fermi energy ($E_F$) pinning by a high density of deep defect states at interfaces. These defects can capture carriers from the bulk, creating barriers at the interfaces. Whereas contact of a clean surface with a metal usually results in charge exchange between the metal and the surface states, in the case of a high density of surface states the barrier height at the interface is not affected by presence of the metal, resulting in pinning of $E_F$ at the defect states.

For polymer LEDs, basic tunneling theory predicts

$$I \propto V^2 \exp(-b/V),$$  \(\text{Eq. (1)}\)

where $b = 4(2m^*)^1/2\phi/3d/(3e\hbar)$ for a triangular barrier of height $\phi$ and $m^*$ is the effective mass of the carrier. A plot of $\ln(I/V^2)$ vs. $I/V$, based on Figure 15, is shown in Figure 17(a). It indicates the Eq. (1) fits the observed behavior from 14 to 21 V, but substantially deviates from it at low bias voltages. One-step tunneling is not expected to be valid at low voltages, since the tunneling distance $x = d\phi/V$ is then too large. For example, $x > 10 \ \mu$m at $V = 1 \ \text{V}$ and $d = 100 \ \mu$m. It has, therefore, been suggested that the excess injection current at low $V$ is due to sequential tunneling of carriers via localized states near $E_F$. That model predicts the following relationship

$$\ln(I) = -\gamma(d/V)^{1/2},$$  \(\text{Eq. (2)}\)

where $\gamma = 4(\Gamma/3)^{1/2}(2m^*/e^2\hbar^3)^{1/4}\phi^{3/4}$, $\Gamma = -\ln(p)$, and $p$ is the probability that a carrier will find another localized state with an energy close to $E_F$. Figure 17(b) display $\ln(I)$ vs. $V^{1/2}$ for the
polymer 14-based diode. The curve is linear for $V < 13$ V, which is just below the bias range at which Eq. (1) fits the results.

Conclusions

A series of copolymers containing units of PPV, PPE and PPP were synthesized by polymerization of diethynylene arylenes and dibromo- (or diiodo-) arylenes in the palladium catalyst system. All the polymers were characterized by $^1$H and $^{13}$C NMR, GPC, UV/Vis, TGA and DSC.

The thermal stabilities and thermal behaviors of these polymers were studied. Most of these polymers start to decompose at above $350^\circ$C. They usually cross-link at about $300^\circ$C.
The conductivities were also studied. At undoped states, all these polymers are nonconductive. After doped by I$_2$ vapor in the vacuum, the polymers are semi-conductive with conductivities of about $1 \times 10^{-4}$-$10 \times 10^{-4}$S/cm.

The optical properties of these polymers were studied. The UV/Vis absorption and photoluminescence spectra were measured. The electron withdrawing group, -CN, can reduce the bandgap energy, thus increasing the $\lambda_{\text{max}}$ values of UV/Vis absorption and emission. The replacement of phenyl with thiophene can enhance the delocalization of $\pi$-orbitals, thus dramatically increasing the $\lambda_{\text{max}}$ of UV/Vis absorption. Increasing the number of PPP units into the repeating unit can increase the localization of $\pi$-orbitals, thus decreasing the $\lambda_{\text{max}}$ of UV/Vis absorption. The photoluminescence spectra show there are several broad peaks at tail regions besides the major emission bands, indicating the defects in the polymer main chains. These defects are due to the interruption of the conjugation of the polymer main chains.

Polymer-based LEDs were studied. The polymer 14-based device showed symmetric $I$-$V$ curve, which is quite different from the non-symmetric curves reported for PPV-based devices. This behavior could be explained by the defects in intersurface of the polymer and the metal or ITO and defects in the polymer main chain. This property could broaden the application of polymer-based LEDs.

Experimental

$^1$H and $^{13}$C-NMR spectra were acquired on a Varian VXR-300 spectrometer. In order to assure the quantitative features of the $^{13}$C-NMR spectra, the relaxation agent chromium(III) acetylacetonate was used in CDCl$_3$ with a relaxation delay of 5 seconds.

Routine GC-MS spectra were obtained on a Hewlett Packard 5970 GC-IR-MS spectrometer at 70 eV. The exact masses were obtained from a Kratos MS 50 mass spectrometer with 10,000 resolution. The infrared spectra were recorded on a Bio-Rad Digilab FTS-7 spectrometer from neat sample. The UV/Vis spectra were obtained on a
Hewlett Packard 8452A diode array UV/VIS spectrometer and $\lambda_{\text{max}}$ were determined at optical densities of 0.2-0.5.

Polymer molecular weights were determined by gel permeation chromatography (GPC) with 6 Microstyrigel columns in series of 500 A, $2 \times 10^3$ A, $2 \times 10^4$ A, $2 \times 10^5$ A. THF was used as an eluent at a flow rate of 1 ml/min. The system was calibrated by polystyrene standards. GPC analyses were performed on a Perkin-Elmer series 601 LC equipped with Beckman solvent delivery system, a Walter Associate R401 refractive index detector and a Viscotek viscometer. Differential scanning calorimetry (DSC) analyses were performed on a Du Pont 910 Differential Scanning Calorimeter. The UV/Vis spectra were measured on a Hewlett Packard 8452A Diode Spectrophotometer. Photoluminescence spectra were measured on a FL 900 fluorometer made by Edinburgh.

Toluene and benzene were distilled over CaH$_2$. THF was distilled from sodium/benzophenone before use. Other reagents were used as received from Aldrich Chemical Co. without further purification unless specified otherwise. Acetic acid, isopropanol and methanol were used as received from Fisher without further purification. All the reactions were performed under argon atmosphere. 1,4-Didodecyloxy-benzene was synthesized in good yield according to the literature. 4-Bromo-2,2,2-trifluoroacetophenone and o-formylphenylboronic acid were synthesized according to the literature.

**1,4-Didecyloxy-2,5-diiodobenzene 10**: To a solution of 1,4-didecoxybenzene (15.557 g, 34.88 mmol) in acetic acid (300 ml) were added ZnCl$_2$ (11.860 g, 87.2 mmol) and BTMAI$\text{Cl}_2$ (30.345 g, 87.2 mmol). The mixture was stirred at 50-60°C for 1 day. The reaction was quenched by adding water (200 ml) and then a saturated sodium bisulfite aqueous solution (300 ml) was added to consume the unreacted BTMAI$\text{Cl}_2$. The precipitate was collected by filtration and washed by water. The product was purified by precipitation from MeOH/THF twice to give a white solid (24.346 g, yield 53.5%). m.p. 59-61°C; Mass: m/z cal. for C$_{36}$H$_{52}$O$_2$I$_2$ = 698.20567, measured (HiRes EI) = 698.20548; $^1$H-NMR: $\delta$ 7.130 (s, 2H), 3.883 (t, 4H, $J$ 6.3 Hz), 1.744 (m, 4H), 1.474-1.100 (br, 36H), 0.875 (t, 6H, $J$ 6.9
1,4-didecyloxy-2,5-bis[(trimethylsilyl)ethynyl]benzene: To a degassed solution of 1,4-didecoxy-2,5-diiodobenzene 10 (3.00 g, 5.15 mmol) in a mixture of benzene (50 ml) and triethylamine (50 ml) were added trimethylsilylacetylene (1.35 g, 13.7 mmol) and a mixture of CuI (21 mg) and PdCl$_2$(PPh$_3$)$_2$ (85 mg). The solution was stirred for 12 hours at room temperature under an argon atmosphere. After the salt was removed by filtration, the solvent was evaporated to afford a yellow solid. Recrystalization of this solid from a mixture of methanol and isopropanol gave a yellow crystal (2.45 g, 90%). m.p. 73-74°C; Mass: m/z cal. for C$_{40}$H$_{70}$O$_2$Si$_2$ = 639.16560, measured (HiRes EI) = 639.16478; $^1$H-NMR: δ 6.845 (s, 2H), 3.897 (t, 4H, J 3.3 Hz), 1.744 (m, 4H), 1.474-1.100 (br, 36H), 0.851 (t, 6H, J 6.9 Hz), 0.212 (s, 18H); $^{13}$C-NMR: δ 153.26, 116.4, 113.16, 100.33, 99.34, 68.71, 31.23, 31.23, 28.94, 28.73, 28.66, 25.32, 22.02, 13.47, -0.72

1,4-Bis(bromomethyl)-2,5-didecyloxy-benzene 11: To a 1000 ml three neck round bottom flask fitted with a mechanical stirrer, a thermometer, and an additional funnel was charged with 1,4-didecoxybenzene 9 (26.76 g, 60 mmol), paraformaldehyde (11.65 g), NaBr (14.86 g, 144 mmol) and CH$_3$COOH (500 ml). A solution of H$_2$SO$_4$ in CH$_3$COOH (20 ml) was added dropwise over 20 minutes. The mixture was then heated to an internal temperature of 80°C overnight. After the solution was allowed to cool to room temperature, the resulting white precipitate was collected by filtration and saved. The filtrate was then poured into ice water. The aqueous layer was extracted by CH$_2$Cl$_2$ (2 x 200 ml). After solvent was removed, the residue was precipitated from methanol. The combined white solid was further purified by precipitation from CH$_2$Cl$_2$/MeOH (29.000 g, yield 76.5%). m.p. 94-96°C; Mass: cal. m/z for C$_{32}$H$_{56}$O$_2$Br$_2$ = 632.26278, C$_{32}$H$_{56}$O$_2$Br$_2$ = 630.26469, C$_{32}$H$_{56}$O$_2$Br$_2$ = 634.26086, measured (HiRes EI) = 632.26230, 630.26418, 634.26101; $^1$H-NMR: δ 6.826 (s, 2H), 4.506 (s, 4H), 3.961 (t, 4H, J 6.3 Hz), 1.790 (m, 4H), 1.543-1.240
(br, 36H), 0.863 (t, 6H, J 6.9 Hz); \textsuperscript{13}C-NMR: \( \delta 150.65, 127.49, 114.62, 68.99, 31.91, 29.66, 29.63, 29.59, 29.34, 29.32, 28.74, 26.06, 22.68, 14.11 \)

\textbf{[2,5-Bis(dodecyloxy)-1,4-phenylene]dimethylenebis(triphenylphosphonium-bromide) 12:} To a solution of 1,4-bis(bromomethyl)-2,5-didecyloxy-benzene 11 (1.870 g, 2.96 mmol) in absolute CHCl\(_3\) (40 ml) was added PPh\(_3\) (1.552 g, 5.92 mmol). The solution was refluxed for 3 days. After the solution was cooled to room temperature, the solvent was removed under vacuum. The residue was dissolved in minimum THF and then was added to hexane, resulting a white precipitate. The solid was collected by filtration and then was dried under vacuum (2.912 g, 85.0%). m.p. 146\(^\circ\)C; \textsuperscript{1}H-NMR: \( \delta 7.72 \) (br, 30H), 6.78(s, 2H), 5.30 (d, 4H, \( J 13 \) Hz), 3.05 (t, 4H, \( J 4.5 \) Hz), 1.543-1.240 (br, 36H), 0.863 (t, 6H, \( J 6.9 \) Hz)

\textbf{1,4-didecyloxy-2,5-diethynyl-benzene 1:} To a degassed solution of 1,4-didecyloxy-2,5-bis[(trimethylsilyl)ethynyl]benzene (2.100 g, 4.74 mmol) in methanol (100 ml) was added a saturated NaOH aqueous solution (0.05 ml). The solution was stirred for 1 hour at 50\(^\circ\)C. Then the solution was cooled to 0\(^\circ\)C and a yellow solid was precipitated. The product was further purified by recrystallization of this solid from methanol to obtain a yellow crystal (1.200 g, 85%). m.p. 106-107.5\(^\circ\)C; Mass: \textit{m/z} cal. for C\(_{34}\)H\(_{54}\)O\(_2\) = 494.41237, measured (HiRes EI) = 494.41165; \textsuperscript{1}H-NMR: \( \delta 6.909 \) (s, 2H), 3.925 (t, 4H, \( J 6.6 \) Hz), 3.288 (s, 2H), 1.756 (m, 4H), 1.440-1.100 (br, 36H), 0.840 (t, 6H, \( J 6.9 \) Hz); \textsuperscript{13}C-NMR: \( \delta 153.34, 117.10, 112.60, 82.08, 79.26, 69.07, 31.35, 29.10, 29.10, 28.79, 28.54,25.34, 22.14, 13.60 \)

\textbf{4,4'-Dibromostilbene 2:} (a) Wittig Reaction.\textsuperscript{30} To a round bottom flask equipped with a condenser was charged with 4-bromophenylmethytriphenylphosphonium bromide (3.072 g, 6 mmol), K\(_2\)CO\(_3\) (0.828 g, 6 mmol) and dry THF (25 ml). Then 4-bromo-benzaldehyde (0.925 g, 5 mmol) and 18-crown-6 (15 mg) were added. The solution then was refluxed for 2 days. After the solution was cooled to the room temperature, the inorganic salt was removed by filtration and solvent was removed under vacuum. The residue was purified by column chromatography to obtain a white solid (0.900 g, yield 53%).
(b) McMurry Reaction: To a flame-dried three-necked flask equipped with a magnetic stirring bar, an additional funnel and an argon outlet was added activated Zn (3.923 g, 60 mmol) and THF (80 ml). The mixture was cooled to -10°C, then TiCl₄ (5.7 g, 30 mmol) was added dropwise, resulting in a green-yellow clean solution. The solution was refluxed for 4 hours. The resulting black solution was allowed to cool to room temperature. Then 4-bromobenzaldehyde (3.700 g, 20 mmol) in THF (20 ml) was added. The solution was refluxed overnight, then was allowed to cool to room temperature. To the black mixture was added aqueous K₂CO₃ solution (10%, 160 ml). After stirring for 1 hour, the precipitate was removed by filtration. The filtrate was extracted with ether (300 ml x 2). The organic layer was dried with MgSO₄. After removal of ether, the residue was purified by column chromatography to obtain a white solid (2.031 g, yield 60.1%).

The cis-isomer was transferred to the trans-isomer by refluxing the isomer mixture in absolute benzene with a catalytic amount of iodine for 8 hours. Mass: cal. m/z for C₁₄H₁₀Br⁻⁻Br⁻⁻Br = 337.91301, C₁₄H₁₀Br⁻⁻Br₂ = 335.91492, C₁₄H₁₀Br⁻⁻Br₂ = 339.91067, measured (HiRes EI) = 337.91248, 335.91423, 339.91067, respectively; ¹H-NMR: 7.461 (d, 4H, J 8.4 Hz), 7.342 (d, 4H, J 8.4 Hz), 6.997 (s, 2H)

1,2-di-(4-bromo-phenyl)-1-cyano-ethene 3: This compound was prepared by a Knoevenagel condensation. To a solution of 4-bromobenzaldehyde (0.925 g, 5 mmol) and 4-bromophenylacetonitrile (0.980 g, 5 mmol) in t-BuOH (100 ml) was added a solution of KO-t-Bu (0.672 g, 6 mmol) in t-BuOH (25 ml) at 55°C. The mixture was stirred at 55°C for 3 hours. After the mixture was cooled, hydrochloric acid (1 M, 50 ml) and water (50 ml) were added. The aqueous layer was extracted with CH₂Cl₂ (100 ml x 2). The combined organic layer was washed with water (50 ml x 2), dried with MgSO₄. After solvent removed, the product was purified by crystallization from hexane to get a white solid (1.230 g, yield 68%). m.p. 126-128°C; Mass: m/z cal. for C₁₅H₁₅N⁻⁻Br⁻⁻Br = 362.90826, C₁₅H₁₅N⁻⁻Br₂ = 360.91017, C₁₅H₁₅N⁻⁻Br₂ = 364.90634, measured (HiRes EI) = 362.90816, 360.91013, 364.90608; ¹H-NMR: δ 7.708 (d, 2H, J 8.4 Hz), 7.561 (d, 2H, J 8.4 Hz), 7.536 (d, 2H, J 8.1
1,4-Bis[2-(4-halophenyl)-2-trifloromethyI-vinyl]-2,5-bis(dodecyloxy)benzene 4:
The product was prepared by a similar procedure according to the literature.36 The mixture of
4'-halo-2,2,2-trifluoroacetophenone35 (2.024 g, 8 mmol), [2,5-bis(dodecyloxy)-1,4-
phenylene]-dimethylene-bis(triphenylphosphoniumbromide) 12 (4.626 g, 4 mmol), KF (9.28 g,
160 mmol) and dibenzo-18-crown-6 (0.8 mmol, 0.31 mg) in CH3CN (140 ml) was stirred at
70-80°C for 5 hours. Then after the solvent was removed, CH2Cl2 was added to dissolve
product. KF was removed by filtration. After most of CH2Cl2 was removed, the product was
precipitated out by adding CH2Cl2 solution into MeOH. 4(a): X = Br, a yellow solid (2.763 g,
yield 74.0%). m.p. 85-87°C; Mass: m/z cal. for C48H62Br2F6O2 = 944.30015, measured
(GC-MS): 944 (M+ (100%), 942 (M+-Br) (39%), 496 (M+ for
C48H62Br2F6O2) (30%); 'H-NMR: δ 7.507 (d, 4H, J 8.1 Hz), 7.454 (d, 2H, J 1.5 Hz), 7.145
(d, 4H, J 8.1 Hz), 6.108 (s, 2H), 3.239 (t, 4H, J 6.6 Hz), 1.554 (m, 4H), 1.258 (br, 36H),
0.863 (t, 6H, J 6.6 Hz); 13C-NMR: δ 150.0, 131.6, 131.4, 131.0, 127.2, 122.8, 122.4, 112.8,
67.9, 31.2, 28.9(2), 28.8, 28.6, 28.5, 28.2, 25.1, 22.0, 13.4; 4(b): X = I, a yellow solid
(4.152g, yield 74.4%). m.p. 91-93°C; Mass: m/z for C48H62Br2I2O2 = 1038.27434, measured
(GC-MS): 1038 (M+ (100%), 912 (M+-I) (31%) 869 (18%), 702 (27%), 71 (22%), 69
(26%), 57 (72%); 'H-NMR: δ 7.703 (d, 4H, J 8.1 Hz), 7.434 (d, 2H, J 1.5), 7.005 (d, 4H, J
8.1 Hz), 6.093 (s, 2H), 3.222 (t, 4H, J 6.6 Hz), 1.555 (m, 4H), 1.294 (br, 36H), 0.853 (t, 6H,
J 6.6 Hz); 13C-NMR: δ 149.84, 137.38, 131.84, 131.84, 130.60, 126.95, 122.66, 112.65, 93.89,
67.75, 31.03, 28.69(2), 28.46, 28.09, 25.09, 21.82, 13.27

2,5-Bis(dodecyloxy)-1,4-bis[2-(5-bromo-2-thienyl)vinyl]benzene 6: The product
was synthesized by a similar procedure in the literature.33 [2,5-Bis(dodecyloxy)-1,4-
phenylene]dimethylene-bis(triphenylphosphoniumbromide) 12 (2.313 g, 2 mmol) and 4-
bromo-2-thiophene carbaldehyde were dissolved in benzene (50 ml). Then KO-t-Bu (0.749 g,
6.7 mmol) was added. The solution was refluxed for 24 hours. The solvent was removed,
and H₂O (30 ml) and CH₂Cl₂ (30 ml) were added. The aqueous layer was extracted with CH₂Cl₂ (30 ml x 2). The combined organic layer was washed with H₂O, dried with MgSO₄. After CH₂Cl₂ was removed, the product was purified by precipitation from THF/MeOH twice to obtain a yellow solid (0.710 g, yield 43%). m.p. 75-78°C; Mass: m/z for C₄₂H₆₀Br²BrO₂S₂ = 820.23822, measured (GC-MS): 820 (M⁺) (100%), 818 (M⁺ for C₄₂H₆₀Br₂O₂S₂) (16%), 822 (M⁺ for C₄₂H₆₀Br²BrO₂S₂) (18%), 742 (13%), 740 (13%), 322 (17%), 321 (66%), 320 (17%), 97 (18%), 82 (12%), 71 (37%), 69 (20%), 57 (100%); ¹H-NMR: δ 7.132 (s, 2H), 7.116 (s, 2H), 6.952 (s, 2H), 6.925 (d, 2H, J 3.8 Hz), 6.761 (d, 2H, J 3.8 Hz), 3.986 (t, 4H, J 6.3 Hz), 1.831 (m, 4H, J 6.9 Hz), 1.553-1.252 (br, 36H), 0.866 (t, 6H, J 6.6 Hz); ¹³C-NMR: δ 150.4, 144.8, 129.9, 125.5, 123.1, 121.1, 110.4, 110.0, 68.9, 31.3, 29.0(2), 28.8, 28.7, 25.6, 22.1, 13.5

2',5'-Bis(dodecyloxy)-p-terphenyl-1,4''-dicarboxaldehyde 13: To an aqueous Na₂CO₃ solution (2 M, 75 ml) was added a solution of o-formylphenylboronic acid*® (1.80 g, 12 mmol) and 2,5-bis(dodecyloxy)-1,4-diiodobenzene (2.30 g, 3 mmol) in DME (75 ml), followed by addition of Pd(PPh₃)₄ (346 mg). The mixture was refluxed for 2 days. After mixture was allowed to cool to room temperature, the aqueous layer was separated and was extracted by CH₂Cl₂ (100 ml x 2). The combined organic layer was washed with water, dried with MgSO₄. After the solvent was removed, the product was precipitated from THF/MeOH twice to obtain a yellow solid (1.00 g, yield 51%). m.p. 86-90°C; Mass: m/z cal. for C₄₄H₆₂O₄ = 654.97388, measured (HiRes EI) = 654.97346; ¹H-NMR: δ 10.049 (s, 2H), 7.240 (d, 4H, J 8.4 Hz), 6.993 (s, 2H), 6.925 (d, 2H, J 3.8 Hz), 6.761 (d, 2H, J 3.8 Hz), 3.986 (t, 4H, J 6.3 Hz), 3.929 (t, 4H, J 6.3 Hz), 1.831 (m, 4H, J 6.9 Hz), 1.553-1.252 (br, 36H), 0.866 (t, 6H, J 6.6 Hz); ¹³C-NMR: δ 191.0, 149.1, 143.4, 133.8, 129.0(2), 128.3, 114.6, 68.4, 30.7, 28.5, 28.4, 28.2, 28.1, 24.9, 21.5, 13.0

1,4''-Bis[2-(4-bromophenyl)vinyl]-2',5'-bis(dodecyloxy)-p-terphenyl 7: KO-t-But (0.596 g, 5.32 mmol) was added a refluxing compound 13 solution (0.620 g, 0.95 mmol) and 4-bromophenylmethyltriphenylphosphonium bromide (1.460 g, 2.85 mmol) in benzene (25 ml). The mixture was refluxed for 1 day. After the solution was cooled to room temperature,
CH₂Cl₂ (25 ml) and a saturated NH₄Cl solution (25 ml) were added. The aqueous layer was extracted by CH₂Cl₂ (25 ml). The combined organic layer was washed with water and dried with MgSO₄. After the solvent was removed, the product was precipitated twice from THF/MeOH to obtain a yellow solid (0.50 g, yield 55%). m.p. 155-160°C; Mass: m/z for C₅₈H₇₂Br₂O₂ = 960.35887, measured (GC-MS): 960 (M⁺) (100%), 962 (M⁺ for C₅₈H₇₂Br₂O₂) (16%), 958 (M⁺ for C₅₈H₇₂Br₂O₂) (18%); ¹H-NMR: δ 7.611 (d, 4H, J 8.1 Hz), 7.539 (d, 4H, J 8.1 Hz), 7.473 (d, 4H, J 8.7 Hz), 7.383 (d, 4H, J 8.7 Hz), 7.112 (s, 2H), 7.086 (s, 2H), 6.991 (s, 2H), 3.914 (t, 4H, J 6.3 Hz), 1.675 (m, 4H), 1.541-1.219 (br, 36H), 0.854 (t, 6H, J 6.9 Hz); ¹³C-NMR: δ 149.8, 137.4, 135.8, 135.0, 131.3, 129.7, 129.3, 128.7, 127.5, 126.7, 125.6, 120.7, 69.1, 31.3, 29.0(2), 28.8, 25.5, 22.1, 13.6

1,4'''-Bis[2-(4-bromophenyl)-2-cyano-vinyl]-2',5'-bis(dodecyl-oxy)-p-terphenyl 8:
To a solution of compound 13 (0.580 g, 0.89 mmol) and 4-bromophenyl-acetonitrile (0.452 g, 2.30 mmol) in t-BuOH (10 ml) and THF (10 ml) was added a KO-t-Bu solution in t-BuOH (10 ml) at room temperature dropwise. The mixture was stirred at 55°C for 4 hours. Then solvent was removed by evaporation under vacuum. CH₂Cl₂ (50 ml) was added. The organic layer was washed with H₂O and dried with MgSO₄. After removal of CH₂Cl₂, the product was precipitated from THF/MeOH to obtain a light yellow solid (0.750 g, yield 84.3%). m.p. 126-127°C; Yellow color luminescence in THF; Mass: m/z for C₆₀H₇₀Br₂N₂O₂ = 1010.37847, measured (GC-MS): 1010 (M⁺) (9%), 1012 (M⁺ for C₆₀H₇₀Br₂N₂O₂) (5%), 1008 (M⁺ for C₆₀H₇₆Br₂N₂O₂) (4%), 676 (13%), 675 (13%), 674 (26%), 673 (13%), 594 (4%), 321 (5%), 71 (31%), 69 (36%), 57 (83%), 44 (100%); ¹H-NMR: δ 7.950 (d, 4H, J 8.7 Hz), 7.722 (d, 4H, J 8.4 Hz), 7.566 (s, 8H), 7.552 (s, 2H), 7.011 (s, 2H), 3.944 (t, 4H, J 6.6 Hz), 3.944 (t, 4H, J 6.6 Hz), 1.705 (m, 4H), 1.358-1.216 (br, 36H), 0.841 (t, 6H, J 6.9 Hz); ¹³C-NMR: δ 149.72, 141.92, 140.28, 132.91, 132.51, 131.67, 131.36, 129.44, 128.51, 125.96, 122.72, 117.15, 115.14, 109.28, 86.02, 68.98, 31.26, 29.02, 28.93, 28.70, 28.62, 25.41, 22.04, 13.50

Polymer 14-20: Polymerization was performed by Heck coupling according to the literature. The monomers (1 mmol of each), palladium chloride (19.4 mg), copper acetate
(3.3 mg) triphenylphosphine (117 mg), triethylamine (50 ml) and THF (20 ml), all dried and degassed, were put into a two-necked 150 ml flask equipped with a magnetic stirrer and a condenser. The reaction mixture was refluxed for 3 days. The precipitated ammonium salt was filtered off and washed with THF. The solvent was removed, the residue was dissolved in hot THF (3 ml), poured into cold methanol (200 ml) to precipitate. The product was collected by filtration. The polymer was further purified by precipitation from MeOH/THF twice or more and dried under vacuum.

**Polymer 14:** Yield 92.5%; A yellow solid; 
\[ ^1\text{H-NMR: } \delta 7.507, 7.114, 7.054, 4.030, 1.085, 1.547-1.244, 0.857; \]
\[ ^13\text{C-NMR: } \delta 152.58, 135.95, 130.95, 127.64, 125.61, 121.69, 115.78, 112.93, 94.07, 86.15, 68.61, 30.92, 28.65, 28.40, 25.09, 21.70, 13.16; \]
\[ \text{GPC: } M_n = 3.50 \times 10^3, M_w = 8.53 \times 10^3, \text{PD} = 2.438; \text{UV/Vis: } \lambda_{\text{max}} = 414 \text{ nm (THF), 458 nm (film)} \]

**Polymer 15:** Yield 89.0%; A yellow solid; 
\[ ^1\text{H-NMR: } \delta 7.865-7.844, 7.638-7.495, 7.002, 6.965, 3.990, 1.813, 1.500, 1.208, 0.823; \]
\[ ^13\text{C-NMR: } \delta 153.21, 131.60, 128.92, 125.43, 117.11, 116.30, 113.38, 110.73, 94.22, 88.50, 87.80, 69.14, 31.49, 29.22, 25.63, 22.27, 13.72; \]
\[ \text{GPC: } M_n = 4.05 \times 10^3, M_w = 8.70 \times 10^3, \text{PD} = 2.145; \text{UV/Vis: } \lambda_{\text{max}} = 428 \text{ nm (THF), 486 nm (film)} \]

**Polymer 16:** For X = Br, yield = 75.4%; A yellow solid; GPC: \( M_n = 1.24 \times 10^4, M_w = 2.47 \times 10^4, \text{PD} = 1.983; \text{UV/Vis: } \lambda_{\text{max}} = 382 \text{ nm (THF)}; \) For X = I, yield = 90.2%; A yellow solid; GPC: \( M_n = 6.36 \times 10^3, M_w = 9.84 \times 10^3, \text{PD} = 1.546; \text{UV/Vis: } \lambda_{\text{max}} = 384 \text{ nm (THF), 427 nm (film)}; \)
\[ ^1\text{H-NMR: } \delta 7.521, 7.468, 7.281, 6.970, 4.001, 3.264, 1.830, 1.533, 1.403-1.218, 0.842; \]
\[ ^13\text{C-NMR: } \delta 152.96, 150.06, 132.33, 131.32, 129.31, 123.42, 122.95, 115.95, 113.13, 112.87, 93.38, 86.65, 68.75, 67.95, 31.21, 28.95, 25.38, 25.13, 21.99, 13.41 \]

**Polymer 17:** Yield 90.0%; A brown solid; 
\[ ^1\text{H-NMR: } \delta 7.258, 7.248, 7.165, 7.019, 6.983, 6.961, 4.029, 1.861, 1.59-1.256, 0.873; \]
\[ ^13\text{C-NMR: } \delta 152.27, 149.81, 144.47, 124.95, \]
123.20, 120.65, 116.42, 115.22, 112.63, 109.43, 89.83, 87.74, 68.44, 30.76, 28.49, 28.26, 24.90, 24.91, 21.55, 13.02; GPC: $M_n = 6.48 \times 10^2$, $M_w = 8.92 \times 10^3$, $DP = 1.377$; UV/Vis: $\lambda_{\text{max}} = 472$ nm (THF), 499 nm (film)

**Polymer 18**: Yield 91.1%; A yellow solid; $^1$H-NMR: $\delta 7.60, 7.04, 4.05, 1.85, 1.56, 1.23, 0.85$; $^{13}$C-NMR: $\delta 153.21, 139.5, 131.64, 126.35, 122.32, 116.43, 113.55, 94.33, 86.59, 69.20, 31.49, 29.23, 25.67, 22.26, 13.70; GPC: $M_n = 3.15 \times 10^3$, $M_w = 4.58 \times 10^3$, $DP = 1.451$; UV/Vis: $\lambda_{\text{max}} = 396$ nm (THF), 432 nm (film)

**Polymer 19**: Yield 65.5%; A yellow solid; $^1$H-NMR: $\delta 7.59, 7.13, 7.02, 4.04, 3.94, 1.85, 1.71, 1.58, 1.25, 0.87$; $^{13}$C-NMR: $\delta 152.88, 149.89, 137.14, 135.59, 131.45, 130.92, 129.39, 127.51, 125.83, 116.07, 115.58, 95.74, 86.64, 69.19, 31.47, 25.61, 22.24, 13.68; GPC: $M_n = 4.05 \times 10^3$, $M_w = 8.7 \times 10^3$, $DP = 2.145$; UV/Vis: $\lambda_{\text{max}} = 378$ nm (THF), 412 nm (film)

**Polymer 20**: Yield 66.0%; A yellow solid; $^1$H-NMR: $\delta 7.961, 7.719, 7.607, 7.025, 4.031, 3.956, 1.846, 1.716, 1.596, 1.229, 0.847$; $^{13}$C-NMR: $\delta 152.94, 149.65, 141.58, 140.21, 133.36, 131.47, 129.35, 128.47, 125.18, 123.80, 117.20, 116.10, 115.08, 113.24, 109.71, 93.60, 87.23, 68.922, 31.20, 28.93, 25.35, 21.98, 13.47; GPC: $M_n = 1.36 \times 10^4$, $M_w = 7.74 \times 10^4$, $DP = 5.707$; UV/Vis: $\lambda_{\text{max}} = 416$ nm (THF), 472 nm (film)
II: SYNTHESIS, CHARACTERIZATION AND STUDY OF CONJUGATED POLYMERS CONTAINING SILOLE UNIT IN THE MAIN CHAIN

Literature Survey

Since the discovery in 1977 that polyacetylene (PA) could be n- or p-doped, either chemically or electrochemically, to nearly the metallic state, the development of the field of conducting polymers has continued to accelerate at a rapid rate. This rapid growth has been stimulated not only by the field's synthetic novelty and multi-disciplinary importance but also by its actual and potential technological application.

One of the fundamental challenges of the field of conducting polymers is the design and synthesis of low bandgap polymer. These low bandgap polymers can have good intrinsic conductivity without doping, good nonlinear optical and photoelectric properties and amphivalent electrochemical characteristics. The electronic and optical properties of conjugated polymers originate mostly from their π-electrons. In the simplest case of a linear chain, the band gap is proportional to the bond length alternation due to the Peierls instability in 1D. The bond length alternation (δr) is defined as the average of the difference between neighboring long and short C-C bonds. Polyacetylenes, which have the lowest bond length alternation, are the most conductive polymers. However, due to their poor environmental stabilities, more attentions are paid to design and synthesis of environmentally stable conjugated polymers with band gap energy lower than 1 eV. Heterocyclic five-membered ring systems (Figure 18) such as polypyrroles, polythiophenes, are of most interest because of their high stability and conductivity. The band gaps (Eg) of heterocyclic five-membered ring systems are determined by both geometrical and heteroatomic effects. In polythiophene, for example, the sulfur changes the bond length by locking in planarity of the butadiene unit. At the same time, the lone pair of electrons on sulfur strongly participates in the aromaticity of the ring. In 1987, Kertesz et al. calculated the effect of the heteroatomic substitutions on the band gap of polyacetylene. The bond length alternation (δr) of polythiophene was shown to actually be smaller than the δr of polyacetylene (see Table 7). However, due to the orbital
symmetry, only the LUMO of the backbone can interact with the $3p_z$ orbital of sulfur, which shifts the LUMO level upward by 0.89 eV, but leaves almost unchanged HOMO level. Thus, the main effect of sulfur in the aromatic ring is to enlarge $E_g$.\textsuperscript{56,78}

In 1989, our group re-examined the thermal polymerization of diethynyldiphenylsilane and the catalytic polymerization of diethynylsilane with MoCl$_5$ or WCl$_6$ was studied, which gave a deep red polymer and a violet polymer respectively.\textsuperscript{51} The resulting polymer was first suggested by Shinar to be polysilole (PS) (Figure 19).\textsuperscript{51} However, one year later, the structure was described as polydiethynylsilane (PDS) having a four-membered backbone ring (Figure 19).\textsuperscript{52(a)} This four-membered structural model is supported by $^{13}$C- and $^{29}$Si-NMR and resonance Raman scattering measurements,\textsuperscript{52(b, c)} and by some ab initio Hartree-Fock calculations of the backbone structure.\textsuperscript{53} The violet polymer, which has a very low bandgap energy value of 2.0 eV, is conductive upon doping with I$_2$ vapor with conductivity of $10^{-1}$ S/cm. The non-linear $\chi^{(3)}$ value at 625 nm was measured to be $3 \times 10^{-9}$ esu, one of the largest measured in conjugated polymers.

![Figure 18: Heterocyclic five-membered ring systems](image)

![Figure 19: The structures of poly(diethynylsilane) and polysilole](image)
In the course of elucidating the structure of PDS, several groups calculated both polysilole and poly(diehtynyloilane)\textsuperscript{53,54,55} to compare with polyacetylene and other heterocyclic five-membered rings (see Figure 20).\textsuperscript{53,54,55} In contrast to the heteroatoms in those heterocyclic five-membered ring systems which were studied, silicon (in PD and PDS) doesn't have a lone pair of electrons to interact with the conjugated \( \pi \)-orbitals of the systems.

![Diagram of polymer structures](image)

**Figure 20:** Conformation of PS, PDS, PA and other heterocyclic ring systems

In 1992, the calculations by Frapper et al. showed that 3p silicon orbitals do not participate in the conjugation because of their sp\(^3\) hybrid states, and therefore the \( \pi \)-electron system is similar to that of polyacetylene.\textsuperscript{54(a)} Table 7 gives the calculated bond alternation and bandgap energies of acetylene isomers, poly(diehtynyloilane) and polysilole. Both PS and PDS have very low bandgap energies, and are good candidates for high intrinsic conductive materials.

A surprising result, found by Grigoras et al.\textsuperscript{53,54} and Frapper et al.,\textsuperscript{54(a,b)} indicated that while both aromatic and quinoid structures of polysilole are more stable by about 15-20
kcal/mol than that of poly(diethynylsilane) due to the four-membered ring strain, the quinoid structure of polysilole is only slightly more stable than its "aromatic" structure by 2 kcal/mol-ring (Figure 20). In 1995, Dr. Hong et al. calculated the conformation and electronic structures of poly(cyclopentadienylene) (PPD) (Figure 20, X = CH₂) and polysilole. The calculation confirmed that the ground state geometries of polysilole (Figure 20, X = SiH₂, SiF₂) and poly(cyclopentadienylene) (Figure 20, X = CH₂, CF₂) are quinoid conformations. The bridging groups affect the band gaps in two principle ways: by decreasing the C1-C4 distance and by pure electronic effects. The large δr values of both PS, PPD and PDS compared to the δr values for PT, PPy and PF imply that the CH₂, SiH₂ bridging groups interact with the polymeric backbone less strongly than do the heteroatoms, S, N and O.

Recently, Bakhshi et al. designed novel donor-acceptor polymers which contained silole backbones, and calculated their bandgap energies. They were found to have very low

<table>
<thead>
<tr>
<th>Table 7: Bond alternation and bandgap energy</th>
</tr>
</thead>
<tbody>
<tr>
<td>Polyacetylene</td>
</tr>
<tr>
<td>Polyfuran</td>
</tr>
<tr>
<td>(A) (Q)</td>
</tr>
<tr>
<td>δr</td>
</tr>
<tr>
<td>E₉</td>
</tr>
<tr>
<td>δr</td>
</tr>
<tr>
<td>E₉</td>
</tr>
</tbody>
</table>

δr: Bond alternation, in angstroms; E₉: Bandgap energy, in eV
bandgap energies. The principal idea behind donor-acceptor polymers is that a regular alternation of conjugated donor- and acceptor-like moieties in a conjugated chain will induce a low bandgap. The designed polymers (Figure 21) can be viewed as analogues of trans-cisoid polyacetylene (cis-PA), but stabilized by different bridging groups (X as an electron donor and Y as an electron acceptor). Although these are not known polymers, they are calculated to have bandgaps lower than 1 eV.

\[ X = \text{CH}_2, \text{SiH}_2, \text{GeH}_2; \ Y = \text{C} = \text{CH}_2, \text{C} = \text{O}, \text{C} = \text{CF}_2 \]

**Figure 21:** The structure of donor-acceptor polymers

Because PS and PDS are much similar to polyacetylene, they should exhibit higher second order hyperpolarizabilities like PA. In 1992, Grigoras et al. calculated the second-order hyperpolarization \( <\gamma> \) of model structures of PS, PDS, PA and PDA (Figure 22) and found that PS and PDS have comparative \( <\gamma> \) values to PA.\(^{53}\)

In 1995, Matsuzaki et al. calculated the longitudinal second hyperpolarizabilities (\( \gamma_{zzz} \)) of both aromatic and quinoid forms of PS and PDS, as well as PA (with polymerization degree \( n \) up to 8) by the finite field (FF) method, including electron correlation at the level of the Moller-Plesset second-order perturbation theory (MP2).\(^{58}\) It is surprising that PS shows the largest \( \gamma_{zzz} \) value among the \( \pi \)-conjugated polymers and the strongest \( n \)-dependence (\( n \) is degree of polymerization) of \( \gamma_{zzz} \).

Up to now polysilole have not been synthesized yet. The most straightforward route to polysilole would be the direct coupling of silole monomers at 2,5-position, which requires two functional groups such as Li, Br, SnR\(_3\) at the 2,5-position. The synthesis and organic chemistry of silole have been recently reviewed.\(^{59}\) The conventional methods to synthesize the siloles include direct or indirect dehydration of silacyclopent-4-en-3-ol (Scheme 10), a
reaction between an alkyne and a disilane or a stannylsilane (Scheme 11), cyclization of 1,4-
dilithio-1,3-butadienes with a polyfunctional compound \( R_nSiX_{4-n} \) (\( n = 0-2 \) (Scheme 12), a
reaction of an alkyne and a silirene (Scheme 13), thermolysis of silirenes (Scheme 14),
dehydrogenation of 1,1-dialkyl-2,5-diphenylsilacyclopentane (Scheme 15),
dehydrohalogenation (Scheme 16), flash vacuum pyrolysis of 1-allylsilacyclopent-3-enes
(Scheme 17), and other methods.\(^5^9\) None of these methods gives a 2,5-difunctional silole.

\[ \gamma = 30.1 \times 10^{-36} \]
\[ \gamma = 27.2 \times 10^{-36} \]
\[ \gamma = 25.1 \times 10^{-36} \]
\[ \gamma = 35.8 \times 10^{-36} \]
\[ \gamma = 32.3 \times 10^{-36} \]
\[ \gamma = 46.6 \times 10^{-36} \]

**Figure 22:** Structures considered for the estimation of second-order polarizability (e.s.u)

![Diagram](attachment:image.png)
HMe₂Si—SiMe₂H + 2 RC≡CR' → cat. "Ni" or "Pd" →
R = R' = Ph, Me, Et; R = Ph, R' = Me, Et

HMe₂Si—SnR₃ + 2 R'C≡CH → cat. "Pd" →
R = Et, Bu, Oct; R' = Ph, CH₂OR

Scheme 11: Reaction between a alkyne and a disilane

Scheme 12: Reaction between an alkyne and a silirene

Scheme 13: Cyclization of 1,4-dilithio-1,3-butadienes
In 1994, Tamao et al. first reported the synthesis of a 2,5-difunctional silole (Scheme 18). They found that diethynyilsilane undergo intramolecular reductive cyclization upon treatment with lithium naphthalenide to form 2,5-dilithiosilole, which could be quenched by electrophiles such as Br₂, R₃SnCl or R₃SiCl. This first reported endo-endo mode of reductive cyclization could be rationalized by the crucial bis(anion radical) intermediate being stabilized by the phenyl and silicon (Figure 23).
However, the attempted polymerization of these 2,5-bifunctional siloles under a variety of condition (including homocoupling of 2,5-dilithiosilole with Fe(III), Ni(II), Cu(II) etc., homocoupling of 2,5-dibromosilole 24 with 2,5-bisstannylsilole 44(a), magnesium reagent 45(b) or zinc reagent 45(a) in the presence of nickel or palladium catalyst), were unsuccessful. Large steric hindrance due to the phenyl group at 3,4-positions of silole and an inherent low reactivity at 2,5-positions of silole may be responsible for these failure. The low reactivity at 2,5-positions of silole may be exemplified by the fact that the silyl groups thereon could not be desilated by a variety of electrophiles. However, the dimer, 3,3',4,4'-tetraphenyl-2,2'-bislole 21 or 5,5'-dibromo-3,3',4,4'-tetraphenyl-2,2'-bislole 22, and tetramer,
dibromoquatersilole 46, of silole were synthesized by oxidative coupling of lithiosiloles via higher order cyanocuprate or with tris(acetylacetonato)iron(III) (Scheme 19).

Scheme 19: Synthesis of dimer and tetramer of silole\textsuperscript{60(a)}

Bisilole 21 and dibromosilole 22 have unusually long UV/Vis absorption with $\lambda_{\text{max}}$ of 398 nm and 417 nm (in chloroform), respectively. For comparison, a carbon analog,
2,2',3,3'-tetraphenyl-1,1'-bicyclopentadiene 23 was synthesized by a similar route to bisilole by the same group two years later.\textsuperscript{61(a)} It is surprising, considering their twisted conformations, that values of $\lambda_{\text{max}}$ of bisiloles are the longest of those of nonfused two-ring $\pi$-conjugated compounds (Table 8). The difference in $\lambda_{\text{max}}$ between 21 and 23 is 58 nm. Recently, the theoretical calculations were carried out to explain this difference.\textsuperscript{62} It was found that the origin of the unusually optical properties was the low-lying LUMO level of the bisilole, arising from $\sigma^*$.\pi* conjugation between a $\pi$-symmetry $\sigma^*$ orbital of exocyclic $\sigma$ bonds on silicon and a $\pi^*$ orbital of the butadiene skeleton.

Table 8: The UV absorption maxima of two-ring $\pi$-conjugated compounds

<table>
<thead>
<tr>
<th>Compound</th>
<th>UV Absorption Maxima</th>
</tr>
</thead>
<tbody>
<tr>
<td>Bisilole 21</td>
<td>398 nm</td>
</tr>
<tr>
<td>Dibromobisilole 22</td>
<td>417 nm</td>
</tr>
<tr>
<td>Bicyclopentadiene 23</td>
<td>340 nm</td>
</tr>
<tr>
<td>2,2'-Bifuran</td>
<td>278 nm</td>
</tr>
<tr>
<td>2,2'-Bithiophene</td>
<td>302 nm</td>
</tr>
<tr>
<td>2,2'-Biselenophene</td>
<td>321 nm</td>
</tr>
</tbody>
</table>

Most recently, in 1996, Tamao et al. evaluated the performance of small molecules containing silole as electron transporting (ET) materials (Scheme 20).\textsuperscript{61(b)} The configuration of the devices is ITO/TPD/ET-material/Mg:Ag (see structures of TPD and ET materials in Scheme 20). Tamao et al. compared the performance of silole derivatives with that of tris(8-hydroxyquinoline)aluminum (Alq), one of the most efficient ET materials reported so far. It
Scheme 20: Synthesis and study of silole derivatives as ET materials

was found that the device PySPy/Alq, consisting of PySPy as the ET material and Alq as the emissive material, exhibited nearly 3 times higher $I$-$V$ efficiency than those obtained with the device using only Alq, and the device with TTSTTT exhibited one order higher $I$-$V$ efficiency.
than that with only Alq. These findings demonstrated for the first time that the silole ring is a promising candidate as a core component for efficient ET materials.

Despite of unsuccess of synthesis of polysilole, the synthesis and study of silole-containing polymers has continued to attract the scientists. In 1991, Corriu et al. synthesized and studied polymers including silole ring in the polymer main chains (Figure 24). However, the silole ring does not participate in the \( \pi \)-conjugation in the polymer main chain as it is insulated by the silicon atom.

The only one example of silole-containing, fully \( \pi \)-conjugated polymers, are the thiophene-silole copolymers and cooligomers shown in Scheme 21. The thiophene-silole-thiophene (TST) was prepared by the nickel-catalyzed intramolecular cyclization of a thiophene-containing 1,6-diyne with a hydrodisilane. TST could be mono- and dibromination with NBS in DMF or mono- and dimetalation by \( \text{n-BuLi/TMEDA} \). By the same method, the other silole-containing compounds such as TTST and TTSTTT were synthesized. By coupling reactions of their brominated monomers and stannylated monomers, a series of thiophene-silole copolymers and oligomers was synthesized. The silole-thiophene 1:1 copolymer could not be prepared by this method. However, its precursors, oligomers (TS)\(_2\)T (S:T = 2:3) and (TS)\(_3\)T (S:T = 3:4) were synthesized.

All these thiophene-silole cooligomers and polymers show much longer absorption in the visible region compared with thiophene homooligomers and homopolymers (Table 9). Significantly, there is a general tendency that a higher silole ratio causes a bathochromic shifts. Of particular interest are the broad absorptions of the copolymer (TST)\(_n\) at 594 and 615 nm,
Scheme 21: The synthesis of thiophene-silole cooligomers and copolymers
Table 9: UV/Vis absorption data for thiophene-silole cooligomers and copolymers and thiophene homooligomers and homopolymers

<table>
<thead>
<tr>
<th>Compound</th>
<th>No. of thiophene and silole</th>
<th>UV absorption $\lambda_{\text{max}}$ (nm)</th>
</tr>
</thead>
<tbody>
<tr>
<td>Terthiophene</td>
<td>$T_1$</td>
<td>353</td>
</tr>
<tr>
<td>Quaterthiophene</td>
<td>$T_4$</td>
<td>391</td>
</tr>
<tr>
<td>Quinquethiophene</td>
<td>$T_5$</td>
<td>418</td>
</tr>
<tr>
<td>Bisilole 21</td>
<td>$S_2$</td>
<td>398</td>
</tr>
<tr>
<td>Dibromosilole 22</td>
<td>$S_2$</td>
<td>417</td>
</tr>
<tr>
<td>Dibromoquartersilole</td>
<td>$S_4$</td>
<td>443</td>
</tr>
<tr>
<td>TST</td>
<td>$T_1 S$</td>
<td>267, 416</td>
</tr>
<tr>
<td>TTSTTT</td>
<td>$T_4 S$</td>
<td>473</td>
</tr>
<tr>
<td>$(TS)_2 T$</td>
<td>$T_4 S_2$</td>
<td>490, 524</td>
</tr>
<tr>
<td>$(TS)_3 T$</td>
<td>$T_4 S_3$</td>
<td>544, 582</td>
</tr>
<tr>
<td>$(TST)_2$</td>
<td>$T_6 S_2$</td>
<td>505</td>
</tr>
<tr>
<td>$(TST)_3$</td>
<td>$T_6 S_3$</td>
<td>549</td>
</tr>
<tr>
<td>$(TST)_n$</td>
<td>$T_{3n} S_n$</td>
<td>594, 615</td>
</tr>
<tr>
<td>$(TTST)_n$</td>
<td>$T_{3n} S_n$</td>
<td>546</td>
</tr>
<tr>
<td>$(TTSTT)_n$</td>
<td>$T_{3n} S_n$</td>
<td>549</td>
</tr>
<tr>
<td>Poly(3-alkylthiophene)</td>
<td>$T_n$</td>
<td>435</td>
</tr>
</tbody>
</table>

which are the longest wavelength ever found for thiophene-based polymers. The $\lambda_{\text{max}}$ value (582 nm) of the seven-ring system $(TS)_3 T$ ($S: T = 3:4$) exceeds the $\lambda_{\text{max}}$ value (549 nm) of the nine-ring system $(TST)_3$ ($S: T = 1:2$). However, the $\lambda_{\text{max}}$ values of dibromobisilole ($S_2$) 22 (417 nm) is about the same as that of TST (416 nm).
Because of the promising optical properties caused by unusual low bandgap of silole, further exploration of the field was conducted. A series of conjugated polymers containing silole was synthesized, and their optical properties were studied.

Results and Discussions

Synthesis and Study of Poly(2,5-thiophene-2,5-silole)

Synthesis. Despite the thiophene-silole copolymers had been synthesized, the copolymer with S:T = 1:1 had not. This copolymer should exhibit the longest $\lambda_{\text{max}}$ of UV absorption among thiophene-silole copolymers. Given the availability of 2,5-difunctional silole, the copolymer with S:T = 1:1, 26, could be synthesized by direct coupling of 2,5-dibromosilole with 2,5-bis(trialkylstannyl)thiophene 25 as shown in Scheme 22.

![Scheme 22: Synthesis of poly(2,5-thiophene-2,5-silole) 26 and model compound 27](image_url)

The polymerization was carried out by refluxing the reactants and the catalyst in THF for 7 days. The polymer was purified by precipitation from MeOH. The polymerization was very slow and the yield was low (about 28%). Polymer 26 is a deep purple solid which is very soluble in organic solvents such as THF and chloroform, giving bright reddish solutions. In
order to characterize polymer 26 by NMR, the model compound 27 was synthesized as the same method. After the reaction mixture was refluxed for 2 days, compound 27 was obtained in a moderate yield (53%) (Scheme 22). Compound 27 is a bright yellow solid and give bright green color luminescence in the solution state.

**Characterization.** The structure of polymer 26 was characterized by comparison with model compound 27. The $^1$H-NMR spectrum of polymer 26 shows two regions around 7.12 and 1.19 ppm representing the aromatic and the hexyl protons respectively. The $^{13}$C-NMR spectra of polymer 26 and model compound 27 are shown in Figure 25. The chemical shifts of sp$^2$ carbons in silole unit of compound 27 are 153.79 and 143.13 ppm. The carbon of thiophene attaching to silole is at 139.37 ppm. Similarly, all these carbons can also be seen clearly in polymer 26 at 153.66, 143.33 and 139.35 ppm, respectively. The molecular weight of polymer 26 measured by GPC was relatively low ($M_w = 2.19 \times 10^3$, PD =1.36). This can be explained by the steric hindrance due to the phenyl groups at 3,4-positions of silole and low reactivity of dibromosilole. This steric hindrance and low reactivity were also seen in the synthesis of model compound 27 with a moderate yield (53%) even at excess trimethylstannylthiophene. The average length of the polymer chain contains about four to five silole and thiophene units. TGA of polymer 26 shows the polymer start to decompose at 334°C. A DSC thermogram of polymer 26 is shown in Figure 24 where an endothermic peak is observed at ~56°C and an exothermic reaction starts at ~165°C, probably due to the crosslinking.

**UV/Vis Absorption.** The UV/Vis absorption spectra of compound 27 and polymer 26 were measured and given in Figure 27. To our surprising, despite of its steric hindrance between phenyl and thiophene and its twisting structure, the UV/Vis maximum absorption $\lambda_{\text{max}}$ (418 nm) of compound 27 is essentially the same as that of TST (416 nm),$^{64}$ which has a smaller steric effect. The $\lambda_{\text{max}}$ value (590 nm) of polymer 26 is much red shifted compared with that of compound 27 (418 nm). However it is still shorter than that of (TST)$_n$ polymer (594 nm, 615 nm) even S:T ratio is 1:1 instead of 1:2. The $\lambda_{\text{max}}$ value of polymer 26 is very close to that of (TS)$_3$T (582 nm). It can be rationalized by the fact that the chain length of polymer 26 is similar to that of (TS)$_3$T. Absorption edge ($\lambda_c$) was estimated as 740 nm, which
Figure 25: $^{13}$C-NMR spectra of compound 27 and polymer 26
Figure 26: DSC thermogram of polymer 26
corresponds to a bandgap energy of 1.67 eV. The UV/Vis absorption of polymer film was also measured. The $\lambda_{\text{max}}$ (626 nm) in the solid state is 36 nm longer than in solution, suggesting that the polymer in the solid state is better aligned than in solution.

**Fluorescence.** Polymer 26 has a red-colored luminescence in THF. The emission spectrum of polymer 26 was recorded and is shown in Figure 28. The deoxygenated THF solution was excited by a laser at 510 nm. The emission maximum is at 631 nm with two shoulders at 674 and 693 nm. The tail of emission band reaches beyond 800 nm.

**Electrical Conductivity.** Polymer 26 is an insulator at its neutral state. When doped by I$_2$ vapor under vacuum, its conductivity increased to $4.3 \times 10^{-5}$ S/cm, which is close to the conductivity of (TS)$_3$T ($3 \times 10^{-5}$ S/cm), but much lower than that of (TST)$_n$ ($9 \times 10^{-3}$ S/cm) and that of well-defined poly(3-alkylthiophene) ($10^2$-$10^3$ S/cm). The reason is probably its low molecular weight and its poor film-forming ability.
Figure 28: The emission spectrum of polymer 26 in THF solution
Synthesis and Study of Poly(2,5-silole-ethynylene) and Poly(2,5-silole-butadiynylene)

Synthesis. The electronic structure of a silole unit is much like that of a butadiene unit. The backbones of poly(silole-ethynyl) and poly(silole-butadiynylene), which have double bond and triple bond repeating units, are much like polydiacetylene. Since silicon can lower the LUMO level of the polymer, they should have even low bandgap energies than polydiacetylene. Poly(silole-ethynylene) or poly(silole-butadiynylene) can be easily synthesized by direct coupling between bis(trialkylstannyl)acetylene or bis(trialkylstannyl)butadiyne and 2,5-dibromosilole 24 (Scheme 23). Unlike thiophene, ethynylene is linear and shouldn’t have any steric hindrance. Indeed, the reaction was much faster than coupling between dibromosilole 24 and 2,5-bis(trialkylstannyl)thiophene 25. The polymerization was carried out by refluxing bis(tributylstannyl)acetylene and dibromosilole 24 in THF in the presence of palladium catalyst for 36 hours. Polymer 28 was precipitated from MeOH as a very deep blue (like black) solid in 93% yield. Polymer 28 is very soluble in organic solvents, exhibiting a deep blue color in THF.

Scheme 23: Synthesis of poly(silole-ethynylene) 28 and poly(silole-butadiynylene) 29 by palladium-catalyzed coupling
However, refluxing of dibromosilole 24 and bis(tributylstannyl)butadiyne in THF solution in the presence of palladium catalyst for only 2 hours resulted in a very viscous, insoluble black gel, obviously the result of crosslinking. Reducing the temperature gave only either no polymer (which could be precipitated from MeOH) or insoluble black gel. The crosslinking could happen because polydiacetylene is not stable and easily crosslinked in the presence of palladium catalyst. In order to avoid crosslinking during the polymerization in the presence of palladium catalyst, synthesis of polymer 29 was attempted by oxidative coupling of 2,5-diethynylsilole 30 in the presence of Cu(I) as a catalyst (Scheme 24).

![Scheme 24: Synthesis of poly(silole-butadiynylene) 29 by oxidative coupling](image)

Attempted syntheses of monomer, 2,5-diethynyl-1,1-dihexyl-3,4-diphenylsilole 30, are presented in Scheme 25. The reaction between dibromosilole 24 and tributylstannylacetylene in THF gave unidentified mixtures. By using an internal acetylene, 1-tributylstannyl-2-trimethylsilyl-ethyne, the coupling reaction went smoothly. The desired product 31 was obtained in high yield, but was contaminated with the byproduct tributylbromotin, which could not be separated from product 31 by column chromatography. The attempted reaction of dibromosilole 24 with trimethylsilylacetylene in Et3NH/benzene in the presence of Pd(PPh3)4/CuI for 4 hours gave no precipitate. However, after 2.2 equivalent of a strong base, 1,8-diazabicyclo-[5,4,0]-undec-7-ene (DBU), was added to this mixture, a white precipitate formed immediately. After further stirring at room temperature for 6 hours, compound 31 was obtained in 57% yield. Detrimethylsilation of compound 31 in MeOH in the presence of a catalytic amount of KOH failed to give the desired product 30, instead a ring-opening product 32 was obtained in 40% yield. No further attempts to prepare monomer 30 was made.
Characterization. Polymer 28 was characterized by NMR, GPC, TGA and DSC. The $^1$H-NMR spectrum (Figure 27) shows only two regions of absorption between 7.3 and 6.9 ppm and between 1.3 and 0.7 ppm, representing the protons of the phenyl and hexyl groups, respectively. The ratio between these two regions is ~2.6, indicating polymer 28 has high molecular weight. The $^{13}$C-NMR spectrum has only two sp$^2$ carbons at ~168 and 138 ppm for silole, four carbons between 161 and 124 ppm for phenyl, only one sp carbon at ~102 ppm for ethynylene carbons and six sp$^3$ carbons for the hexyl groups. This spectrum fits the structure of polymer 28 perfectly. Because of the symmetry of the structure of the polymer, only one sp ethynylenic carbon is expected.

Scheme 25: Attempted synthesis of 2,5-diethynyl-3,4-diphenylsilole 30
Figure 29: $^1$H-NMR and $^{13}$C-NMR spectra of polymer 28
The molecular weight of polymer 28 was measured by GPC; \( M_w = 6.52 \times 10^3 \) with PD = 3.6. The average number of units in the polymer chain is \( \sim 16 \). The TGA spectrum of polymer 28 (Figure 30) shows there is a small endothermic peak at \( \sim 60^\circ C \), then the polymer starts to give off the heat slowly until \( \sim 170^\circ C \). An exothermic peak at \( \sim 240^\circ C \) is probably due to the crosslinking at 202°C.

Due to the crosslinking during the polymerization and failure of synthesis by the other route, polymer 29 cannot be fully characterization by NMR and GPC. So its characterization and electronic properties won’t be discussed in this thesis.

**UV/Vis Absorption.** The UV/Vis spectra of polymer 28 in THF and in the solid state are shown in Figure 29. The UV/Vis absorption \( \lambda_{\text{max}} \) in THF is 604 nm, which is the longest absorption wavelength among those of normal conjugated polymers having similar structure, such as poly(phenylene-ethynylene) \( (\lambda_{\text{max}} = 425 \text{ nm}),^{25(0)} \) and poly(thiophene-ethynylene) \( (\lambda_{\text{max}} = 438 \text{ nm}) \) (see Table 10 in page 80). The lack of the aromaticity of the silole unit and \( \sigma^*-\pi^* \) conjugation between a \( \pi \)-symmetry \( \sigma^* \) orbital of exocyclic \( \sigma \) bonds on silicon, and a \( \pi^* \) orbital of the skeleton\(^{62} \) enhance overlap of \( \pi \)-electrons along the chain, thus dramatically lowering the bandgap energy. The absorption tail extends up to 740 nm. The bandgap energy can be estimated as low as 1.67 eV. Polymer 28 can form a shining smooth metallic film. The pattern of the UV/Vis absorption of the film is very close to that of its THF solution, red-shifted by 28 nm but otherwise very similar.

**Fluorescence.** Unlike PPV and PPE, polymer 28 has only a weak photoluminescence. Polymer 28 is much like polydiacetylene, which also has a very weak photoluminescence. Polymer 28 was excited at a fixed wavelength at 550 nm. Emission peaks were observed at 622 and 655 nm (Figure 32). The reason for the sharp peak at 655 nm is unclear. The small shoulder at 575 nm is possibly due to emission from the high vibration states in the excited state to the ground state.

**Electrical Conductivity.** The conductivity of undoped 28 is small \( (4.5 \times 10^{-7} \text{ S/cm}) \). The conductivity increased dramatically upon doping by I\(_2\) vapor under vacuum. The maximum conductivity \( (2.3 \times 10^1 \text{ S/cm}) \) was essentially achieved in the first 5 minutes. Further doping did not increase the conductivity. Despite its low bandgap energy, the
Figure 30: DSC thermogram of polymer 28
Figure 31: The UV/Vis absorption spectra of polymer 28 in THF solution and in the solid state

conductivity of polymer 28 is still much lower than those of polythiophene, polyacetylene and polydiacetylene.

Synthesis and Study of Poly(silole-ethynylene-paraphenylene-ethynylene)

**Synthesis.** As discussed in the first part of my thesis, there has been a wide variety of poly(arylene-ethynylene) derivatives synthesized and studied (see Figure 7). Since the silole unit has the electronic effect of lowering the LUMO level of the polymer main chain, the synthesis and study of poly(arylene-ethynylene) derivatives containing silole units are of much interest. The synthesis of poly(silole-ethynylene-paraphenylene-ethynylene) could be achieved via dibromosilole 24. The polymerization was first carried out by reacting 24 with 1,4-diethynylbenzene and DBU in Et\(_2\)NH/THF in the presence of Pd(PPh\(_3\))\(_2\)Cl\(_2\)/CuI. The organic salt precipitated during the polymerization. However, during work up, no polymer precipitated from MeOH. This is presumably due to low molecular weight of the polymer which resulted from low reactivity between dibromosilole 24 and 1,4-diethynylbenzene. Reactivity of 1,4-bis(trialkylstannyl)benzene should be higher than that of 1,4-diethynylbenzene because the reductive elimination step during the polymerization does not need a base
Figure 32: The emission spectrum of polymer 28
to participate. The reaction of 1,4-diethynylbenzene with n-BuLi, followed by chlorotributyltin gave one major product 1,4-bis(tributylstannyl)benzene 34(a) according to TLC. However, purification by column chromatography resulted in partial destannylation to give several mixtures which were difficult to separate. When 1,4-bis(tributylstannyl)benzene (Scheme 26) was assumed to be in 100% yield, and was allowed to react directly with dibromosilole 24 and catalytic Pd(PPh₃)₂Cl₂ without further purification, a very low molecular polymer was obtained. This is because the yield of 1,4-bis(tributylstannyl)benzene was not 100%. However, when diethynylbenzene reacted with n-BuLi, followed by quenching with chlorotrimethyltin, the resulting product 1,4-bis(trimethylstannyl)benzene 34(b) (a white solid), could be purified by crystallization from hexane in 46% yield. Stirring 2,5-dibromo-3,4-diphenylsilole 24, 1,4-bis(trimethylstannyl)benzene 34(b) and catalytic Pd(PPh₃)₂Cl₂ in THF at 50°C for seven days gave polymer 33 in 81% yield (Scheme 26). Polymer 33 is very soluble in THF, CHCl₃, CH₂Cl₂. However, after storage in the atmosphere for several days, the polymer was not totally soluble in THF. This indicates polymer 33 is unstable in the atmosphere, possibly due to the crosslinking.

![Chemical structure](attachment:image.png)

**Scheme 26: Synthesis of poly(silole-ethynylene-paraphenylene-ethynylene) 33**

**Characterization.** The ¹H-NMR spectrum (Figure 33) of polymer 33 shows two groups of protons, one at between 7.41 and 6.97 ppm for the phenyl groups and the other at between 1.63 and 0.87 ppm for the hexyl groups. The ¹³C-NMR spectrum (Figure 33) shows
two sp² carbons of silole ring at ~162 and 137 ppm, three big peaks and one small peak between 131 to 123 ppm for the phenyl carbons, two sp ethynylene carbons at ~99 and 92 ppm, and six sp³ hexyl carbons between ~32 and 10 ppm. The NMR spectra fit the structure of polymer 33 quite well. The molecular weight of 33 measured by GPC was 1.43 x 10⁴ (PD = 3.082). The thermal stability of 33 was studied by TGA. The polymer starts to lose weight rapidly at ~285°C (4.5% weight loss) up at ~593°C (39% weight loss). After 593°C, the weight loss is very slow. At 800°C, the weight loss is ~44%. The thermal behavior of 33 was also studied by DSC as shown in Figure 34. The polymer has an endothermic peak at ~73°C and starts an exothermic reaction slowly at ~128°C to give a broad exothermic peak. This probably is due to slow crosslinking.

UV/Vis Absorption. The UV/Vis absorption spectra of polymer 33 in THF solution and in the solid state were measured and shown in Figure 35. The λ_max of polymer 33 is 494 nm in THF solution and 526 nm in the solid state. These numbers are the largest among those of poly(arylene-ethynylene) analogs, except poly(anthracene-ethynylene). Table 10 shows λ_max values of these polymers. The λ_max of polymer 33 is 65 nm longer than that of poly(2,5-dialkoxy-1,4-phenyleneethynylene). By changing the thiophene unit of poly(thiophene-ethynylene-paraphenylene-ethynylene) to silole unit of polymer 33, the λ_max can change 69 nm from 425 to 494 nm. However, compared with the λ_max (604 nm) of poly(silole-ethynylene) 28, the λ_max (494 nm) of polymer 33 is 110 nm shorter. This is the further evidence that the silicon dramatically lowers the bandgaps in these systems.

Fluorescence. Polymer 33 has a weaker photoluminescence than PPV and PPE, but a stronger photoluminescence than polymer 28. It seems apparently that increasing the ratio of phenyl rings to ethynylene enhances the photoluminescence. Polymer 33 was excited at a fixed wavelength at 430 nm. The emission peak was observed at 549 nm (Figure 36). The quantum yield was not measured.

Electrical Conductivity. Like other poly(aryleneethynylene) polymers, polymer 33 is an insulator when undoped. After exposure to I₂ vapor under vacuum, the conductivity increased to 4.3 x 10⁻³ S/cm.
Figure 33: The $^1$H-NMR and $^{13}$C-NMR spectra of polymer 33
Figure 34: DSC thermogram of polymer 33
Figure 35: The UV/Vis spectra of polymer 33 in THF solution and the solid state
Table 10: The $\lambda_{\text{max}}$ of poly(aryleneethynylene) in solution and the solid state

<table>
<thead>
<tr>
<th>Polymers</th>
<th>$\lambda_{\text{max}}$ (solution) (nm)</th>
<th>$\lambda_{\text{max}}$ (film) (nm)</th>
<th>Ref.</th>
</tr>
</thead>
<tbody>
<tr>
<td><img src="image1.png" alt="Structure 1" /></td>
<td>429 (THF)</td>
<td>455</td>
<td>25(d)</td>
</tr>
<tr>
<td><img src="image2.png" alt="Structure 2" /></td>
<td>425 (THF)</td>
<td>453</td>
<td>25(d)</td>
</tr>
<tr>
<td><img src="image3.png" alt="Structure 3" /></td>
<td>508 (THF)</td>
<td>521</td>
<td>25(d, g)</td>
</tr>
<tr>
<td><img src="image4.png" alt="Structure 4" /></td>
<td>404 (THF)</td>
<td>412</td>
<td>25(a)</td>
</tr>
<tr>
<td><img src="image5.png" alt="Structure 5" /></td>
<td>403 (CH$_3$Cl)</td>
<td>410</td>
<td>25(a)</td>
</tr>
<tr>
<td><img src="image6.png" alt="Structure 6" /></td>
<td>360 (HCOOH)</td>
<td>350</td>
<td>25(a)</td>
</tr>
<tr>
<td><img src="image7.png" alt="Structure 7" /></td>
<td>438 (THF)</td>
<td></td>
<td>25(f)</td>
</tr>
<tr>
<td><img src="image8.png" alt="Structure 8" /></td>
<td>426 (CH$_3$Cl)</td>
<td>432</td>
<td>25(a)</td>
</tr>
<tr>
<td><img src="image9.png" alt="Structure 9" /></td>
<td>462 (CH$_3$Cl)</td>
<td>460</td>
<td>25(a)</td>
</tr>
<tr>
<td><img src="image10.png" alt="Structure 10" /></td>
<td>366 (THF)</td>
<td>366</td>
<td>27</td>
</tr>
<tr>
<td><img src="image11.png" alt="Structure 11" /></td>
<td>494 (THF)</td>
<td>526</td>
<td>this work</td>
</tr>
</tbody>
</table>
Figure 36: The emission spectrum of polymer 33
Synthesis and Study of Poly(2,5-silole-ethynylene-dimethylsilylene-ethynylene)

Conjugated polymers interrupted by a silylene unit have σ-π interaction along their main chains. The study of polymers having σ-π conjugated systems is an interesting subject, because of their potential utility as electronically conducting, semiconducting, and emitting polymers. In 1994, our group synthesized and fully studied liquid crystalline properties of silylene-containing polymer, poly(ethynylene-dialkysilylene-ethynylene-1,4-phenylene) (Scheme 27). Like a phenyl, the silole unit has a rigid structure. Therefore, poly(2,5-silole-ethynylene-dialkydsilylene-ethynylene) has a rigid, rod-like backbone, and should have liquid crystalline properties. Silole, much like butadiene, lowers the bandgap energy, thus changing electronic properties of the polymer.

Scheme 27: Synthesis of poly(ethynylene-dialkysilylene-ethynylene-1,4-phenylene)

Synthesis. The synthesis of poly(2,5-silole-ethynylene-dimethylsilylene-ethynylene) is similar to that of polymer 33 (Scheme 28). Bis(trimethylstannylethynyl)-dimethylsilane reacted with dibromosilole in THF in the presence of catalytic Pd(PPh3)2Cl2 to give
polymer 36 in 52% yield. Polymer 36 is a brown solid and very soluble in organic solvents, the THF solution being reddish.

\[
\begin{align*}
\text{Ph} & \quad \text{Ph} \\
\text{Br} & \quad \text{Br} \\
\text{hex} & \quad \text{hex} \\
\end{align*}
\quad + \quad 
\begin{align*}
\text{Me}_3\text{Sn} & \quad \equiv \quad \text{Si} \\
& \quad \equiv \quad \text{SnMe}_3 \\
\end{align*}
\]

\[
\text{Pd(PPh}_3\text{)}_2\text{Cl}_2 \\
\text{THF} \\
\text{Yield 51.6%}
\]

\[
\begin{align*}
\text{Ph} & \quad \text{Ph} \\
\text{hex} & \quad \text{hex} \\
\end{align*}
\quad \equiv \quad 
\begin{align*}
\text{Me} & \quad \equiv \quad \text{Si} \\
& \quad \equiv \quad \text{Me} \\
\end{align*}
\]

\[
36 \quad (M_w = 6.08 \times 10^3, \text{PD} = 1.129)
\]

**Scheme 28: Synthesis of poly(silole-ethynylene-dimethylsilylene-ethynylene) 36**

**Characterization.** The \(^1^H\)-NMR spectrum (Figure 37) of polymer 36 shows one broad peak at 7.09 ppm for the phenyl protons, four peaks between 1.55 and 0.87 ppm for the hexyl groups, one peak at 0.16 ppm for the protons of the dimethylsilyl group. The \(^{13}\)C-NMR spectrum (Figure 37) shows two carbons at ~162 and 137 ppm for the silole unit, three peaks for the phenyl carbons, two ethynylene carbons at ~105 and 102 ppm, six hexyl carbons between 32.32 and 10.29 ppm, and one dimethylsilyl carbon at 0.11 ppm. The molecular weight of polymer 36 is \(6.06 \times 10^3\) (PD = 1.129) according to GPC. TGA shows the polymer starts to decompose at ~366°C (weight loss 4%). The weight loss is rapid between 366°C and 562°C (weight loss 38%). After 562°C, the weight loss is very slow. At 800°C, the weight loss is ~43%. DSC thermogram (Figure 38) shows a strong endothermic peak at ~74°C. This endothermic peak indicates polymer 36 starts to melt at 74°C. An exothermic reaction starts at ~261°C to give a peak at 337°C. This peak is probably due to the crosslinking.
Figure 37: The $^1$H and $^{13}$C NMR spectra of polymer 36
UV/Vis Absorption. The UV/Vis spectra of polymer 36 in THF solution and film were shown in Figure 39. The absorption maximum in THF solution is 400 nm. To fully study the σ-π interaction, the UV/Vis absorption of the model compound 31 was measured. The λ_max of 31 is 336 nm in THF, which is about 64 nm shorter than that of polymer 36. It indicates extensive conjugation despite the interruption by the silylene unit. The silylene unit participates in the conjugation by the σ-π interaction. Compare with the λ_max value of poly(ethynylene-dialkysilylene-ethynylene-1,4-phenylene) 37 (~364 nm), the λ_max of polymer 36 is 36 nm red-shifted, which shows the silole unit decreases the bandgap. The peak shape of the film was also recorded and is similar to that in THF solution, but red-shifts 26 nm. The absorption edge in THF solution is ~460 nm.

Electrical Conductivity. The conductivity of polymer 36 in the neutral state is 4.49 x 10^-4 S/cm. After doped by I_2 vapor, the conductivity didn’t change much with σ = 3.67 x 10^-5 S/cm.

All these polymers should exhibit high χ^(3) values. The study of third-nonlinear optical properties of these polymers should be very interesting.

Attempted Synthesis of Poly(2,5-silole-1,2,3-butatriene)

Organic polymeric materials with π-conjugated systems, which exhibit very large nonlinear optical response, are polyacetylene, poly(phenylenevinylene) (PPV), polythiophene, polydiacetylene and other conjugated polymers (see Table 2). Cumuenes such as butatriene and hexapentaene also have extended π-electrons and linear structures. However, conjugated polymers containing the cumulene unit have been paid little attention. Theoretical calculations of polyenes, polyenynes and cumulenes within the correlated Pariser-Parr-Pople (PPP) model (defined over the π-frame by Albert et al.) showed that for the same chain length, cumulenes have the largest polarizability and third harmonic generator (THG) coefficients. The optical gaps for these systems were also calculated, with cumulenes possessing the smallest gap at 0.75 eV, polyenes at 2.86 eV, and polyenynes at 4.37 eV. Prasad reported the first experimental determination of the third-order nonlinear optical susceptibilities of some cumulenes in 1993. The |γ| values increase dramatically with increasing of the numbers of
Figure 38: DSC thermogram of polymer 36
Figure 39: UV/Vis absorption spectra of polymer 36 in THF solution and in the solid state.
cumulene units. The $\langle \gamma \rangle$ values are comparative to those of polyacetylenes and polydiacetylenes. In 1995, our group first reported the synthesis and study of cumulene-containing polymers (Scheme 29). It was found that the second hyperpolarizability $\langle \gamma \rangle$ measured for 38 is about two orders of magnitude higher than the cumulene molecule with the highest second hyperpolarizability ($\langle \gamma \rangle = 10^{31}$ esu) reported by Prasad. Replacement of the phenyl in the main chain of polymer 38 with the silole unit should dramatically enhance nonlinear optical property because polysilole shows the largest $\langle \gamma \rangle$ value according to the calculation. The route (Scheme 30) to polymer 41, containing silole and butatriene units, was similar to that for polymer 38.

Several routes were tried to synthesize monomer 39 (Scheme 31). Direct coupling between 2,5-bis(tributylstannyl)silole 42 and benzoyl chloride in the presence of Pd(PhCH$_2$)(PPh$_3$)$_2$Cl in THF for 3 days gave no reaction. Steric hindrance may be responsible for this failure. When 2,5-dibromosilole 24 was treated with 2 equivalents of n-BuLi, then quenched by either benzonitrile or benzoyl chloride, none of the desired product was isolated in either case. The possible reason for the failure is that 2,5-dibromosilole 24 could not be completely converted into 2,5-dilithiosilole by treatment with 2 equivalents of n-BuLi. However, 2,5-dilithiosilole, generated directly from diethynylsilane and 4 equivalents of lithium naphthalenide according to the literature, was quenched by benzoyl chloride to give the desired product 39 in 30% yield after purification by column chromatography several times.

However, polymerization of 39 with dilithium acetylide failed. The reaction of monomer 39 and dilithium acetylene yielded a red/black solid with broad $^{13}$C-NMR peaks at 127-128 ppm. There is no peaks observed for acetylenic carbons. Further GPC analysis showed the molecular weight was less 1000, indicating no polymerization. A possible explanation is the charge transfer process where one electron was transferred from the dilithium acetylene to silole. This possible charge transfer process was also seen in the attempted preparation of polycumulene carried in our group by using benzil or diketone acetylene and dilithium acetylide (Scheme 32).
Scheme 29: Synthesis of polymer containing butatriene unit

Scheme 30: Proposed route to poly(2,5-silole-1,2,3-butatriene)
Scheme 31: Synthesis of monomer 39

Scheme 32: Attempted synthesis of polycumulene\(^{68(b)}\)
Attempted Synthesis of Poly(2.5-silo-1.2-vinylene)

Poly(3,4-diphenyl-1,1-dihexyl-2,5-silo-1,2-vinylene) 43, which has alternate double bonds on the backbone, is a polyacetylene derivative. Recently, theoretical calculations have shown that $\sigma^*-\pi^*$ conjugation, between a $\pi$-symmetry $\sigma^*$ orbital of exocyclic $\sigma$ bond on silicon and a $\pi^*$ orbital of the butadiene skeleton, lowers the LUMO level, thus further decreasing the bandgap without decreasing the stability, and polysilo has the largest second hyperpolarizabilities. Furthermore, the dialkylsilylene group can provide long side chains to increase the solubility of the polymer. Poly(3,4-diphenyl-1,1-dihexyl-2,5-silo-1,2-vinylene) 43 should exhibit high conductivity and high second order hyperpolarizibility like polyacetylene. The synthesis and study of 43 should be of particular interest. The synthesis of polymer 43 was tried by direct coupling between 1,2-bis(tributylstanny)ethylene and dibromosilo 24 in the presence of palladium catalyst. Thus, the mixture of catalytic Pd(PPh$_3$)$_2$Cl$_2$, bis(tributylstanny)ethylene and dibromosilo 24 in THF was refluxed for more than 7 days, resulting in a black solution (Scheme 33). After the solvent was removed, the polymer could not be precipitated from MeOH. Only a sticky liquid, which was hard to dry and further purify and characterize, was isolated. The molecular weight of 43 was $\sim$1000 measured by GPC. The number of repeating units is $\sim$3-4. The low reactivity of dibromosilo 24 and 1,2-bis(tributylstanny)ethylene was the main reason. The UV/Vis absorption maximum of its THF solution was only at 420 nm, which is far below the UV/Vis absorption maximum of polyacetylene (495 nm), poly(3,4-dialkyl-2,5-thiophene-vinylene) (550 nm), and even PPV (459 nm). Changing catalyst from Pd(PPh$_3$)$_2$Cl$_2$ to Pd(PPh$_3$)$_4$ or to bis[1,2-bis(diphenylphosphino)ethane]-palladium (0) gave similar result. The low reactivity of 24 is the main reason for the failure. Use of compound 2,5-diido-silo as a starting material should give polymer 43 with high molecular weight. However, the further attempts to polymer 43 were not carried out.
Scheme 33: Attempted synthesis of poly(2,5-silole-vinylene) 43

Conclusions

Polysilole attract scientist recently because their theoretical calculations have shown their special electronic property and their attractive applications. Despite the fact that polysilole has not been synthesized, the silole-containing polymers were synthesized and exhibited unique properties. Silole unit can enhance the \( \pi \)-electron overlap along the main chain, lowering the LUMO energy. The values of the UV/Vis absorption maxima of these polymers are always much longer than those of their thiophene and phenyl analogues. The polymers were not significantly photoluminescent and did not dope to high electroconductivity.

Experimental

\(^1\)H and \(^13\)C-NMR spectra were acquired on a Varian VXR-300 spectrometer. In order to assure the quantitative features of the \(^13\)C-NMR spectra, the relaxation agent chromium(III) acetylacetonate was used in CDCl\(_3\) with a relaxation delay of 5 seconds.
Routine GC-MS spectra were obtained on a Hewlett Packard 5970 GC-IR-MS spectrometer at 70 eV. The exact masses were obtained from a Kratos MS 50 mass spectrometer with 10,000 resolution. The infrared spectra were recorded on a Bio-Rad Digilab FTS-7 spectrometer from neat sample. The UV/Vis spectra were obtained on a Hewlett Packard 8452A diode array UV/VIS spectrometer and $\lambda_{\text{max}}$ were determined at optical densities of 0.2-0.5.

Polymer molecular weights were determined by gel permeation chromatography (GPC) with 6 Microstyrage1 columns in series of 500 A, 2 x $10^3$ A, 2 x $10^4$ A, 2 x $10^5$ A. THF was used as an eluent at a flow rate of 1 ml/min. The system was calibrated by polystyrene standards. GPC analyses were performed on a Perkin-Elmer series 601 LC equipped with Beckman solvent delivery system, a Walter Associate R401 refractive index detector and a Viscotek viscometer. Differential scanning calorimetry (DSC) analyses were performed on a Du Pont 910 Differential Scanning Calorimeter. The UV/Vis spectra were measured on a Hewlett Packard 8452A Diode Spectrophotometer. Photoluminescence spectra were measured on a FL 900 fluorometer made by Edinburgh.

Toluene and benzene were distilled over CaH$_2$. THF was distilled from sodium/benzophenone before use. Other reagents were used as received from Aldrich Chemical Co. without further purification unless specified otherwise. Acetic acid, isopropanol and methanol were used as received from Fisher without further purification. All the reactions were performed under argon atmosphere. Diethynyldihexylsilane was synthesized by reaction of ethynylmagnesium bromide and dichlorodihexylsilane. 2-Trimethylstannylthiophene was synthesized by reaction of 2-bromothiophene with n-BuLi, followed by quenching with chlorotrimethyltin. Bis(tributylstannyl)acetylene or 1,4-bis(tributylstannyl)butadiyne was synthesized by reaction of bis(trimethylsilyl)acetylene or 1,4-bis(trimethylsilyl)butadiyne with (tributyltin) oxide in the presence of catalyst tetrabutylammonium fluoride (TBAF). $^{70}$ 1,2-Bis(tributylstannyl)ethylene was synthesized by hydrostannylation of tributylstannylacetylene with tributyl hydride.$^{71}$ 1,4-Bis(trimethylstannylethynyl)benzene $^{34}$ was prepared according to the literature by using 1,4-diethynyldibenzene, 2 equivalents of n-BuLi and trimethylchlorotin.$^{79}$
Bis(phenylethynyl)dihexylsilane: To a degassed solution of iodobenzene (20.4 g, 0.1 mol) and diethynyldihexylsilane (12.4 g, 50 mol) in Et3N (30 ml) was added Pd(PPh3)2Cl2 (350 mg) and CuI (140 mg) at room temperature. The mixture was stirred overnight. The salt was removed by filtration and was washed by hexane. After the solvent was removed, the product was purified by column chromatography (hexane as an eluent) (16.5 g, yield 82.5%). Mass: cal. m/z for C28H60Si = 424.86999, measure (HiRes EI) = 424.86987; 1H-NMR (300 MHz, CDCl3): δ 7.539-7.500 (m, 6H), 7.346-7.293 (m, 4H), 1.636-1.278 (m, 16H), 0.950-0.870 (m, 10H); 13C-NMR (300 MHz, CDCl3): δ 132.09, 132.03, 128.07, 128.13, 122.83, 106.53, 89.44, 32.69, 31.50, 23.66, 22.58, 14.81, 14.11

2,5-Bibromo-1,1-dihexyl-3,4-diphenylsilole 2d: The procedure was performed according to the literature.6 A solution of lithium naphthalenide was prepared by stirring a mixture of naphthalene (7.68 g, 60 mmol) and lithium powder (0.420 g, 60 mmol) in THF (100 ml) for 3 hours at room temperature under argon atmosphere. To the solution of lithium naphthalenide was added a solution of bis(phenylethynyl)dihexylsilane (6.000 g, 15 mmol) in THF (50 ml) dropwise at room temperature, and then the mixture was stirred for 15 minutes. The resulting 2, 5-dilithiosilole solution was cooled to -78°C, followed by dropwise addition of a solution of chlorotriphenylsilane (8.835 g, 30 mmol) in THF (50 ml) in order to quench the excess lithium naphthalenide. After stirring for 20 minutes, bromine (4.800 g, 30 mol) was added dropwise to the reaction mixture over 20 minutes at -78°C. The resulting yellow-white suspension was gradually warmed up to room temperature and was stirred at room temperature for overnight. After removal of salts and most part of the hexaphenyl disilane produced by filtration, the filtrate was concentrated, followed by addition of saturated aqueous Na2S2O3. The mixture was extracted with ether (200 ml x 2). The combined extract was washed with brine, dried over MgSO4, filtered and evaporated. The residue was heated to 55°C in vacuum (0.5 mmHg) to remove the most naphthalene by sublimation. The residue was purified by column chromatography (hexane as an eluent), followed by crystallization from hexane (5.000 g, yield 60%). m.p. 63°C (lit.60 64°C); Mass: cal. m/z for C28H60·79Br81BrSi
= 584.54951, C_{28}H_{60}^{81}Br_2Si = 586.54760, C_{28}H_{60}^{79}Br_2Si = 582.55142, measured (HiRes EI) = 584.467772, 586.54723, 582.55160, respectively; 'H-NMR: δ 7.200-7.100 (m, 10H), 6.963-6.932 (m, 16H), 1.600-1.320 (m, 16H), 1.015 (s, J = 8.4 Hz, 4H), 0.912 (m, J = 6.6 Hz, 6H); 
^{13}C-NMR: δ 156.81, 137.17, 128.94, 127.50, 127.34, 121.83, 32.68, 31.41, 22.86, 22.57, 14.13, 9.70

1,1-Dihexyl-3,4-diphenyl-2,5-bis(thienyl)silole 27: To a mixture of 2,5-dibromo-1,1-dihexyl-3,4-diphenylsilole 24 (0.224 g, 0.4 mmol) and 2-trimethylstannyl-thiophene (0.296 g, 1.2 mmol) in THF (15 ml) was added the catalyst Pd(PPh_{3})_{2}Cl_{2} (28 mg, 0.04 mmol). The solution was refluxed for 2 days. After the solvent was removed, the residue was purified by column chromatography (hexane as an eluent) (0.102 g, yield 37%). Mass: cal. m/z for C_{36}H_{44}S_{2}Si = 568.96296, measure (HiRes EI) = 568.96270; 'H-NMR: δ 7.05-7.21 (m, 6H), 6.97-7.00 (m, 4H), 6.90 (d, J = 6.0 Hz, 2H), 6.88 (d, J = 3.0 Hz, 2H), 1.19-1.55 (m, 20H), 0.85 (t, J = 6.0 Hz, 6H); ^{13}C-NMR: δ 153.79, 143.13, 139.37, 129.62, 128.39, 127.07, 126.85, 126.14, 125.50; UV/Vis: λ_{max} (THF) = 418 nm

Poly(2,5-thiophene-1,1-dihexyl-3,4-diphenyl-2,5-silole) 26: To a solution of 2,5-dibromo-1,1-dihexyl-3,4-diphenylsilole 24 (1.038 g, 1.854 mmol) and 2,5-bis(trimethylstannyl)thiophene 25 (0.760 g, 1.854 mmol) in THF (25 ml) was added Pd(PPh_{3})_{2}Cl_{2} (66 mg). The solution was refluxed for 7 days. After the solvent was removed, the polymer was precipitated by adding the polymer solution in THF into methanol (0.249 g, yield 27.8%). 'H-NMR: δ 6.90-7.36(br, 12H), 1.19 (br, 20H), 0.81 (br, 6H); ^{13}C-NMR: δ 153.66, 143.33, 141.39, 129.47, 128.34, 127.22, 32.57, 31.52, 23.53, 22.67, 14.16; GPC: M_{w} = 2.19 \times 10^{3}, M_{a} = 1.61 \times 10^{3}, PD = 1.360; UV/Vis: λ_{max} (THF) = 590 nm, λ_{max} (film) = 620 nm; Emission: λ_{max} = 631 nm, 674 nm, 693 nm (excitation at 510 nm); Electrical conductivity: 4.3 \times 10^{-5} S/cm (doped by I_{2})

Poly(1,1-dihexyl-3,4-diphenyl-2,5-silole-ethynylene) 28: To a solution of 2,5-dibromo-1,1-dihexyl-3,4-diphenylsilole 24 (0.527 g, 0.94 mmol) and bis(tributylstannyl)-
acetylene (0.569 g, 0.94 mmol) in THF (10 ml) was added Pd(PPh₃)₂Cl₂ (70 mg). The mixture was refluxed for 36 hours. After THF was removed, MeOH was added to precipitate the product. The precipitate was collected by filtration and then dissolved in small amount of THF. The polymer was precipitate again by adding MeOH (0.37 g, yield 93%). The polymer is very deep blue with metallic sheer. 'H-NMR: δ 7.259, 7.108, 6.997, 1.239, 0.872; ¹³C-NMR: δ 160.78, 138.11, 131.24, 128.93, 126.97, 124.63, 102.82, 32.58, 31.19, 22.87, 22.48, 14.00, 10.64; GPC: Mₘ = 6.52 x 10⁴, Mₕ = 1.81 x 10³, PD = 3.600; UV/Vis: λ_max (THF) = 604 nm; Emission: 640 nm and 655 nm (excitation at 550 nm); Electrical conductivity: 2.3 x 10⁻² S/cm (doped by I₂)

2,5-Di(trimethylsilylethynyl)-3,4-diphenylsilole 31: To a degassed solution of 2,5-dibromo-1,1-dihexyl-3,4-diphenylsilole 2₄ (1.147 g, 2.05 mmol) and Pd(PPh₃)₂Cl₂ (0.230 g)/CuI (0.057 g) in Et₂NH (10 ml)/benzene (10 ml) was added trimethylsilylacetylene (0.602 g, 6.14 mmol). After the mixture was stirred for 4 hours at room temperature, no salt was precipitated. Therefore, 1,8-diazabicyclo-[5,4,0]-undec-7-ene (DBU) (0.685 g, 4.51 mmol) was added into the mixture. The white precipitate was formed immediately. The mixture was stirred for 6 hours. The resulting salt was removed by filtration. After the solvent was removed, the residue was purified by column chromatography (hexane/ethyl acetate = 10:1 as eluents) to afford 31 (0.700 g, yield 57.0%). m.p. 65°C; Mass: cal. m/z for C₃₅H₇₆Si₃ = 619.29382, measure (HiRes EI = 619.29476; 'H-NMR: δ 7.24 (br, 8H), 1.74 (m, 4H), 1.57 (m, 4H), 1.48 (m, 8H), 1.18 (t, J = 6 Hz, 4H), 1.05 (t, J = 4 Hz, 6H), 0.25 (s, 18H); ¹³C-NMR: δ 162.82, 137.42, 129.25, 127.34, 127.04, 123.40, 105.64, 105.12, 32.82, 31.52, 23.21, 22.57, 14.19, 10.74, 0.00

3-Dihexylmethoxysilyl-4,5-diphenyl-octa-3,5-dien-1,7-diyne 32: To a solution of 2,5-bis(trimethylsilylethynyl)-1,1-dihexyl-3,4-diphenylsilole 2₄ (0.624 g, 1.05 mmol) in MeOH (15 ml)/THF (15 ml) was added a catalytic amount of KOH. The mixture was stirred at room temperature. The solvent was removed and the residue was purified by column chromatography to afford 32 (0.320 g, yield 63.2%). m.p. 60°C; Mass: cal. m/z for C₃₅H₆₅OSi
Poly(1,1-dihexyl-3,4-diphenyl-2,5-silole-ethynylene-1, 4-phenylene-ethynylene)

To a degassed solution of 2,5-dibromo-1,1-dihexyl-3,4-diphenylsilole 24 (0.200 g, 0.443 mmol) and 1,4-bis(trimethylstannylethynyl)benzene 34 (0.248 g, 0.443 mmol) in THF (10 ml) was added Pd(PPh3)2Cl2 (18 mg). The mixture was stirred at 50°C for 7 days. After removal of the solvent, the polymer was precipitated by adding the saturated THF solution of the polymer into MeOH (0.188 g, yield 81.0%). 1H-NMR: δ 7.41, 7.21, 6.97, 1.63, 1.45, 1.31, 0.87; 13C-NMR: δ 162.40, 137.40, 131.02, 129.06, 126.98, 123.88, 99.19, 92.62, 32.59, 31.28, 23.16, 22.35, 13.93, 10.66; GPC: Mw = 7.15 x 10^3, Mn = 2.32 x 10^3, PD = 3.082; UV/Vis: λmax (THF) = 494 nm, λmax (film) = 526 nm; Emission: 549 nm (excitation at 430 nm); Electrical conductivity: 4.3 x 10⁻³ S/cm (doped by I2)

Bis(trimethylstannylethynyl)-dimethylsilane 35: To a solution of diethynyldimethylsilane (4.310 g, 40 mmol) in THF (80 ml) was added EtMgCl (87.7 mmol, 43.8 ml, 2.0 M in THF) at -78°C dropwise. The mixture was allowed to warm to room temperature, and then refluxed for 2 hours. The solution was again cooled to -78°C. A solution of chlorotrimethyltin (87.7 mmol, 87.7 ml, 1.0 M in hexane) was added dropwise. The mixture was allowed to warm to room temperature and stirred overnight. A saturated NH4Cl aqueous solution (100 ml) was added. The aqueous layer was extracted by ether (100 ml x 2). The combined organic layer was washed with water, dried with MgSO4. After the solvent was removed, the product was purified by crystallization from hexane (10.600 g, yield 61.2%). Mass: cal. m/z for C12H24SiSn2 = 433.82806, measure (GC-MS): 415.85 (41.47%) (M⁺-15), 416.85 (36.60%) (M⁺-15), 417.85 (86.09%) (M⁺-15), 421.85 (80.29%) (M⁺-15), 420.85 (46.95%) (M⁺-15), 419.85 (95.45%) (M⁺-15), 418.85 (58.56%) (M⁺-15), 162.35
(24.25%), 163.35 (75.13%), 164.35 (28.24%), 165.35 (100%), 167.35 (15.62%); $^1$H-NMR: δ 0.28 (s, 6H), 0.25 (s, 18H); $^{13}$C-NMR: δ 114.97, 113.62, 0.89, -7.68

Poly(l,l-dihexyl-3,4-diphenyl-2,5-siloIe-ethynyIene-dimethylsilyIene-ethynyIene) 36: To a degassed solution of 2,5-dibromo-l,l-dihexyl-3,4-diphenylsilole 24 (1.074 g, 1.917 mmol) and bis(tributylstannylethynyl)dimethylsilane 25 (0.831 g, 1.917 mmol) in THF (25 ml) was added Pd(PPh$_3$)$_2$Cl$_2$ (66 mg). The mixture was refluxed for 7 days. After the solvent was removed, the polymer was precipitated by adding the saturated THF solution of the polymer into MeOH to afford a deep brown solid (0.501 g, yield 51.6%). $^1$H-NMR: δ 7.09 (br, 10H), 1.55, 1.38, 1.01, 0.87, 0.16; $^{13}$C-NMR: δ 162.86, 136.78, 128.79, 126.66, 122.92, 105.43, 101.70, 32.39, 31.03, 22.75, 22.16, 13.75, 10.29, 0.10; GPC: $M_n$ = 5.39 x 10$^3$, $M_a$ = 4.77 x 10$^3$, PD = 1.129; UV/Vis: $\lambda_{\text{max}}$(THF) = 400 nm, $\lambda_{\text{max}}$(film) = 426 nm; Electrical conductivity: 4.6 x 10$^{-6}$ S/cm (undoped), 5.4 x 10$^{-5}$ S/cm (doped by I$_2$)

2,5-Bisbenzoyl-l,l-dihexyl-3,4-diphenylsilole 38: 1, 1-Dihexyl-3,4-diphenyl-2,5-dilithiosilole was prepared from bis(phenylethynyl)dihexylsilane (4.000 g, 10 mmol) and lithium naphthalenide (40 mmol) in THF (70 ml) after excess lithium naphthalenide was quenched by triphenylchloride (5.900 g, 20 mmol) as described in preparation of 2,5-dibromo-l,l-dihexyl-3,4-diphenylsilole 24. Then the mixture was cooled to -78°C. Benzoyl chloride (2.810 g, 20 mmol) was added dropwise. The mixture was allowed to warm to room temperature and was stirred overnight. The resulting white precipitate was removed by filtration. A NH$_4$Cl aqueous solution was added. The aqueous layer was extracted by Et$_2$O (100 ml x 2). The combined organic layer was dried with MgSO$_4$. After the solvent was removed, the naphthalene was removed by vacuum sublimation at 55°C. The residue was purified by column chromatography 4 times to afford a yellow solid (1.830 g, yield 30%). m.p. 67°C; Mass: cal. m/z for C$_{47}$H$_{70}$OSi = 635.10220, measure (HiRes EI) = 635.10271; $^1$H-NMR: δ 7.66 (dd, J = 9 and 3 Hz, 4H), 7.30 (m, 2H), 7.18 (m, 6H), 6.89 (d, J = 1, 4H), 6.85 (dd, J = 15 and 3 Hz, 4H), 1.40 (m, 4H), 1.26-1.07 (m, 16H), 0.81 (t, J = 6 Hz, 6H); $^{13}$C-
NMR: $\delta$ 199.29, 156.74, 145.62, 137.08, 132.43, 129.03, 127.91, 127.49, 127.26, 32.90, 31.20, 23.63, 22.39, 4.01, 11.36
Conducting polymers are at present intensively studied because of their potential technological application.\textsuperscript{1,2} The modern era of conducting polymers began at the end of the 1970s when Heeger and MacDiarmid discovered that polydiacetylene, synthesized by Shirakawa's method,\textsuperscript{80} could increase a 12 order of magnitude of conductivity upon charge-transfer oxidative doping.\textsuperscript{81} However, polyacetylenes are environmental unstable, which constitutes a major obstacle to practical applications. Recent interest in conducting polymers have shifted to poly(heterocycles) because of their higher environmental stability and structure versatility which allows the modulation of their electronic and electrochemical properties by manipulation of the monomer structure.

An important step in the development of conjugated poly(heterocycles) occurred in 1979 when it was shown that highly conducting and homogeneous free-standing films of poly(pyrrole) could be produced by oxidative electropolymerization of pyrrole.\textsuperscript{82} Since then, the electrochemical polymerization has been rapidly extended to other aromatic compounds such as thiophene,\textsuperscript{83,84} furan,\textsuperscript{84} indole, carbazole, azulene, pyrene,\textsuperscript{85} benzene,\textsuperscript{86} fluorene,\textsuperscript{87} and aniline. Among them, polythiophenes have rapidly become a "hot subject" for their potential applications. These applications include rechargeable battery electrodes, electrochromic devices,\textsuperscript{88(a,b)} chemical and optical sensors,\textsuperscript{88(c)} light-emitting diodes,\textsuperscript{24(c),88(d)} molecular-based devices,\textsuperscript{88(e,f)} and nonlinear optical active materials.\textsuperscript{10,88(g,h)}

Polythiophenes have already been considered in several reviews devoted to conducting polymers in general or to some of their more specific aspects such as electropolymerization,\textsuperscript{89} electrochemistry,\textsuperscript{90} environmental stability,\textsuperscript{91} or optical properties.\textsuperscript{88(a)} There are two reviews which focused only on polythiophenes in 1986\textsuperscript{9(a)} and 1992.\textsuperscript{5(a)} However, considerable progress has been accomplished in their preparation and characterization of their structures.
and electronic and electrochemical properties, and the essential part of the work on functional polythiophenes and their practical applications in the past several years.

Polythiophenes are usually prepared by either electrochemical polymerization or chemical polymerization. The chemical polymerization is usually carried out in an electric cell. Noble metals such as platinum and gold, and optically transparent tin oxide or indium-tin oxide (ITO) are used as electrodes. The solvents employed are anhydrous aprotic solvents of high dielectric constants and low nucleophilicities, such as acetonitrile, benzonitrile, nitrobenzene and propylene carbonate. These solvents lead to highest current efficiency of electropolymerization. The electrolytes are anions, derived from strong acids such as ClO$_4^-$, PF$_6^-$, BF$_4^-$ and AsF$_6^-$, associated with lithium or tetraalkylammonium cations. The solution is prepared by dissolving the monomer and electrolyte in the solvent in about 0.1-0.5 M concentration in a three-electrode cell (third electrode as a reference) and degassed. Polythiophenes are generally grown on the electrode in potential static or galvanostatic conditions and by recurrent potential sweeps, or current pulses.

The chemical syntheses are the most adequate methods of preparation of oligomers of defined structures. Several methods have been developed for preparation of polythiophenes (Scheme 34). Ni-catalyzed polycondensation of 2-bromo-5-(bromozincio)-thiophene$^{92}$ or 5-(2-bromo-thienyl)-magnesium bromide,$^{93}$ generated from either oxidative addition of Zn or Mg to 2,5-dibromothiophene or metathesis of the corresponding organolithium reagents with ZnBr$_2$ or MgBr$_2$, and Pd-catalyzed polycondensation of 2,5-diiodothiophene with 2,5-distannylthiophene, are the most used methods. Oxidative coupling of 2,5-dilithiothiophene in the presence of Fe(acac)$_3$ or CuCl$_2$,$^{94}$ is much simpler than Ni or Pd-catalyzed coupling, but gives lower molecular polymers (Scheme 34). Oxidative coupling of 2,5-dilithiothiophene and Ni or Pd-catalyzed coupling can ensure the 2,5-linkage necessary for conjugation. However, all these methods require extremely pure monomers and can not tolerate electrophilic functional groups. Recently, direct slow oxidative polymerization of thiophene using FeCl$_3$ as an oxidant$^{95}$ was developed and widely used for preparation of polythiophenes because it doesn't need pure monomers, is easy and suitable for large scale production, and gives high molecular weight polymers (Scheme 34). But it also result in conjugation-breaking α-β'
In order to synthesize polythiophene with electron-withdrawing substituents directly attached to the thiophene and with substituents having electrophilic groups such as acids, esters and carbonyls, several novel methods were developed recently. In 1995, polythiophenes with a carbonyl group directly attached to the ring were prepared via an Ullman reaction by refluxing 2,5-dibromothiophene with activated Cu in DMF (Scheme 34). The other method is oxidative coupling of 2,5-bis(chloromercuro)-3-alkylthiophene by Cu and PdCl₂. These two methods, however, gave regiorandom polymers when 3-substituted thiophene was used as a monomer, which may result in an out-of-plane conformation for the units along the polymer chain.

Scheme 34: Chemical preparation of polythiophenes
A variety of polythiophenes with functionalized substituents have been synthesized. The functional groups can be introduced into the 3- or 4-position to increase the solubility of the polymer and provide special properties to polythiophene. The most studied polythiophenes are the 3-substituted polythiophenes (Figure 40). 3-Substituted polythiophenes display properties superior to those of 3,4-disubstituted polythiophenes because of their less steric effects between substituents grafted on adjacent thiophene. An ethereal or thioetheral side chain allows the fine-tuning of electrical and optical properties of polythiophene. A crown ether side chain results in ionochromic activity in the polythiophene. A chiral side chain yields an optically active polythiophene. An alkanesulfonic acid side chain produces a water-soluble self-doping conducting polymer.

Fused ring systems (Figure 41) have also been prepared, mainly by electrochemical polymerization. These ring systems, which have a bridge to enhance planar structures, have lower bandgaps than 3-substituted polythiophenes. Poly(isothianapthene) has been shown to

![Figure 40: Examples of 3-substituted polythiophene](image-url)
Results and Discussions

Synthesis and Study of Silicon-Bridged Polythiophene

Despite the fact that a variety of polythiophenes have been synthesized, polythiophenes containing silicon have not been reported. As discussed in part II, a silicon unit can dramatically change the electronic structure of the polymers. The $\sigma^*$ orbital on silicon interacts with $\pi$-conjugated orbitals, thus decreasing the LUMO level energy dramatically.
We designed polythiophene 51 having a bridging silicon unit (Figure 42). The bridge, dialkylsilylene, not only provides two big alkyl groups to increase the solubility, and enhances the planar structure of the polymer, but also fully interacts with the π-orbitals, decreasing the bandgap energy. The polymer structure can also be viewed as containing a silole unit. This silicon-bridged polythiophene should have a very low bandgap, and therefore high conductivity and other special properties.

![Figure 42: The design of silicon bridged polythiophene](image)

**Synthesis and Characterization.** The key compound leading to the silicon-bridged polythiophene 51 was designed as silicon-fused bithiophene 48. First, 3,3'-dibromo-2,2'-bithiophene 44 was synthesized starting with 2-bromothiophene (Scheme 35) according the literature.105

![Scheme 35: Synthesis of 3,3'-dibromo-2,2'-bithiophene 44](image)
In 1955, Gilman developed a method to synthesize 9-silafluorenes 45 by a reaction of 2,2'-dilithiobiphenyl, generated from 2,2'-dibromobiphenyl and 2 equivalents of n-BuLi, with RₙSiQₙ (n = 2-4) (Scheme 36). The reaction was carried out by slow addition of a 2,2'-dilithiobiphenyl solution into a refluxing RₙSiQₙ ether solution. When dialkylchlorosilanes (R₂SiCl₂) were used, the yields were very low (18-25%) with byproduct 46 being formed. When RSiCl₃ or SiCl₄ was used, the yields were improved to higher than 70%, and the product 9-chloroalkyl- (or dichloro-)silafluorene could be transformed into 9-dialkyllsilafuorene 45 by reacting it with alkylithium or Gringard reagents.

Scheme 36: Synthesis of 9-silafluorene

The synthesis of 3,3'-dihexylsilylene-2,2'-bithiophene 47 was first tried by using Gilman's procedure. A 3,3'-dilithio-2,2'-bithiophene solution, generated from 3,3'-dibromo-2,2'-bithiophene 44 and 2 equivalents of n-BuLi at -78°C, was added into a refluxing dichlorodihexylsilane solution. After work-up, the yield was only 10%. We rationalized that 3,3'-dilithio-2,2'-bithiophene was not stable at high temperature. It would easily undergo metal-hydrogen exchange to give more stable 5,5'-dilithio-2,2'-bithiophene. The reaction condition was then modified by directly adding dichlorodihexylsilane into a 3,3'-dilithio-2,2'-bithiophene solution at -78°C and keeping the solution stirring at -78°C for 5 hours (Scheme 37). After then, the reaction mixture was allowed to warm to room temperature and stirred at room temperature overnight. The yield of 47 was improved up to 81.2%. 

![Scheme 36: Synthesis of 9-silafluorene](image-url)
Compound 47 was then converted to 5,5'-dibromo-3,3'-dihexylsilylene-2,2'-bithiophene 48 by reacting it with 2 equivalents of NBS in DMF in 90% yield (Scheme 37). It was important that the solution was stirred at room temperature for less than 20 minutes. The longer stirring time resulted in the desilated product.

Scheme 37: Synthesis of 3,3'-dihexylsilylene-2,2'-bithiophene 47 and 5,5'-dibromo-3,3'-dihexylsilylene-2,2'-bithiophene 48

Compounds 47 and 48 have blue luminescence in THF solution. The UV/Vis absorption maximum ($\lambda_{\text{max}}$) of 47 is 338 nm, which is much larger than those of bithiophenes, such as 2,2'-bithiophene (302 nm), 3,4'-dimethyl-2,2'-bithiophene (302 nm) and 3,3'-methylene-2,2'-bithiophene (312 nm), but still smaller than those of dithieno[3,2-b:2',3'-d]thiophene (346 nm) and cyclopenta[2,1-b:3',4'-b']dithiophene-4-one (474 nm) (Figure 43). The $\lambda_{\text{max}}$ of 47 is 26 nm longer than that of its carbon analog 3,3'-methylenes-2,2'-bithiophene. The UV/Vis absorption maximum of 48 was 358 nm, 20 nm longer than that of 47. The $\lambda_{\text{max}}$ values of 47 and 48 show that silicon has more effect on electronic structure than carbon by dramatically lowering the LUMO energy.
Other evidence showing that a silicon group induces a bathochromic shift is found in a 1992 paper by Tour's group. Tour et al. synthesized and studied a series of trimethylsilyl-capped \( \alpha \)-oligothiophenes (Figure 44). It was found that the \( \lambda_{\text{max}} \) values of trimethylsilyl-capped \( \alpha \)-oligothiophenes were usually larger than those of non-capped \( \alpha \)-oligothiophenes. They rationalized that \( \pi \)-\( \pi \)-conjugation contributed to these differences. It is surprising that the \( \lambda_{\text{max}} \) (338 nm) of compound 47 is 18 nm longer than that of 5,5'-bis(trimethylsilyl)-2,2'-bithiophene (320 nm) because 5,5'-bis(trimethylsilyl)-2,2'-bithiophene has two silicon units while 47 has only one. This shows that the bridged silicon group enhances the planarity and participates in conjugation more effectively than a terminal silicon group.

Compound 48 was used as a monomer to couple with 2,5-bis(trimethylstannyl)-thiophene in the presence of catalytic \( \text{Pd(PPh}_3)_2\text{Cl}_2 \) to give a silicon-bridged polythiophene 49 (Scheme 38). The reaction was carried out in refluxing THF for 7 days. After the solvent was removed, polymer 49 was precipitated from MeOH to obtain a black metallic solid. Polymer 49 was very soluble in organic solvents such as CHCl\(_3\), CH\(_2\)Cl\(_2\), THF, and toluene,
and gave a reddish solution in THF. The polymer film, made from a THF solution by evaporating the solvent slowly, was very smooth and deep red with a metallic luster.

Polymer 49 was characterized by \(^1\)H- and \(^{13}\)C-NMR (Figure 45) and GPC. The \(^1\)H-NMR spectrum shows a broad peak at \(-7.10\) ppm for the protons of thiophene and two broad peaks at \(-1.29\) and \(-0.89\) ppm for the hexyl protons. The \(^{13}\)C-NMR spectrum shows exactly six peaks between 148 and 124 ppm corresponding to the six thiophene carbons (due to the symmetry of the polymer, only six thiophene carbons show the peaks in the repeating unit), and six hexyl carbons between 33 and 11 ppm. The six thiophene peaks are assigned to the six thiophene carbons as shown in Figure 45. The GPC chromatogram shows a main peak overlapping with a small shoulder, indicating that polymer 49 contains polymers of two different lengths. The weight molecular weight of polymer 49 is \(3.16 \times 10^4\) and PD is 3.21.
The synthesis of silicon-bridged polythiophene \textbf{49} was tried by different methods. Direct slow oxidative polymerization of 3,3'-dihexylsilylene-2,2'-bithiophene \textbf{47} with FeCl\textsubscript{3} in CHCl\textsubscript{3} (Scheme 33) gave a low molecular weight polymer in very low yield (10\%). The \textsuperscript{1}H- and \textsuperscript{13}C-NMR spectra showed that the polymer lost the silicon bridge. It is believed that polymerization of thiophene with FeCl\textsubscript{3} in CHCl\textsubscript{3} involves a radical or a carbocation intermediate and produces HCl gas as a byproduct. Under these conditions, the desilation could easily happen by either radicals (or carbocations) or HCl during the polymerization.

Ullman reaction by refluxing \textbf{48} with activated Cu powder in DMF for 6 days (Scheme 34) also gave a low molecular weight polymer. Oxidative coupling of 5,5'-dilithio-3,3'-dihexylsilylene-2,2'-bithiophene, generated from either \textbf{48} and 2 equivalents of n-BuLi or \textbf{47} and 2 equivalents of n-BuLi/TMEDA, with Fe(acac\textsubscript{3}) also caused the desilated product. The only good method is palladium-catalyzed coupling of monomer 5,5'-bis(trialkylstanny1)-3,3'-dihexylsilylene-2,2'-bithiophene \textbf{50} with 5,5'-dibromo-3,3'-dihexylsilylene-2,2'-bithiophene \textbf{48} (Scheme 35).
Figure 45: The $^1$H- and $^{13}$C-NMR spectra of polymer 49
However, the synthesis of monomer 5,5'-bis(trialkylstannyl)-3,3'-dihexylsilylene-2,2'-bithiophene 50 was painful. First, 5,5'-dibromo-3,3'-dihexylsilylene-2,2'-bithiophene 48 was reacted with 2 equivalents of n-BuLi, followed by quenching with excess Bu$_3$SnCl. TLC showed only one spot, which had almost the same $R_f$ as the debrominated compound 47. The product, after purification by column chromatography, was confirmed as 47 by NMR.

Second, direct metathesis of 47 with 2 equivalents of n-BuLi/TMEDA or excess LDA, followed by quenching with excess Bu$_3$SnCl, after purification by column chromatography, still failed to give the desired product 50 and gave the starting material 47 instead. Third, palladium-catalyzed reaction of 48 with excess hexabutylditin also gave the product 47 after column chromatography. After all above failures, we realized that the product 50 had the same $R_f$ as 47 and the destannylation could happen during the purification by column chromatography.

Thus compound 50 was assumed to be obtained in 100% yield by either direct metathesis of 47 with excess LDA or treatment of 48 with 2 equivalents of n-BuLi, followed by quenching with excess Me$_3$SnCl, and was used without further purification to couple with 48 in the presence of Pd(PPh$_3$)$_2$Cl$_2$ catalyst (Scheme 39). Indeed, after the mixture was refluxed in THF for 7 days, the desired polymer 51 was obtained in 87.4% yield. Polymer 51 is a black solid which is very soluble in common organic solvent such as THF, toluene, chloroform, and methylene chloride. The THF solution is purple with red luminescence.

Polymer 51 was characterized by $^1$H- and $^{13}$C-NMR and GPC. The $^1$H-NMR spectrum shows only one peak at 7.10 ppm for the thiophene protons, and the hexyl protons between 1.24 and 0.85 ppm. The $^{13}$C-NMR spectrum shows clearly only four sp$^2$ thiophene carbons between 147.16 and 125.68 ppm and six sp$^3$ hexyl carbons between 32.65 and 11.56 ppm (Figure 46). The GPC shows the molecular weight of polymer 51 is $2.03 \times 10^4$ (PD = 1.896).

**Thermal Behavior.** The thermal behavior of polymers 49 and 51 was studied by thermogravimetric analysis (TGA) and differential scanning calorimetric analysis (DSC) under argon atmosphere. The TGA was performed from room temperature to 800°C with a ramp of 15°C/min. From the TGA spectra (Figure 47), polymers 49 and 51 have similar thermal
Scheme 39: Synthesis of Silicon-bridged polythiophene 51

stabilities. They start to lose weight rapidly at about 430°C. Polymer 49 slowly loses weight after 600°C and has a 70% char yield at 800°C, while polymer 51 slowly loses weight after 532°C and has a 59% char yield at 800°C.

The thermal transitions of polymers 49 and 51 were studied by DSC. Polymer 49 has endothermic peaks at ~278°C and 497°C and an exothermic peak at ~302°C before it
Figure 46: The $^1$H- and $^{13}$C-NMR spectra of polymer 51
Figure 47: The TGA spectra of polymers 49 and 51
Table 11: Thermal property analysis of polymers 49 and 51

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<th>Polymer</th>
<th>TGA</th>
<th>DSC</th>
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<td></td>
<td>$T_d$ (°C) (4% weight lose)</td>
<td>Char Yield (at 800°C)</td>
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<td>49</td>
<td>430</td>
<td>80%</td>
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<tr>
<td>51</td>
<td>439</td>
<td>59%</td>
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starts to decompose. Polymer 51 has an endothermic peak at 334°C. The thermal results are summarized in Table 11.

**UV/Vis Absorption.** The UV/Vis absorption spectra of polymers 49 and 51 in THF and the solid state were measured. Both polymers exhibit unusually long absorption maximum wavelengths. The $\lambda_{max}$ values of polymers 49 and 51 in THF are 556 nm and 582 nm, respectively (Figure 48). These $\lambda_{max}$ values are the longest among polythiophenes known to us. Compared with the $\lambda_{max}$ of normal 3-substituted polythiophenes (448 nm for 3-octylpolythiophene, for example), the absorptions of polymers 49 and 51 are about 110 nm to 130 nm red-shifted. Compared with the $\lambda_{max}$ of 49 (in which the silylene bridge to thiophene ratio is 1/3), the $\lambda_{max}$ of 51 (in which the silylene bridge to thiophene ratio is 1/2) is 26 nm red-shifted. All these data show that the silicon on the main chain dramatically decreases the bandgap energy. A silicon unit can change the electronic structures of the polymers by interacting the $\sigma^*$ orbital of silicon with $\pi$-conjugated orbitals, thus decreasing the LUMO energy. The absorption edges ($\lambda_e$) of polymers 49 and 51 were estimated in order to calculate the bandgap energy. Polymer 49 has a $\lambda_e$ of 650 nm and the bandgap energy of polymer 49 can be estimated as 1.90 eV. Polymer 51 has a more extended absorption tail up to 720 nm compared with polymer 49 ($\lambda_e - \lambda_{max} = 92$ nm for polymer 49, while $\lambda_e - \lambda_{max} = 138$ nm for polymer 51). The bandgap of polymer 51 is estimated to be as low as 1.71 eV, one of the lowest bandgaps of polythiophenes ever reported.
The solid state UV/Vis absorptions of polymers 49 and 51 were also measured. These two polymers can form smooth "metallic" films. The UV/Vis absorptions of these films are similar to those of their THF solutions, but red-shifted. The film of polymer 49 has a $\lambda_{\text{max}}$ of 582 nm, 26 nm red-shifted compared with its THF solution. The film of polymer 51 has a $\lambda_{\text{max}}$ of 592 nm, only 10 nm red-shifted compared with that of its THF solution. The results of the UV/Vis absorptions of polymers 49 and 51 are summarized in Table 12.

![Figure 48: The UV/Vis absorptions of polymers 49 and 51](image)

**Table 12: The results of UV/Vis absorptions of polymers 49 and 51**

<table>
<thead>
<tr>
<th>Polymer</th>
<th>THF</th>
<th>Film</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>$\lambda_{\text{max}}$</td>
<td>$\lambda_e$</td>
</tr>
<tr>
<td>49</td>
<td>556 nm</td>
<td>650 nm (1.90 eV)</td>
</tr>
<tr>
<td>51</td>
<td>582 nm</td>
<td>720 nm (1.71 eV)</td>
</tr>
</tbody>
</table>
**Fluorescence.** Polymers 49 and 51 have weak reddish photoluminescence in THF. The fluorescence spectra of the polymers in THF were measured (Figure 49). Polymer 49 has two emission peaks at 627 and 662 nm when excited at 510 nm. These two emission peaks are probably due to two different lengths of polymers as shown in GPC (a main peak overlapping with a small shoulder). Polymer 51 has a weaker emission peak at 680 nm when excited at 540 nm.

![Fluorescence spectra](image)

**Figure 49:** The fluorescence spectra of polymers 49 (excited at 510 nm) and 51 (excited at 540 nm)

**Electrical Conductivity.** Polythiophenes are most often used as conducting polymers. The narrow bandgaps of the silicon-bridged polythiophenes suggest that they might be highly conductive upon doping. The shiny, deep purple films of polymers 49 and 51 were nonconductive in their neutral states. But exposure to I₂ vapor under vacuum changed the conductivities of the films dramatically with a few seconds, and turned them into black shiny films. The I₂-doped polymers exhibited a conductivity of 4-6 × 10² S/cm (determined by the
Table 13: Conductivities of I\textsubscript{2}-doped polymers \textit{49} and \textit{51}

<table>
<thead>
<tr>
<th>Polymer</th>
<th>I\textsubscript{2}-doped conductivity (S/cm)</th>
</tr>
</thead>
<tbody>
<tr>
<td>\textit{49}</td>
<td>~400</td>
</tr>
<tr>
<td>\textit{51}</td>
<td>~600</td>
</tr>
</tbody>
</table>

two-probe method) as listed in Table 13, one of the highest conductivities of polythiophenes (the highest conductivity is up to \(> 1,000 \text{ S/cm}\)).

\textit{Synthesis and Study of Butadiene-Linked Polythiophenes}

Polyacetylenes are the most conductive but environmentally unstable polymers. Polythiophenes, which increase the aromaticity along the main chains and the environmental stabilities, decrease the conductivity. We rationalized that decreasing the aromaticity of polythiophene by putting butadiene units into the polythiophene main chains will increase the mobility of \(\pi\)-electrons over the main chain, lowering the bandgap energy without significantly decreasing the environmental stability. Butadiene-linked polythiophenes, which are more like hybrids of polyacetylenes and polythiophenes, should exhibit special properties.

In 1985, Tanaka et al. synthesized polyene-linked polythiophenes \textit{52} by electrochemical polymerization (Scheme 40), and studied their UV/Vis absorptions and conductivities. However, the polymers they synthesized were insoluble and were only characterized by IR spectroscopy. The conductivity of these polymers were very low (\(\sigma = 2.8 \times 10^{-2} \text{ S/cm}\)).

\textit{Synthesis and Characterization.} One of the monomers for butadiene-linked polythiophenes is monomer \textit{55}, which was synthesized, starting from n-octylamine, as shown in Scheme 41. Phase transfer N-alkylation\textsuperscript{112} of n-octylamine with propargyl bromide yielded N,N-di(3-propynyl)-N-octylamine \textit{53} in 76\% yield. Palladium-catalyzed coupling of \textit{53} with 2-iodothiophene in Et\textsubscript{3}N gave product \textit{54} in 90\% yield. The zirconocene-mediated
Scheme 40: Synthesis of polyene-linked polythiopehnes by electrochemical polymerization

\[
\text{Scheme 41: Synthesis of monomer 55}
\]

intramolecular cyclization\(^{113}\) of 54, followed by quenching with acid, gave the desired monomer 55 in 80% yield.

The other monomer 57 has a butadiene linkage, but the silicon unit was introduced into the monomer in order to decrease the LUMO energy (Scheme 42). The Pd-catalyzed
coupling of diethynyldihexylsilane with two equivalents of 2-iodothiophene gave compound 56 in 87% yield. Treatment of 56 with Cp₂ZrEt₂, generated from Cp₂ZrCl₂ and 2 equivalents of EtMgBr, followed by quenching with acid, gave a silacyclobutene derivative 57 in 45% yield. This reaction was discovered by the Takahashi's group in 1995. The intermediate from a reaction of di(phenylethynyl)dialkylsilane and Cp₂ZrEt₂ was the fused ring compound 58, which was then quenched with acid to give silacyclobutene 59 (Scheme 43)."*"  

\[
\text{Hex Si} \quad \equiv \quad \equiv \\
\text{Hex} + \text{2 eq. I} 
\xrightarrow{\text{cat. } \text{Pd} \text{Et}_3\text{N}} 
\quad \text{Yeild 87\%} \\
\quad 56 
\xrightarrow{1) \text{Cp}_2\text{ZrEt}_2} 
\quad \text{H}^+ 
\quad \text{Yield 45\%} \\
\quad 57 
\]

Scheme 42: Synthesis of monomer 57

\[
\text{[Ph} \equiv \equiv \) \text{SiR}_2 \quad \xrightarrow{\text{Cp}_2\text{ZrEt}_2} 
\quad \text{Ph} 
\quad 58 
\quad \xrightarrow{\text{H}^+} 
\quad \text{Ph} 
\quad 59 
\]

Scheme 43: The formation of silacyclobutene"*

Compared with the UV/Vis absorption of terthiophene (\(\lambda_{\text{max}} = 353 \text{ nm in benzene}\)), the UV/Vis absorption maxima of 55 and 57 are about 17 nm (\(\lambda_{\text{max}} = 370 \text{ nm in THF with two shoulders at 354 nm and 390 nm for 55}\)) and 19 nm (\(\lambda_{\text{max}} = 372 \text{ nm in THF with two shoulders at 356 nm and 388 nm for 57}\)) red-shifted, respectively, indicating that the energies
between the LUMO and the HOMO of 55 and 57 are smaller than that of terthiophene. The \( \lambda_{\text{max}} \) of 57 is almost the same as that of 55 (only 2 nm red-shifted) even 57 has a silicon unit in the structure. The ring strain in 57 might offset the silicon effect. However, the silole-linked derivative of 57, TST (which has less ring strain), has a UV/Vis absorption maximum of 416 nm, much longer than those of other compounds (Figure 50).

![Chemical structures](image)

**Figure 50**: The comparison of the UV/Vis absorptions

Polymerization of monomers 55 and 57 was carried out by several methods. Bromination of 55 and 57 at the 5,5'-positions failed to give the desired 5,5'-dibromo-products. The direct metathesis of 55 and 57 at the 5,5'-positions with n-BuLi or LDA also failed. Due to the difficulties of introducing functional groups at the 5,5'-positions, the polymerization was carried out by direct oxidative coupling of the corresponding monomer using FeCl₃ in chloroform (Scheme 44). Polymer 60 was obtained in 46% yield. Polymer 60 was a deep red solid and very soluble in organic solvents. However, monomer 57 failed to give the desired polymer 61, as the desilation happened during the polymerization.
Polymer $\text{60}$ was characterized by $^1\text{H}$- and $^{13}\text{C}$-NMR (Figure 51) and GPC. The $^1\text{H}$-NMR spectrum shows two peaks at 7.70 and 7.52 ppm for the protons on thiophene, one peak at 7.01 ppm for the vinyl protons, one peak at 4.21 ppm for the protons of cyclic methylene, one peak at 3.74 ppm for the protons of the acyclic methylene attached to nitrogen, and several peaks for the alkyl protons between 1.67 and 0.88 ppm. The $^{13}\text{C}$-NMR spectrum shows clearly six sp$^2$ carbons between 140.00 and 111.78 ppm, two sp$^3$ carbons attached to nitrogen at 58.50 and 56.66 ppm, and seven sp$^3$ carbons between 38.73 and 10.97 ppm for octyl group. GPC shows polymer $\text{60}$ has $M_w$ of $1.39 \times 10^4$ with PD of 1.84.

**Thermal Behaviors.** The thermal behavior of polymer $\text{60}$ was studied by thermogravimetric analysis (TGA) and differential scanning calorimetric analysis (DSC) under argon atmosphere. Polymer $\text{60}$ started to lose weight (decompose) at $\sim 348^\circ\text{C}$ (3% weight loss) slowly up to $448^\circ\text{C}$ (11% weight loss). After $448^\circ\text{C}$, the polymer lost weight dramatically. At $531^\circ\text{C}$, the weight loss was $\sim 58\%$. At $800^\circ\text{C}$, the char yield was $37\%$ (Figure 52). DSC shows the polymer melted at $\sim 130^\circ\text{C}$ and crosslinking started at $\sim 243^\circ\text{C}$ to give a large exothermic peak at $306^\circ\text{C}$ (Figure 53).
Figure 51: The $^1$H- and $^{13}$C-NMR spectra of polymer 60
Figure 52: The TGA spectrum of polymer 60

Figure 53: The DSC thermogram of polymer 60
UV/Vis absorption. The UV/Vis absorption of polymer 60 was measured (Figure 54). The $\lambda_{\text{max}}$ of polymer 60 in THF solution is 514 nm, which is larger than that of usual 3-substituted polythiophene (448 nm for 3-octylpolythiophene, for example), but much smaller than those of silicon-bridged polythiophenes (see Table 10). The absorption edge ($\lambda_e$) of polymer 60 was estimated as 630 ± 5 nm. The bandgap can be estimated as about 1.96 eV.

In 1985, the butadiene-linked polythiophene 52 (m = 2) (Scheme 39) film was prepared by electrochemical polymerization. The neutral polymer 52 (m = 2) film was red with a $\lambda_{\text{max}}$ of only 390 nm. Polymer 60 can form a smooth reddish film by spin-coating technique. The UV/Vis absorption of the polymer 60 film was also measured. The $\lambda_{\text{max}}$ of the polymer 60 film is 534 nm, about 20 nm longer than that of its THF solution. Compared with the $\lambda_{\text{max}}$ of the polymer 52 (m = 2) film, the $\lambda_{\text{max}}$ of polymer 60 film is 144 nm longer.

![Figure 54: The UV/Vis absorption of polymer 60 in THF and the solid state](image)
Fluorescence. Polymer 60 has weak luminescence in THF solution. The fluorescence spectrum of its THF solution was measured as shown in Figure 55. Polymer 60 has an emission peak at 554 nm with a small shoulder at ~470 nm when excited at 440 nm.

Electrical Conductivity. The conductivity of polymer 60 was studied as comparison with that of the butadiene-linked polythiophene 52 (m = 2), prepared by electrochemical polymerization. The dark green film of polymer 52 (m = 2) had the highest conductivity of 0.5 S/cm when oxidized by Bu₃NClO₄. The polymer 60 film is nonconductive at its neutral state. After exposure to I₂ vapor under vacuum, the conductivity of the polymer 60 film increased dramatically up to σ of ~2 S/cm, which is 4 fold larger than that of polymer 52 (m = 2). However, it is still low considering its low bandgap energy.

Figure 55: The emission spectrum of polymer 60 in THF solution
Conclusions

Novel silicon-bridged and butadiene-linked polythiophenes were synthesized. Their thermal and optical properties were studied. The silicon-bridged polythiophenes exhibited very unusual optical properties. They have the longest UV/Vis absorption maximum wavelengths among the polythiophenes ever reported, which indicates they have the lowest bandgap energies. The bridged silicon can not only enhance the planarity of the polymers, but also participate in the conjugation along the main chains by interacting $\sigma^*$ orbital of silicon with $\pi$-conjugated orbitals of the backbones, thus decreasing the bandgap energies dramatically. The conductivities of the silicon-bridged polythiophenes were measured as high as ~400-600 S/cm, one of the highest conductivities among polythiophenes. The butadiene-linked polythiophene also showed that its UV/Vis absorption maximum wavelength was longer than that of normal 3-substituted polythiophene. However, the butadiene-linked polythiophene didn't exhibit high conductivity as expected.

Experimental

$^1$H and $^{13}$C-NMR spectra were acquired on a Varian VXR-300 spectrometer. In order to assure the quantitative features of the $^{13}$C-NMR spectra, the relaxation agent chromium(III) acetylacetonate was used in CDCl$_3$ with a relaxation delay of 5 seconds.

Routine GC-MS spectra were obtained on a Hewlett Packard 5970 GC-IR-MS spectrometer at 70 eV. The exact masses were obtained from a Kratos MS 50 mass spectrometer with 10,000 resolution. The infrared spectra were recorded on a Bio-Rad Digilab FTS-7 spectrometer from neat sample. The UV/Vis spectra were obtained on a Hewlett Packard 8452A diode array UV/VIS spectrometer and $\lambda_{max}$ were determined at optical densities of 0.2-0.5.

Polymer molecular weights were determined by gel permeation chromatography (GPC) with 6 Microstyragel columns in series of 500 A, 2 x $10^3$ A, 2 x $10^4$ A, 2 x $10^5$ A. THF was used as an eluent at a flow rate of 1 ml/min. The system was calibrated by polystyrene
standards. GPC analyses were performed on a Perkin-Elmer series 601 LC equipped with Beckman solvent delivery system, a Walter Associate R401 refractive index detector and a Viscotek viscometer. Differential scanning calorimetry (DSC) analyses were performed on a Du Pont 910 Differential Scanning Calorimeter. The UV/Vis spectra were measured on a Hewlett Packard 8452A Diode Spectrophotometer. Photoluminescence spectra were measured on a FL 900 fluorometer made by Edinburgh.

Toluene and benzene were distilled over CaH₂. THF was distilled from sodium/benzophenone before use. Other reagents were used as received from Aldrich Chemical Co. without further purification unless specified otherwise. Acetic acid, isopropanol and methanol were used as received from Fisher without further purification. All the reactions were performed under argon atmosphere. Diethynylhexylsilane was synthesized by reaction of ethynylmagnesium bromide with dichlorodihexylsilane. 2,5-

Bis(trimethylstannyl)thiophene was synthesized by reaction of 2,5-dibromothiophene with 2 equivalents of n-BuLi, followed by quenching with chlorotrimethyltin. 3,3’-Dibromo-2,2’-bithiophene 44 was synthesized according to the literature, starting from 2-bromothiophene.¹⁰⁵

3,3-Dihexylsilylene-2,2’-bithiophene 47: A flask was charged with n-BuLi (60 mmol, 24 ml, 2.5 M in hexane) and THF (500 ml). The solution was cooled to -78°C and a solution of 3,3’-dibromo-2,2’-bithiophene (30 mmol, 9.720 g) in THF (100 ml) was added dropwise. The solution was stirred at -78°C for 1 hour, resulting in a white precipitate. Then a solution of dichlorodihexylsilane (30 mmol, 8.070 g) in THF (100 ml) was added. The solution was stirred at -78°C for 5 hours, and then was allowed to warm to room temperature and was stirred overnight. A saturated NH₄Cl solution (300 ml) was added. The aqueous solution was extracted with Et₂O (300 ml). The combined organic layer was washed with water and dried (MgSO₄). After the solvents were removed, the product was purified by column chromatography to yield a bright green-yellow liquid (8.812 g, yield 81.2%). Mass: cal. m/z for C₂₀H₃₀S₂Si = 362.67570, measured (HiRes EI): 362.67532; IR: ν (cm⁻¹) 3070, 2950, 2844, 1462; ¹H-NMR: δ 7.20 (d, 2H, J = 6 Hz), 7.06 (d, 2H, J = 6 Hz), 1.27 (br, 16H), 0.92 (t, 6H, J = 9 Hz), 0.63 (d, 4H, J = 6 Hz); ¹³C-NMR: δ 149.17, 141.53, 129.58,
130

124.91, 32.83, 31.43, 24.15, 22.54, 14.08, 11.88; Element analysis: cal. for C_{20}H_{30}S_{2}Si, C 66.24; H 8.34; S 17.68; Si 7.74, found C 66.10; H 8.40

5,5'-Dibromo-3,3'-dihexylsilylene-2,2'-bithiophene 48: To a solution of 3,3'-dihexylsilylene-2,2'-bithiophene 47 (5 mmol, 1.810 g) in DMF (40 ml) was added NBS (11 mmol, 1.958 g). After the mixture was stirred at room temperature for 20 minutes only (don’t exceed 20 minutes), water (50 ml) was added. The solution was extracted with EtO (50 ml x 3). The combined organic layer was washed with water (50 ml) and dried (MgSO₄). After ether was removed, the product was purified by column chromatography (hexane as eluent) to yield a green-yellow liquid (2.374 g, yield 91.3%). Mass: cal. m/z for C_{20}H_{28}^{79}Br_{2}^{81}BrS_{2}Si = 519.97493, C_{20}H_{28}^{79}Br_{2}S_{2}Si = 517.97786, C_{20}H_{28}^{81}Br_{2}S_{2}Si = 521.97301, measured (HiRes EI) = 519.97466, 517.97655, 521.97301, respectively; IR: ν (cm⁻¹) 3074, 2952, 2918, 2850, 1667, 1462; ¹H-NMR: δ 7.00 (s, 2H), 1.32-1.22 (m, 16H), 0.88 (t, 4H, J = 6 Hz), 0.86 (t, 6H, J = 3 Hz); ¹³C-NMR: δ 148.90, 140.96, 132.12, 111.40, 32.77, 31.35, 23.96, 22.52, 14.05, 11.62; Element analysis: for C_{20}H_{28}Br_{2}S_{2}Si, C 64.62; H 5.42; Br 30.70; S 12.32; Si 5.40, found C 64.13; H 5.52

Polymer 49: The monomers, 2,5-bis(trimethylstannyl)thiophene 25 (1.5 mmol, 0.614 g) and 5,5'-dibromo-3,3'-dihexylsilylene-2,2'-bithiophene 48 (1.5 mmol, 0.780 g), and the catalyst Pd(PPh₃)₂Cl₂ (55 mg) were dissolved in THF (50 ml) and the solution was refluxed for 7 days. After the solvent was removed, the polymer was purified by precipitation from MeOH/THF twice to yield a black shining solid (0.630 g, yield 95%). GPC: Mₙ = 3.16 x 10⁴, PD = 3.21, n = 67.7; IR: ν (cm⁻¹) 3056, 2956, 2923, 2853, 1546; ¹H-NMR: δ 7.11 (br, 4H), 1.29 (br, 20H), 0.89 (br, 6H); ¹³C-NMR: δ 147.29, 142.50, 138.37, 135.00, 125.40, 124.59, 32.41, 30.92, 23.60, 22.06, 13.57, 11.24; Element analysis: cal. for (C_{24}H_{30}S_{3}Si)ₙ, C 65.10; H 6.83; S 21.72; Si 6.34, found C 64.84, H 6.85; UV/Vis: λ_max (THF) = 556, 574 nm, λ_max (film) = 582 nm; Emission: λ_max = 627, 662 nm (λ_excit. = 510 nm); Electrical conductivity: ~400 S/cm (doped by I₂ vapor)
Polymer 51: To a solution of 5,5'-dibromo-3,3'-dihexylsilylene-2,2'-bithiophene 48 (0.773 mmol, 0.402 g) in THF (20 ml) was added n-BuLi (1.85 mmol, 0.74 ml, 2.5 M in hexane) dropwise at -78°C. After addition, the mixture was stirred at -78°C for one half hour, resulting in a white-yellow precipitate. Then chlorotrimethyltin (2.411 mmol, 2.411 ml, 1.0 M in hexane) was added. The mixture was allowed to warm to room temperature and stirred for 2 hours. Then the solvent and BuBr were removed under vacuum. A solution of 5,5'-dibromo-3,3'-dihexylsilylene-2,2'-bithiophene 48 (0.773 mmol, 0.402 g) and the catalyst Pd(PPh3)2Cl2 (56 mg) in THF (20 ml) was added. The mixture was refluxed for 7 days, resulting in a deep purple solution. After the solution was cooled down, water (30 ml) was added. The aqueous layer was extracted by CH2Cl2 (30 ml x 2). The combined organic layer was washed with water (40 ml) and dried with MgSO4. After the solvent was removed, the residue was dissolved in a minimum amount of THF and was added into methanol (200 ml). A deep purple (like black) solid was precipitated and collected by vacuum filtration. The polymer was purified by precipitation one more time from THF/MEOH (0.486 g, yield 87.4%). GPC: Mw = 2.03 x 10^4, PD = 1.896, n = 56.29; IR: ν (cm⁻¹) 3054, 2957, 2923, 2854, 1549, 1464; ¹H-NMR: δ 7.10 (br, 2H), 1.24 (br, 20H), 0.93-0.85 (br, 6H); ¹³C-NMR: δ 147.16, 142.75, 138.23, 125.68, 32.65, 31.16, 23.85, 22.33, 13.85, 11.56; Element analysis: cal. for (C26H28S2Si)n, C 66.61; H 7.83; S 17.78; Si 7.78, found C 66.34, H 7.50; UV/Vis: λmax (THF) = 582 nm, λmax (film) = 592 nm; Emission: λmax = 680 nm (λexcit. = 540 nm); Electrical conductivity: ~600 S/cm (doped by I2 vapor)

Di(3-propynyl)-octylamine 53: To a solution of n-octylamine (0.20 mol, 25.850 g) and propargyl bromide (0.42 mol, 49.980 g) in toluene (600 ml) was added K2CO3 (0.80 mmol, 110.400 g) and tetrabutylammonium bromide (0.20 mol, 6.440 g). The mixture was stirred at 70°C for 8 hours. Then KOH (0.20 mol, 11.200 g) was added and the mixture was kept stirring at 70°C overnight. The resulting white solid was removed by filtration. The solvent was distilled off and the product was distilled off under vacuum (b.p. 87-90°C/0.15 mmHg) to yield a colorless liquid (31.063 g, yield 75.8%). Mass: cal. m/z for C14H23N = 205.18305, measured (HiRes EI) = 205.18256; IR: ν (cm⁻¹) 3305, 2951, 2924, 2819, 1711, 1658; ¹H-NMR: δ 7.00 (br, 2H), 1.24 (br, 20H), 0.93-0.85 (br, 6H); ¹³C-NMR: δ 142.75, 138.23, 125.68, 32.65, 31.16, 23.85, 22.33, 13.85, 11.56; UV/Vis: λmax (THF) = 320 nm, λmax (film) = 330 nm; Electrical conductivity: ~200 S/cm (doped by I2 vapor)
\begin{align*}
132; & \quad ^1{H}-\text{NMR: } \delta 3.404 (d, 4H, J = 2.4 \text{ Hz}), 2.477 (t, 2H, J = 7.5 \text{ Hz}), 2.184 (t, 2H, J = 2.4 \\
& \quad \text{Hz}), 1.451-1.386 (m, 2H), 1.247 (br, 10H), 0.845 (t, 3H, J = 6.9 \text{ Hz}); ^{13}{C}-\text{NMR: } \delta 78.795, \\
& \quad 72.783, 52.973, 42.003, 31.788, 29.416, 29.1888, 27.286, 22.613, 14.058
\end{align*}

\textbf{Di-[1-(2-thienyl)-3-propynyl]-octylamine 54:} To a degassed solution of di-(3-propynyl)-octylamine \textbf{53} (20 mmol, 4.100 g) and 2-halothiophene (42 mmol) in triethylamine (200 ml) was added the catalysts \text{Pd}(\text{PPh}_3)_2\text{Cl}_2 (0.140 g)/\text{CuI} (0.076 g). The solution was stirred at room temperature for 36 hours. The resulting salt was removed by filtration. After the solvent was removed, the residue was directly purified by column chromatography to yield a liquid (for bromothiophene: 3.246 g, yield 44%; for iodothiophene: 6.639 g, yield 90%).

\text{Mass: cal. } m/z \text{ for } C_{22}H_{27}NS_2 = 369.15850, \text{ measured (HiRes EI) } = 369.15622; \quad \text{IR: } \nu (\text{cm}^{-1}) 3100, 2908, 2842, 2224, 1410; \quad ^1{H}-\text{NMR: } \delta 7.21 (dd, 2H, J = 3 \text{ Hz}), 7.19 (s, 2H), 6.95 (dd, 2H, J = 3 and 3 Hz), 3.72 (s, 4H), 2.64 (t, 2H, J = 6 Hz), 1.56 (m, 2H), 1.34-1.26 (m, 10H), 0.91 (t, 3H, J = 9 Hz); ^{13}{C}-\text{NMR: } \delta 131.67, 126.63, 126.50, 122.98, 88.73, 78.11, 53.03, 43.30, 31.68, 29.30, 29.09, 27.35, 27.21, 22.51, 13.98

\textbf{Compound 55:} A suspension of magnesium turnings (166.6 mmol, 4.000 g) and mercuric chloride (16.810 mmol, 4.566 g) in THF (200 ml) was stirred for 15 minutes. A solution of bis(cyclopentadienyl)zirconium dichloride (30.649 mmol, 8.950 g) and di-[1-(2-thienyl)-3-propynyl]-octylamine \textbf{54} (25.626 mmol, 9.456 g) in THF (200 ml) was added rapidly dropwise, and the mixture was stirred overnight at room temperature. The reaction solution was decanted from the unreacted magnesium, then rapidly quenched with HCl (3 M, 100 ml), and extracted with CH\textsubscript{2}Cl\textsubscript{2} (150 ml x 3). The organic layer was washed with water (300 ml) and a saturated NaHCO\textsubscript{3} (300 ml) solution, and dried with MgSO\textsubscript{4}. After the solvent was removed, the residue was dissolved in a small amount of CH\textsubscript{2}Cl\textsubscript{2}, and MeOH was added to precipitate a yellow solid (7.14 g, yield 80%). m.p. 65-68\degree C; \text{Mass: cal. } m/z \text{ for } C_{22}H_{29}NS_2 = 371.17415, \text{ measured (HiRes EI) } = 371.17507; \quad \text{IR: } \nu (\text{cm}^{-1}) 3006, 2917, 2750, 1617, 1495; \quad ^1{H}-\text{NMR: } \delta 7.31 (d, 2H, J = 3 \text{ Hz}), 7.09 (s, 2H), 7.07-7.03 (m, 4H), 3.75 (s, 4H), 2.69 (t, 2H, J = 6 Hz), 1.64 (m, 2H), 1.34 (m, 10H), 0.90 (t, 3H, J = 6 Hz); ^{13}{C}-\text{NMR:}
To a degassed solution of diethynyldihexylsilane (3.92 mmol, 0.973 g) and 2-iodothiophene (8.28 mmol, 1.950 g) in triethylamine (30 ml) was added the catalysts Pd(PPh3)2Cl2 (56 mg)/CuI (18 mg). The mixture was stirred at room temperature for 18 hours. The resulting white precipitate was removed by filtration. After the solvent was removed, the residue was purified by column chromatography (hexane/ethyl acetate = 10:1 as eluents) to yield a colorless liquid (1.400 g, 86.7%). Mass: cal. $m/z$ for C24H32S2Si = 412.17148, measured (HiRes EI) = 412.17084; IR: $\nu$ (cm$^{-1}$) 3110, 2964, 2857, 2148, 1510, 1468; $^1$H-NMR: $\delta$ 7.31 (dd, 2H, $\gamma = 3$ and 1 Hz), 7.28 (dd, 2H, $\gamma = 6$ and 1 Hz), 6.99 (dd, 2H, $\gamma = 6$ and 3 Hz), 1.60 (m, 4H), 1.49-1.36 (m, 16H), 0.95 (t, 5H, $\gamma = 6$ Hz); $^{13}$C-NMR: $\delta$ 133.13, 127.72, 126.82, 122.80, 99.24, 93.81, 32.62, 31.45, 23.57, 14.59, 14.12

To a solution of bis(cyclopentadienyl)zirconium dichloride (4.11 mmol, 1.200 g) in THF (30 ml) was added EtMgBr (8.22 mmol, 8.22 ml, 1.0 M in THF) dropwise at -78°C. The solution was stirred at -78°C for 1 hour. A solution of dihexyl-di(2-thienylethyny)silane 56 (3.29 mmol, 1.358 g) in THF (20 ml) was added dropwise. The mixture was allowed to warm to room temperature and was stirred for 1 hour. The reaction mixture was then quenched with 3N HCl, extracted with diethyl ether, and washed with water, NaHCO3, and water. The extract was dried over MgSO4. After the solvent was removed, the residue was purified by column chromatography to afford a yellow-green liquid (0.605 g, yield 44.4%). Mass: cal. $m/z$ for C24H34S2Si = 424.18713, measured (HiRes EI) = 414.18656; IR: $\nu$ (cm$^{-1}$) 3110, 2954, 2920, 2852, 1611, 1461; $^1$H-NMR: $\delta$ 7.62 (s, 1H), 7.31 (d, 2H, $\gamma = 6$ Hz), 7.14 (d, 2H, $\gamma = 3$ Hz), 7.09-6.95 (m, 5H), 1.57 (m, 4H), 1.42-1.28 (m, 16H), 0.90 (t, 6H, $\gamma = 6$ Hz); $^{13}$C-NMR: $\delta$ 148.04, 147.47, 145.76, 144.27, 142.21, 127.66, 127.60, 126.51, 125.76, 125.11, 122.86, 121.39, 32.77, 31.34, 23.75, 22.52, 14.22, 14.02; UV/Vis: $\lambda_{\text{max}}$ (THF) = 372 nm, 356 nm (shoulder), 388 nm (shoulder)
**Polymer 60**: To a degassed solution of FeCl₃ (2.06 mmol, 0.336 g) in CHCl₃ (12 ml) was added monomer 55 (0.516 mmol, 0.191 g). Argon was purged through the solution during the reaction. The solution was stirred at room temperature for 4 hours. The resulting blue-black solid was collected by filtration and washed with MeOH. The polymer was undoped by stirred the solid in hydrazine/CH₂Cl₂ solution for 2 hours. Then water was added. The organic layer was washed with water and dried (MgSO₄). After the solvent was removed, the polymer was precipitated from MeOH to yield a dark red solid (0.086 g, 46.0%). GPC: $M_w = 1.39 \times 10^4$, $M_z = 2.37 \times 10^4$, $PD = 1.84$; IR: $\nu$ (cm⁻¹) 3005, 2915, 2754, 1615, 1495; $^1$H-NMR: $\delta$ 7.70, 7.52, 7.02, 4.21, 3.74, 1.67, 1.30, 0.90; $^{13}$C-NMR: $\delta$ 137.70, 132.44, 130.88, 124.08, 111.78, 68.16, 58.50, 56.66, 38.37, 30.37, 28.94, 23.75, 22.99, 14.09, 10.97; UV/Vis: $\lambda_{max}$ (THF) = 514 nm, $\lambda_{max}$ (film) = 534 nm; Emission: $\lambda_{max} = 554$ nm ($\lambda_{excit.} = 440$ nm); Electrical conductivity: $\sim 2$ S/cm (doped by I₂)
FUTURE WORK

From our study of the above polymers, we learned that the silicon unit on the main chain decreases the bandgap energy. Polysilole, which was predicted as a promising polymer for high conductivity and third-order nonlinearity,\(^{57,58}\) is still a synthetic challenge for chemists. It was found by calculations that the silicon unit decreases the LUMO by \(\sigma^*-\pi^*\) conjugation.\(^{62}\) The introduction of an electron donating group, such as O or S, into polymer side chain was presumed to be able to decrease the bandgap by raising the energy of the HOMO.\(^{115}\) We thus designed the sulfur-bridged polysilole (Scheme 45). The combination of sulfur and silicon in the main chain should provide a polymer having high environmental stability and a very low bandgap.

The synthesis of sulfur-bridged silole was designed as shown in Scheme 45. We have already synthesized compound 64 (R = Ph, R’ = Me), starting from 3,4-dibromothiophene. 3,4-Dibromothiophene was treated with 1 equivalent of n-BuLi, followed by quenching with 1-chlorodimethylsilyl-2-phenylethyne 62 (R = Ph, R’ = Me) to give compound 63. Repeat this step on 63 gave compound 64 (R = Ph, R’ = Me) in ~80% overall yield. Direct treatment of 3,4-dibromothiophene with 2 equivalents or excess n-BuLi, followed by quenching with 62 in one step failed to give 64, but instead gave a mixture. Compound 64 (R = Ph, R’ = Me) was treated with excess LDA, followed by quenching with chlorotrimethylsilane to give compound 65 (R = Ph, R’ = Me, X = SiMe3) in ~76% yield. Based on this result, compound 65 (X = Br or I) should be obtainable when the reaction is quenched by Br2 or I2. The key step is intramolecular cyclization of 65 (X = Br, I) by palladium catalyst. We rationalize that palladium inserts into the thiophene-halogen bond, complexes with the triple bond, and then isomerizes to a cyclic intermediate, which could be quenched with Me3SnSnMe3 to give the cyclic product 66 (Scheme 46). Once 66 is obtained, it can be easily transformed into 67 by reacting with I2. 67 can be used to couple with 66 in the presence of a palladium catalyst to give the desired polymer 68. Polymer 68 is a promising polymer having a very low bandgap and high conductivity and third nonlinearity. Furthermore, desulfurization of 68 will give polysilole.
The other promising polymer is a ladder polymer containing sulfur and silicon as shown in Figure 56. The planar structure can be enhanced by the sulfur and silicon bridges. The long alkyl groups on the silicon can provide good solubility of the polymer. Polymer 69 should exhibit better properties than 68.

Scheme 45: Design of synthesis of sulfur-bridged polysilole
Scheme 46: Proposed mechanism for intramolecular cyclization of 65

Figure 56: Design of a ladder polymer
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